Ethyl Acetate

Other names:	1-Acetoxyethane
	Acetic acid, ethyl ester
	Acetic ether
	Acetidin
	Acetoxyethane
	Aethylacetat
	CH3COOC2H5
	Essigester
	Ethyl acetic ester
	Ethyl ester of acetic acid
	Ethyl ethanoate
	Ethylacetaat
	Ethyle (acetate d')
	Ethylester kyseliny octove
	Etile (acetato di)
	NSC 70930
	Rcra waste number U112
	UN 1173
	Vinegar naphtha
	ac. acetic ethyl ester
Inchi:	InChI=1S/C4H8O2/c1-3-6-4(2)5/h3H2,1-2H3
InchiKey:	XEKOWRVHYACXOJ-UHFFFAOYSA-N
Formula:	C4H8O2
SMILES:	CCOC(C)=O
Mol. weight [g/mol]:	88.11
CAS:	141-78-6

Physical Properties

Property code	Value	Unit	Source
af	0.3620		KDB
affp	835.70	kJ/mol	NIST Webbook
aigt	699.82	K	KDB
basg	804.70	kJ/mol	NIST Webbook
basg	799.90 ± 0.20	kJ/mol	NIST Webbook
chl	-2235.40 ± 3.90	kJ/mol	NIST Webbook
chl	-2238.54 ± 0.48	kJ/mol	NIST Webbook
chl	-2246.00	kJ/mol	NIST Webbook

chl	-2256.00	kJ/mol	NIST Webbook
dm	1.90	debye	KDB
dvisc	0.0004370	Pa×s	A volumetric and viscosity study for the binary mixtures of 1-hexyl-3-methylimidazolium tetrafluoroborate with some molecular solvents
dvisc	0.0004274	Pa×s	Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents
dvisc	0.0004260	Pa×s	Densities and Viscosities of Ternary Mixtures of Cyclohexane + Cyclohexanone + Some Alkyl Acetates at 298.15 K
fll	2.20	% in Air	KDB
flu	9.00	% in Air	KDB
fpc	285.93	K	KDB
fpo	268.71	K	KDB
gf	-327.60	kJ/mol	KDB
gyrad	3.3480		KDB
hf	-445.43 ± 0.84	kJ/mol	NIST Webbook
hf	-443.20	kJ/mol	KDB
hf	-446.90	kJ/mol	NIST Webbook
hf	-443.80	kJ/mol	NIST Webbook
hf	-444.80 ± 0.40	kJ/mol	NIST Webbook
hfl	-479.86 ± 0.46	kJ/mol	NIST Webbook
hfl	-480.57 ± 0.79	kJ/mol	NIST Webbook
hfl	-482.00 ± 4.00	kJ/mol	NIST Webbook
hfl	-478.82 ± 0.73	kJ/mol	NIST Webbook
hfus	8.90	kJ/mol	Joback Method
hvap	33.65	kJ/mol	Joback Method
ie	10.09 ± 0.02	eV	NIST Webbook
ie	9.90 ± 0.05	eV	NIST Webbook
ie	10.01 ± 0.05	eV	NIST Webbook
ie	9.90	eV	NIST Webbook
ie	10.01 ± 0.05	eV	NIST Webbook
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ie	10.24	eV	NIST Webbook
ie	10.11 ± 0.02	eV	NIST Webbook
ie	10.45	eV	NIST Webbook
ie	10.16	eV	NIST Webbook
log10ws	-0.04		Aqueous Solubility Prediction Method

log10ws	-0.04		Estimated Solubility Method
logp	0.569		Crippen Method
mcvol	74.660	ml/mol	McGowan Method
nfpaf	%!d(float64=3)		KDB
pc	3830.00 ± 81.06	kPa	NIST Webbook
pc	3882.00 ± 3.87	kPa	NIST Webbook
pc	3900.00	kPa	Critical Properties of the Reacting Mixture in the Esterification of Acetic Acid with Ethanol
pc	3851.70 ± 40.00	kPa	NIST Webbook
pc	4018.00 ± 202.65	kPa	NIST Webbook
pc	4280.00 ± 405.30	kPa	NIST Webbook
pc	3882.00	kPa	KDB
rhoc	307.66 ± 5.29	kg/m3	NIST Webbook
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tb	350.70 ± 2.00	K	NIST Webbook
tb	350.10 ± 2.00	K	NIST Webbook
tb	350.30 ± 2.00	K	NIST Webbook

tb	350.20 ± 2.00	K	NIST Webbook
tb	350.30 ± 2.00	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.30 ± 0.20	K	NIST Webbook
tb	350.16	K	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tb	347.45 ± 2.00	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.26	K	KDB
tb	350.27	K	Vapor liquid equilibria for the quaternary reactive system ethyl acetate + ethanol +water + acetic acid and some of the constituent binary systems at 101.3 kPa
tb	350.24	K	Vapor liquid equilibria for the binary mixtures of 2,3-butanediol with n-butanol, n-butyl acetate, and ethyl acetate at 101.3 kPa
tb	350.24	K	Liquid-liquid equilibria of water + 3-hydroxy-2-butanone + ethyl ethanoate
tb	350.20	K	Solubility and tie-line data for ternary aqueous mixtures of cyclopentanol with organic solvents at T = 298.2 K: Experiments and NRTL model
tb	350.15	K	The isobaric vapor liquid equilibria of ethyl acetate p acetonitrile p bis(trifluoromethylsulfonyl)imide-based ionic liquids at 101.3 kPa
tb	350.29	K	Isobaric vapor-liquid equilibrium of a ternary system of ethyl acetate + propyl acetate + dimethyl sulfoxide and binary systems of ethyl acetate + dimethyl sulfoxide and propyl acetate + dimethyl sulfoxide at 101.3 kPa
tb	350.35	K	Isobaric Vapor - Liquid Equilibrium for Ethyl acetate + Methanol + Ionic Liquids Ternary systems at 101.3 kPa
tb	349.90 ± 0.20	K	NIST Webbook

tb	350.20	K	Isobaric Vapor Liquid Equilibrium for Three Binary Systems of Acetaldehyde + Ethanol, Ethyl Acetate, 1-Butanol at 101.3 kPa
tb	350.25 ± 1.00	K	NIST Webbook
tb	350.28	K	Isobaric Vapor-Liquid Phase Equilibrium Measurements for Allyl Alcohol with Chloroform, Ethyl Acetate, and Methyl Propionate at 101.3 kPa
tb	350.25	K	Isobaric Vapor-Liquid Equilibria and Excess Quantities for Binary Mixtures of an Ethyl Ester + tert-Butanol and a New Approach to VLE Data Processing
tb	350.15	K	Vapor-Liquid Equilibrium of Binary Mixtures Containing Ethyl Acetate + 2-Methyl-1-propanol and Ethyl Acetate + 2-Methyl-1-butanol at 101.3 kPa
tb	350.21	K	Volumetric Behavior and Saturated Pressure for Carbon Dioxide + Ethyl Acetate at a Temperature of 313.15 K
tb	350.15	K	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
tb	350.19	K	Experimental Determination of Vapor Liquid Equilibria. Binary Systems of Methyl Acetate, Ethyl Acetate, and Propyl Acetate with 1-Propanol at 0.6 MPa
tb	350.13 ± 0.30	K	NIST Webbook
tb	350.20	K	NIST Webbook
tb	350.00	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	349.25 ± 0.30	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.30 ± 0.20	K	NIST Webbook
tb	350.30	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.19 ± 0.50	K	NIST Webbook

tb	350.25 ± 0.30	K	NIST Webbook
tb	350.30 ± 0.20	K	NIST Webbook
tb	350.30 ± 0.30	K	NIST Webbook
tb	375.00 ± 2.00	K	NIST Webbook
tb	350.35	K	NIST Webbook
tb	350.05 ± 0.30	K	NIST Webbook
tb	350.21 ± 0.50	K	NIST Webbook
tb	350.25 ± 0.50	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.30 ± 0.30	K	NIST Webbook
tb	350.29 ± 0.15	K	NIST Webbook
tb	350.15 ± 2.00	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.25 ± 0.50	K	NIST Webbook
tb	350.25 ± 0.30	K	NIST Webbook
tb	350.65 ± 2.00	K	NIST Webbook
tb	349.65 ± 2.00	K	NIST Webbook
tb	352.15 ± 2.00	K	NIST Webbook
tb	350.25 ± 1.00	K	NIST Webbook
tb	350.30 ± 1.00	K	NIST Webbook
tb	350.30 ± 1.00	K	NIST Webbook
tb	349.97 ± 2.00	K	NIST Webbook
tb	351.15 ± 2.00	K	NIST Webbook
tb	350.45 ± 1.00	K	NIST Webbook
tb	350.25 ± 1.00	K	NIST Webbook
tb	350.30 ± 0.40	K	NIST Webbook
tb	349.15 ± 2.00	K	NIST Webbook
tb	349.65 ± 2.50	K	NIST Webbook
tb	350.26 ± 0.06	K	NIST Webbook
tb	350.30 ± 0.15	K	NIST Webbook
tb	350.00 ± 2.00	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.15 ± 2.00	K	NIST Webbook
tb	349.95 ± 2.00	K	NIST Webbook
tb	350.20 ± 0.50	K	NIST Webbook
tb	350.30 ± 0.20	K	NIST Webbook
tb	350.25 ± 0.50	K	NIST Webbook
tb	350.21 ± 0.50	K	NIST Webbook
tb	349.65 ± 2.00	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.30 ± 0.50	K	NIST Webbook
tb	350.45 ± 0.60	K	NIST Webbook

tb	350.00 ± 2.00	K	NIST Webbook
tb	350.25 ± 1.00	K	NIST Webbook
tb	350.30 ± 0.20	K	NIST Webbook
tc	548.90 ± 20.00	K	NIST Webbook
tc	523.20	K	NIST Webbook
tc	523.30 ± 0.05	K	NIST Webbook
tc	522.70 ± 2.00	K	NIST Webbook
tc	513.00 ± 6.00	K	NIST Webbook
tc	523.30	K	KDB
tc	523.29	K	Development of a Predictive Equation of State for CO2 + Ethyl Ester Mixtures Based on Critical Points Measurements
tc	523.30 ± 1.00	K	NIST Webbook
tc	523.30 ± 1.00	K	NIST Webbook
tf	189.25	K	Aqueous Solubility Prediction Method
tf	189.50	K	KDB
tt	189.30 ± 0.20	K	NIST Webbook
tt	189.30 ± 0.05	K	NIST Webbook
vc	0.286	m3/kmol	KDB
zc	0.2551730		KDB
zra	0.25		KDB

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	149.47	J/mol×K	450.00	NIST Webbook
cpg	131.06	J/mol×K	380.00	NIST Webbook
cpg	125.82	J/mol×K	360.00	NIST Webbook
cpg	142.80	J/mol×K	425.00	NIST Webbook
cpg	136.22	J/mol×K	400.00	NIST Webbook
cpl	169.60	J/mol×K	298.15	NIST Webbook
cpl	169.60	J/mol×K	298.15	NIST Webbook
cpl	169.06	J/mol×K	298.15	NIST Webbook
cpl	169.30	J/mol×K	298.15	NIST Webbook
cpl	169.50	J/mol×K	298.10	NIST Webbook
cpl	168.82	J/mol×K	303.61	NIST Webbook
cpl	157.70	J/mol×K	290.00	NIST Webbook
cpl	169.20	J/mol×K	293.60	NIST Webbook
cpl	170.59	J/mol×K	298.32	NIST Webbook

cpl	168.94	J/mol×K	298.15	NIST Webbook	
cpl	167.40	J/mol×K	298.15	NIST Webbook	
dvisc	0.0004520	Pa×s	293.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure	
dvisc	0.0004280	Pa×s	298.15	Densities, Excess Molar Volumes, Viscosities, Speeds of Sound, Excess Isentropic Compressibilities, and Relative Permittivities for Alkyl (Methyl, Ethyl, Butyl, and Isoamyl) Acetates + Glycols at Different Temperatures	
dvisc	0.0003870	Pa×s	308.15	Densities, Excess Molar Volumes, Viscosities, Speeds of Sound, Excess Isentropic Compressibilities, and Relative Permittivities for Alkyl (Methyl, Ethyl, Butyl, and Isoamyl) Acetates + Glycols at Different Temperatures	

dvisc	0.0004030	Pa×s	303.15	Density, dynamic viscosity, and derived properties of binary mixtures of methanol or ethanol with water, ethyl acetate, and methyl acetate at T = (293.15, 298.15, and 303.15) K	
dvisc	0.0004260	Pa×s	298.15	Density, dynamic viscosity, and derived properties of binary mixtures of methanol or ethanol with water, ethyl acetate, and methyl acetate at T = (293.15, 298.15, and 303.15) K	
dvisc	0.0004030	Pa×s	303.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure	
dvisc	0.0004430	Pa×s	293.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0004210	Pa×s	298.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003990	Pa×s	303.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	

dvisc	0.0003810	Pa×s	308.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003630	Pa×s	313.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003460	Pa×s	318.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003300	Pa×s	323.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003150	Pa×s	328.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003010	Pa×s	333.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0002880	Pa×s	338.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0002750	Pa×s	343.15	Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters	
dvisc	0.0003806	Pa×s	303.15	Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Sulfolane with Ethyl Acetate, n-Propyl Acetate, and n-Butyl Acetate at Temperature of (303.15, 308.15, and 313.15) K	

dvisc	0.0003622	Pa×s	308.15	Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Sulfolane with Ethyl Acetate, n-Propyl Acetate, and n-Butyl Acetate at Temperature of (303.15, 308.15, and 313.15) K	
dvisc	0.0003426	Pa×s	313.15	Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Sulfolane with Ethyl Acetate, n-Propyl Acetate, and n-Butyl Acetate at Temperature of (303.15, 308.15, and 313.15) K	
dvisc	0.0003790	Pa×s	303.15	Volumetric and Transport Properties of Binary Liquid Mixtures of Phenylacetonitrile with Aliphatic Esters at Temperatures of (303.15 to 313.15) K	
dvisc	0.0003610	Pa×s	308.15	Volumetric and Transport Properties of Binary Liquid Mixtures of Phenylacetonitrile with Aliphatic Esters at Temperatures of (303.15 to 313.15) K	
dvisc	0.0003440	Pa×s	313.15	Volumetric and Transport Properties of Binary Liquid Mixtures of Phenylacetonitrile with Aliphatic Esters at Temperatures of (303.15 to 313.15) K	

dvisc	0.0004520	Pa×s	293.15	Density, dynamic viscosity, and derived properties of binary mixtures of methanol or ethanol with water, ethyl acetate, and methyl acetate at T = (293.15, 298.15, and 303.15) K
dvisc	0.0004260	Pa×s	298.15	Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure
hfust	10.48	kJ/mol	189.30	NIST Webbook
hfust	10.48	kJ/mol	189.30	NIST Webbook
hfust	10.48	kJ/mol	189.30	NIST Webbook
hvapt	32.40 ± 0.10	kJ/mol	344.00	NIST Webbook
hvapt	31.90	kJ/mol	350.00	NIST Webbook
hvapt	34.00	kJ/mol	320.00	NIST Webbook
hvapt	31.00 ± 0.10	kJ/mol	363.00	NIST Webbook
hvapt	31.90 ± 0.10	kJ/mol	351.00	NIST Webbook
hvapt	33.80 ± 0.10	kJ/mol	326.00	NIST Webbook
hvapt	33.40 ± 0.10	kJ/mol	331.00	NIST Webbook
hvapt	31.40 ± 0.10	kJ/mol	343.00	NIST Webbook
hvapt	34.60 ± 0.10	kJ/mol	313.00	NIST Webbook
hvapt	36.70	kJ/mol	322.00	NIST Webbook
hvapt	35.70	kJ/mol	319.50	NIST Webbook
hvapt	34.10	kJ/mol	345.00	NIST Webbook
hvapt	31.94	kJ/mol	350.30	NIST Webbook
hvapt	32.22	kJ/mol	349.80	KDB
pvap	136.34	kPa	359.44	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	24.78	kPa	313.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	21.72	kPa	310.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	687.53	kPa	423.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	202.90	kPa	373.15	Vapor Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation
pvap	195.35	kPa	371.50	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	190.92	kPa	370.69	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	187.47	kPa	370.07	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	184.43	kPa	369.49	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	180.88	kPa	368.83	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	176.37	kPa	367.98	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	174.65	kPa	367.64	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	171.33	kPa	367.00	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	167.81	kPa	366.29	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	28.18	kPa	316.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	160.53	kPa	364.81	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	157.94	kPa	364.25	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	155.58	kPa	363.72	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	150.43	kPa	362.62	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	146.37	kPa	361.73	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	142.55	kPa	360.88	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	139.07	kPa	360.03	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	31.95	kPa	319.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	133.45	kPa	358.74	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	130.00	kPa	357.92	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	127.89	kPa	357.35	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	124.00	kPa	356.37	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	122.31	kPa	355.91	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	120.68	kPa	355.50	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	118.56	kPa	354.93	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	116.46	kPa	354.43	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	114.65	kPa	353.91	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	111.21	kPa	352.94	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	108.13	kPa	352.12	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	106.84	kPa	351.77	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	104.65	kPa	351.12	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	101.32	kPa	350.15	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	100.68	kPa	349.93	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	100.07	kPa	349.76	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	99.36	kPa	349.56	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	98.97	kPa	349.45	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	96.74	kPa	348.75	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	94.63	kPa	348.09	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	90.81	kPa	346.86	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	86.85	kPa	345.53	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	82.56	kPa	344.11	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	79.83	kPa	343.03	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	76.02	kPa	341.69	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	72.26	kPa	340.23	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	69.34	kPa	339.09	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	36.12	kPa	322.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	40.73	kPa	325.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	45.80	kPa	328.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	51.36	kPa	331.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	57.46	kPa	334.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	64.13	kPa	337.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	71.40	kPa	340.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa

pvap	79.32	kPa	343.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	87.93	kPa	346.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	97.26	kPa	349.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	107.36	kPa	352.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	66.49	kPa	337.91	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	62.83	kPa	336.33	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	59.91	kPa	335.02	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)

pvap	55.72	kPa	333.07	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	52.56	kPa	331.51	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	49.27	kPa	329.77	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	44.98	kPa	327.41	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	40.61	kPa	324.80	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	38.43	kPa	323.47	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	34.85	kPa	321.01	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	31.33	kPa	318.46	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	27.63	kPa	315.44	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	24.34	kPa	312.53	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	22.44	kPa	310.67	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	
pvap	20.64	kPa	308.86	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)	

pvap	18.33	kPa	306.25	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	16.04	kPa	303.35	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	11.89	kPa	297.17	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
pvap	25.30	kPa	313.15	Isothermal Vapor-Liquid Equilibria of ethyl acetate + dibromomethane, or + bromochloromethane or + 1,2-dichloroethane or +1-bromo-2-chloroethane at T = 313.15 K
pvap	25.08	kPa	313.15	Total Vapor Pressure Measurements for 2-Ethoxyethanol with Methyl Acetate, Ethyl Acetate, Propyl Acetate, and Ethyl Propionate at 313.15 K and for 2-Ethoxyethanol with Methyl Formate at 308.15 K

pvap	50.00	kPa	330.01	Isobaric Vapor Liquid Equilibria for Binary Mixtures of Isoamyl Acetate + Ethyl Acetate at 50 and 100 kPa
pvap	100.00	kPa	349.53	Isobaric Vapor Liquid Equilibria for Binary Mixtures of Isoamyl Acetate + Ethyl Acetate at 50 and 100 kPa
pvap	118.27	kPa	355.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	130.03	kPa	358.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	134.15	kPa	359.00	Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa
pvap	101.30	kPa	350.28	Isobaric Vapor-Liquid Phase Equilibrium Measurements for Allyl Alcohol with Chloroform, Ethyl Acetate, and Methyl Propionate at 101.3 kPa
pvap	101.30	kPa	349.95	Effect of Ionic Liquids on the Binary Vapor-Liquid Equilibrium of Ethyl Acetate + Methanol System at 101.3 kPa

pvap	45.95	kPa	328.15	Isothermal Vapor-Liquid Equilibria for Binary Mixtures of Methyl Nonafluorobutyl Ether + Acetone, Cyclopentyl Methyl Ether, Ethyl Acetate, n-Heptane, Methanol, and Toluene
pvap	111.76	kPa	353.15	Vapor Liquid Equilibrium and Excess Enthalpy Data for Systems Containing N,N-Dimethylacetamide
pvap	1500.00	kPa	462.75	Measurement and modeling of high pressure VLE for methyl acetate or ethyl acetate with 2-butanol. Isobaric data at 1.5 MPa
pvap	101.30	kPa	350.35	Isobaric Vapor - Liquid Equilibrium for Ethyl acetate + Methanol + Ionic Liquids Ternary systems at 101.3 kPa
pvap	31.37	kPa	318.40	Vapour liquid equilibrium for the ethyl ethanoate + 1-butene, +cis-2-butene, +trans-2-butene, +2-methylpropene, +n-butane and +2-methylpropane
pvap	101.30	kPa	350.29	Isobaric vapor-liquid equilibrium of a ternary system of ethyl acetate + propyl acetate + dimethyl sulfoxide and binary systems of ethyl acetate + dimethyl sulfoxide and propyl acetate + dimethyl sulfoxide at 101.3 kPa

pvap	1222.70	kPa	452.30	Isothermal (vapour + liquid) equilibrium (VLE) for binary mixtures containing diethyl carbonate, phenyl acetate, diphenyl carbonate, or ethyl acetate
pvap	557.10	kPa	412.80	Isothermal (vapour + liquid) equilibrium (VLE) for binary mixtures containing diethyl carbonate, phenyl acetate, diphenyl carbonate, or ethyl acetate
pvap	205.70	kPa	373.00	Isothermal (vapour + liquid) equilibrium (VLE) for binary mixtures containing diethyl carbonate, phenyl acetate, diphenyl carbonate, or ethyl acetate
pvap	101.30	kPa	350.15	The isobaric vapor liquid equilibria of ethyl acetate p acetonitrile p bis(trifluoromethylsulfonyl)imide-based ionic liquids at 101.3 kPa
pvap	93.32	kPa	347.70	Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components
pvap	79.99	kPa	343.18	Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components

pvap	66.66	kPa	337.99	Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components
pvap	53.33	kPa	331.90	Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components
pvap	40.00	kPa	324.44	Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components
pvap	163.99	kPa	365.52	Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10)
rfi	1.36920		298.15	Thermodynamic study of (alkyl esters + a,x-alkyl dihalides) I: HE and V E for 25 binary mixtures {xCu-1H2u-1CO2C2H5 + (1-x)a,x-ClCH2(CH2)v-2CH2Cl}, where u = 1 to 5, a = 1 and v = x = 2 to 6
rfi	1.37244		298.15	Isobaric Vapor-Liquid Equilibrium Data for Binary Systems of Anisole with Methyl Acetate, Ethyl Acetate, n-Propyl Acetate, and Isopropyl Acetate at 93.9 kPa

rfi	1.36981	298.15	Measurement of VLE Data by Using an Experimental Installation with Automatic Control: Modeling of Binary Systems of Methyl Acetate or Ethyl Acetate with n-Heptane or 2,2,4-Trimethylpentane at Both 0.1 and 1.5 MPa
rfi	1.36980	298.15	Density, Refractive Index, and Speed of Sound at 298.15 K and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Ethyl Acetate + 1-Pentanol and Ethanol + 2-Methyl-1-propanol
rfi	1.37130	298.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K
rfi	1.36730	303.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K
rfi	1.36310	308.15	Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K

rfi	1.36990	298.15	Densities, Viscosities, and Refractive Indices for Binary and Ternary Mixtures of N,N-Dimethylacetamide (1) + 2-Methylbutan-2-ol (2) + Ethyl Acetate (3) at 298.15 K for the Liquid Region and at Ambient Pressure
rfi	1.36978	298.15	Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Ethyl Acetate + Ethyl Lactate and Methyl Acetate + Ethyl Lactate
rfi	1.36978	298.15	Liquid-Liquid Equilibrium Diagrams of Ethanol + Water + (Ethyl Acetate or 1-Pentanol) at Several Temperatures
rfi	1.37190	293.15	Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents
rfi	1.37241	293.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K
rfi	1.36977	298.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K
rfi	1.36712	303.15	Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K

rfi	1.37190	293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.37000	298.15	Vapor-Liquid Equilibria for Ethyl Acetate + Methanol at (0.1, 0.5, and 0.7) MPa. Measurements with a New Ebulliometer
rfi	1.37010	293.15	Liquid-Liquid Equilibrium of (Water + Pentane-2,4-dione + Ethyl Ethanoate) and (Water + Pentane-2,4-dione + Cyclohexane) at (298.15 and 313.15) K
rfi	1.37120	298.15	Determination and Correlation of Vapor Liquid Equilibrium Data for the Ethyl Acetate + Hexamethyl Disiloxane System at 101.3 kPa
rfi	1.35880	293.15	Solubilities of Triphenylphosphine Oxide in Selected Solvents
rfi	1.36983	298.15	Physical Properties of Binary and Ternary Mixtures of Ethyl Acetate, Ethanol, and 1-Octyl-3-methyl-imidazolium Bis(trifluoromethylsulfonyl)imide at 298.15 K
rfi	1.35880	293.15	Solubilities of Methyldiphenylphosphine Oxide in Selected Solvents

rfi	1.37519	288.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.36994	298.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.36460	308.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rfi	1.36980	298.15	Vapor-Liquid Equilibrium Data for Binary Mixtures of Dimethyl Carbonate with Methyl Acetate, Ethyl Acetate, n-Propyl Acetate, Isopropyl Acetate, n-Butyl Acetate, and Isoamyl Acetate at 93.13 kPa
rfi	1.35368	328.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K

rfi	1.36640	303.15	Densities, speeds of sound, isentropic compressibilities, refractive indexes, and viscosities of tetrahydrofuran with haloalkane or alkyl ethanoate at T = 303.15 K
rfi	1.36750	303.15	Liquid-liquid equilibrium for ternary systems of ethyl acetate/isopropyl acetate + 2,2,3,3-tetrafluoro-1-propanol + water at 298.15, 318.15 K
rfi	1.36720	303.15	Density, refraction index and vapor-liquid equilibria of N-methyl-2-hydroxyethylammonium butyrate plus (methyl acetate or ethyl acetate or propyl acetate) at several temperatures
rfi	1.37000	298.15	Correlation and prediction of mixing thermodynamic properties of ester-containing systems: Ester + alkane and ester + ester binary systems and the ternary dodecane + ethyl pentanoate + ethyl ethanoate
rfi	1.36977	298.15	Ternary (liquid + liquid) equilibria of the azeotrope (ethyl acetate + 2-propanol) with different ionic liquids at T = 298.15 K
rfi	1.36712	303.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures

rfi	1.36977	298.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures
rfi	1.37241	293.15	Mixing properties of binary mixtures presenting azeotropes at several temperatures
rfi	1.37270	298.15	Effects of the presence of ethylacetate or benzene on the densities and volumetric properties of mixture (styrene + N,N-dimethylformamide)
rfi	1.37000	298.15	Molecular interactions in (2,4,6-trimethyl-1,3,5-trioxane + n-alkyl acetates) at T=(298.15, 303.15, and 308.15) K
rfi	1.36977	298.15	Properties of ionic liquid HMIMPF6 with carbonates, ketones and alkyl acetates
rfi	1.35970	318.15	Thermodynamic properties of (an ester + an alkane). XVI. Experimental HEm and V Em values and a new correlation method for (an alkyl ethanoate + an n-alkane) at 318.15 K
rfi	1.37240	293.15	A novel static analytical apparatus for phase equilibrium measurements
rfi	1.35940	318.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents

rfi	1.36720		303.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents
rfi	1.37490		288.20	Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents
rfi	1.37020		293.20	Vapor liquid equilibria for the ternary system of carbon dioxide + ethanol + ethyl acetate at elevated pressures
rfi	1.35919		318.15	Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K
rhol	856.81	kg/m3	328.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	888.75	kg/m3	303.15	Excess Volumes and Excess Isentropic Compressibilities of Binary Liquid Mixtures of Trichloroethylene with Esters at 303.15 K
rhol	894.40	kg/m3	298.15	Revision of the volumetric method for measurements of liquid liquid equilibria in binary systems

rhol	894.60	kg/m3	298.15	Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa
rhol	906.55	kg/m3	288.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	894.38	kg/m3	298.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	875.81	kg/m3	313.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	901.00	kg/m3	293.00	KDB
rhol	897.72	kg/m3	298.15	Effect of Inorganic Salts on the Isobaric Vapor Liquid Equilibrium of the Ethyl Acetate Ethanol System

rhol	843.86	kg/m3	338.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	830.62	kg/m3	348.15	Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T= (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2 K) and p = 101.3 kPa
rhol	900.48	kg/m3	293.15	Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ ionic liquid - hydroxyacetone - water mixtures
rhol	894.36	kg/m3	298.15	Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ ionic liquid - hydroxyacetone - water mixtures
rhol	888.19	kg/m3	303.15	Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ ionic liquid - hydroxyacetone - water mixtures
rhol	881.98	kg/m3	308.15	Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ ionic liquid - hydroxyacetone - water mixtures

rhol	875.73	kg/m3	313.15	Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ ionic liquid - hydroxyacetone - water mixtures
rhol	900.57	kg/m3	293.20	Liquid-liquid equilibrium data for ternary systems of water + acetic acid + acetate esters at 293.2 K and 303.2 K and ~ 95 kPa
rhol	876.10	kg/m3	313.20	Liquid-liquid equilibrium data for ternary systems of water + acetic acid + acetate esters at 293.2 K and 303.2 K and ~ 95 kPa
rhol	894.51	kg/m3	298.15	Standard partial molar volumes of some electrolytes in ethylene carbonate based mixtures
rhol	900.26	kg/m3	293.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	894.14	kg/m3	298.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	875.10	kg/m3	313.15	Ternary Excess Molar Volumes of {Methyltrioctylammonium Bis[(trifluoromethyl)sulfonyl]imide + Methanol + Methyl Acetate or Ethyl Acetate} Systems at (298.15, 303.15, and 313.15) K

rhol	888.50	kg/m3	303.15	Ternary Excess Molar Volumes of {Methyltrioctylammonium Bis[(trifluoromethyl)sulfonyl]imide + Methanol + Methyl Acetate or Ethyl Acetate} Systems at (298.15, 303.15, and 313.15) K
rhol	894.70	kg/m3	298.15	Ternary Excess Molar Volumes of {Methyltrioctylammonium Bis[(trifluoromethyl)sulfonyl]imide + Methanol + Methyl Acetate or Ethyl Acetate} Systems at (298.15, 303.15, and 313.15) K
rhol	893.90	kg/m3	298.40	Vapor Liquid Equilibrium at 350 K, Excess Molar Enthalpies at 298 K, and Excess Molar Volumes at 298 K of Binary Mixtures Containing Ethyl Acetate, Butyl Acetate, and 2-Butanol
rhol	881.52	kg/m3	308.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	894.40	kg/m3	298.15	Liquid-Liquid Equilibrium in Ternary Systems Containing Ethylene Glycol, Monofunctional Benzene Derivative, and Ethyl Acetate
rhol	876.25	kg/m3	313.15	Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5)

rhol	882.48	kg/m3	308.15	Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5)
rhol	888.66	kg/m3	303.15	Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5)
rhol	894.79	kg/m3	298.15	Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5)
rhol	900.88	kg/m3	293.15	Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5)
rhol	899.50	kg/m3	298.15	Isobaric Vapor-liquid Equilibrium for Three Binary Systems of Ethyl Acetate + Propyl Acetate, Ethyl Acetate + Propylene Carbonate, and Propyl Acetate + Propylene Carbonate at 101.3 kPa
rhol	863.30	kg/m3	323.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K

rhol	869.62	kg/m3	318.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	875.88	kg/m3	313.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	882.09	kg/m3	308.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	888.26	kg/m3	303.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K

rhol	894.37	kg/m3	298.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	900.44	kg/m3	293.15	Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K
rhol	876.30	kg/m3	313.15	Apparent molar volume and apparent molar isentropic compressibility for thebinary systems {methyltrioctylammoniumbis(trifluoromethylsulfonyl)imide + ethyl acetate or ethanol} at different temperatures underatmospheric pressure
rhol	882.70	kg/m3	308.15	Apparent molar volume and apparent molar isentropic compressibility for thebinary systems {methyltrioctylammoniumbis(trifluoromethylsulfonyl)imide + ethyl acetate or ethanol} at different temperatures underatmospheric pressure

rhol	887.85	kg/m3	303.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K	
rhol	894.40	kg/m3	298.15	Apparent molar volume and apparent molar isentropic compressibility for thebinary systems {methyltrioctylammoniumbis(trifluoromethylsulfonyl)imide + ethyl acetate or ethanol} at different temperatures underatmospheric pressure	
rhol	888.50	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl ethanoate + cyclohexane at 303.15 K: Experimental data, correlation and prediction by ERAS model	
rhol	888.50	kg/m3	303.15	Volumetric and transport properties of ternary mixtures containing 1-propanol + ethyl ethanoate + cyclohexane or benzene at 303.15 K: Experimental data, correlation and prediction by ERAS model	
rhol	888.42	kg/m3	303.15	Studies of viscosities of dilute solutions of alkylamine in non-electrolyte solvents. II. Haloalkanes and other polar solvents	

rhol	894.40	kg/m3	298.15	Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K	
rhol	893.40	kg/m3	298.15	Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents	
rhol	894.50	kg/m3	298.15	(Liquid + liquid) equilibria for mixtures of dodecane and ethanol with alkylsulfate-based ionic liquids	
rhol	863.36	kg/m3	323.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	869.68	kg/m3	318.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	875.95	kg/m3	313.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	882.17	kg/m3	308.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	

rhol	888.35	kg/m3	303.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	894.47	kg/m3	298.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	900.57	kg/m3	293.15	Study on thermo physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates	
rhol	875.14	kg/m3	313.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K	
rhol	888.50	kg/m3	303.15	Apparent molar volume and apparent molar isentropic compressibility for thebinary systems {methyltrioctylammoniumbis(trifluoromethylsulfonyl)imide + ethyl acetate or ethanol} at different temperatures underatmospheric pressure	
rhol	851.36	kg/m3	332.70	Isothermal vapor liquid equilibria for different binary mixtures involved in the alcoholic distillation	
sfust	55.27	J/mol×K	189.30	NIST Webbook	

speedsl	1098.43	m/s	308.15	Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K
speedsl	1098.00	m/s	308.15	Density and Speed of Sound of Binary Mixtures of N-Methylacetamide with Ethyl Acetate, Ethyl Chloroacetate, and Ethyl Cyanoacetate in the Temperature Interval (303.15 to 318.15) K
speedsl	1081.00	m/s	313.15	Density and Speed of Sound of Binary Mixtures of N-Methylacetamide with Ethyl Acetate, Ethyl Chloroacetate, and Ethyl Cyanoacetate in the Temperature Interval (303.15 to 318.15) K
speedsl	1058.00	m/s	318.15	Density and Speed of Sound of Binary Mixtures of N-Methylacetamide with Ethyl Acetate, Ethyl Chloroacetate, and Ethyl Cyanoacetate in the Temperature Interval (303.15 to 318.15) K
speedsl	1120.80	m/s	303.15	Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K

speedsl	1119.00	m/s	303.15	Density and Speed of Sound of Binary Mixtures of N-Methylacetamide with Ethyl Acetate, Ethyl Chloroacetate, and Ethyl Cyanoacetate in the Temperature Interval (303.15 to 318.15) K	
speedsl	1143.23	m/s	298.15	Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K	
speedsl	1187.92	m/s	288.15	Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K	
speedsl	1165.80	m/s	293.15	Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K	
srf	0.03	N/m	323.20	KDB	
srf	0.02	N/m	298.15	Surface Tension Data of Aqueous Binary Mixtures of Methyl, Ethyl, Propyl, and Butyl Acetates at 298.15 K	

srf	0.02	N/m	298.15	Concentration Dependence of Surface Tension for Very Dilute Aqueous Solutions of Organic Non-Electrolytes	
tcondl	0.13	W/m×K	333.37	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.13	W/m×K	323.25	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.13	W/m×K	328.39	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.13	W/m×K	318.40	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.13	W/m×K	338.34	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.13	W/m×K	343.27	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.12	W/m×K	345.43	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.12	W/m×K	348.46	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	
tcondl	0.14	W/m×K	313.45	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase	

tcondl	0.14	W/m×K	303.53	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.14	W/m×K	298.52	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.14	W/m×K	295.44	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.15	W/m×K	290.42	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.15	W/m×K	285.31	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.15	W/m×K	281.34	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.15	W/m×K	274.35	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.15	W/m×K	268.04	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.16	W/m×K	260.17	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.16	W/m×K	253.89	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase

tcondl	0.16	W/m×K	249.13	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase
tcondl	0.14	W/m×K	308.44	Measurement of the thermal conductivity of five aliphatic esters in the liquid phase

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	332.82	K	54.99	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	335.16	K	59.98	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	339.42	K	70.00	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	340.92	K	75.00	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	343.18	K	80.00	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil

tbp	344.93	K	84.99	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	346.58	K	89.91	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	348.20	K	95.01	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	351.26	K	104.99	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	352.67	K	109.98	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	355.36	K	119.90	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	356.62	K	124.98	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	357.86	K	129.93	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil
tbp	359.08	K	134.95	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil

tbp	360.24	K	139.96	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	362.48	K	149.95	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	363.56	K	154.97	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	364.60	K	159.96	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	365.64	K	164.99	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	367.61	K	174.99	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	368.58	K	180.00	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	
tbp	369.50	K	185.00	Vapor Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil	

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)
Coeff. A	1.47020e+01
Coeff. B	-3.08216e+03
Coeff. C	-4.45420e+01
Temperature range (K), min.	258.37
Temperature range (K), max.	372.76

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	8.09498e+01
Coeff. B	-6.86243e+03
Coeff. C	-9.83731e+00
Coeff. D	7.27641e-06
Temperature range (K), min.	189.60
Temperature range (K), max.	523.30

Datasets

Mass density, kg/m3

Temperature, K - Liquid	Pressure, kPa - Liquid	Mass density, kg/m3 - Liquid
298.15	100.00	894.0
298.15	1000.00	894.7
298.15	2000.00	895.7
298.15	3000.00	896.7
298.15	3800.00	897.4
298.15	5000.00	898.8
298.15	10000.00	903.4
298.15	15000.00	907.7
298.15	20000.00	912.0

298.15	25000.00	916.1
298.15	30000.00	921.1
298.15	35000.00	926.3
303.15	100.00	887.6
303.15	1000.00	888.5
303.15	2000.00	889.5
303.15	3000.00	890.5
303.15	3800.00	891.3
303.15	5000.00	892.7
303.15	10000.00	897.5
303.15	15000.00	902.0
303.15	20000.00	906.3
303.15	25000.00	910.7
303.15	30000.00	915.9
303.15	35000.00	921.2
308.15	100.00	881.4
308.15	1000.00	882.3
308.15	2000.00	883.3
308.15	3000.00	884.4
308.15	3800.00	885.3
308.15	5000.00	886.7
308.15	10000.00	891.7
308.15	15000.00	896.4
308.15	20000.00	900.9
308.15	25000.00	905.4
308.15	30000.00	910.7
308.15	35000.00	916.2
313.15	100.00	875.2
313.15	1000.00	876.1
313.15	2000.00	877.2
313.15	3000.00	878.3
313.15	3800.00	879.2
313.15	5000.00	880.6
313.15	10000.00	885.8
313.15	15000.00	890.7
313.15	20000.00	895.4
313.15	25000.00	899.8
313.15	30000.00	905.3
313.15	35000.00	910.9
318.15	100.00	868.8
318.15	1000.00	869.8
318.15	2000.00	871.0
318.15	3000.00	872.2
318.15	3800.00	873.1

318.15	5000.00	874.5
318.15	10000.00	880.0
318.15	15000.00	885.1
318.15	20000.00	889.8
318.15	25000.00	894.5
318.15	30000.00	900.1
318.15	35000.00	905.9
323.15	100.00	862.7
323.15	1000.00	864.0
323.15	2000.00	865.1
323.15	3000.00	866.2
323.15	3800.00	867.1
323.15	5000.00	868.6
323.15	10000.00	874.2
323.15	15000.00	879.4
323.15	20000.00	884.3
323.15	25000.00	889.2
323.15	30000.00	895.0
323.15	35000.00	900.9
328.15	100.00	856.4
328.15	1000.00	857.8
328.15	2000.00	859.2
328.15	3000.00	860.5
328.15	3800.00	861.5
328.15	5000.00	862.5
328.15	10000.00	868.4
328.15	15000.00	873.9
328.15	20000.00	879.0
328.15	25000.00	884.1
328.15	30000.00	890.1
328.15	35000.00	896.2
333.15	100.00	850.4
333.15	1000.00	851.7
333.15	2000.00	853.0
333.15	3000.00	854.4
333.15	3800.00	855.4
333.15	5000.00	856.8
333.15	10000.00	862.9
333.15	15000.00	868.5
333.15	20000.00	873.8
333.15	25000.00	879.0
333.15	30000.00	885.2
333.15	35000.00	891.4
338.15	100.00	844.0

338.15	1000.00	845.4
338.15	2000.00	846.8
338.15	3000.00	848.2
338.15	3800.00	849.3
338.15	5000.00	850.8
338.15	10000.00	857.1
338.15	15000.00	863.0
338.15	20000.00	868.5
338.15	25000.00	873.9
338.15	30000.00	880.3
338.15	35000.00	886.7
343.15	100.00	837.5
343.15	1000.00	839.1
343.15	2000.00	840.6
343.15	3000.00	842.0
343.15	3800.00	843.2
343.15	5000.00	844.5
343.15	10000.00	851.2
343.15	15000.00	857.3
343.15	20000.00	863.0
343.15	25000.00	868.5
343.15	30000.00	875.1
343.15	35000.00	881.6
348.15	100.00	829.4
348.15	1000.00	830.5
348.15	2000.00	832.3
348.15	3000.00	833.8
348.15	3800.00	835.0
348.15	5000.00	836.8
348.15	10000.00	845.5
348.15	15000.00	851.8
348.15	20000.00	857.7
348.15	25000.00	863.5
348.15	30000.00	870.2
348.15	35000.00	877.0
353.15	1000.00	824.3
353.15	2000.00	826.0
353.15	3000.00	827.5
353.15	3800.00	828.8
353.15	5000.00	830.6
353.15	10000.00	837.7
353.15	15000.00	844.3
353.15	20000.00	850.4
353.15	25000.00	856.1

353.15	30000.00	862.7
353.15	35000.00	869.6
358.15	1000.00	818.4
358.15	2000.00	820.0
358.15	3000.00	821.6
358.15	3800.00	822.8
358.15	5000.00	824.7
358.15	10000.00	832.1
358.15	15000.00	838.9
358.15	20000.00	845.2
358.15	25000.00	851.0
358.15	30000.00	857.6
358.15	35000.00	864.7
363.15	1000.00	811.8
363.15	2000.00	813.5
363.15	3000.00	815.3
363.15	3800.00	816.6
363.15	5000.00	818.5
363.15	10000.00	826.3
363.15	15000.00	833.3
363.15	20000.00	839.7
363.15	25000.00	845.8
363.15	30000.00	852.9
363.15	35000.00	860.2
368.15	1000.00	804.9
368.15	2000.00	806.8
368.15	3000.00	808.6
368.15	3800.00	810.0
368.15	5000.00	812.1
368.15	10000.00	820.2
368.15	15000.00	827.5
368.15	20000.00	834.3
368.15	25000.00	840.6
368.15	30000.00	848.0
368.15	35000.00	855.5
373.15	1000.00	798.5
373.15	2000.00	800.4
373.15	3000.00	802.4
373.15	3800.00	803.8
373.15	5000.00	806.0
373.15	10000.00	814.4
373.15	15000.00	822.0
373.15	20000.00	829.0
373.15	25000.00	835.5

373.15	30000.00	843.0
373.15	35000.00	850.6
378.15	1000.00	791.8
378.15	2000.00	793.8
378.15	3000.00	795.8
378.15	3800.00	797.3
378.15	5000.00	799.6
378.15	10000.00	808.5
378.15	15000.00	816.5
378.15	20000.00	823.7
378.15	25000.00	830.5
378.15	30000.00	838.4
378.15	35000.00	846.3
383.15	1000.00	784.8
383.15	2000.00	787.1
383.15	3000.00	789.1
383.15	3800.00	790.8
383.15	5000.00	793.2
383.15	10000.00	802.4
383.15	15000.00	810.7
383.15	20000.00	818.2
383.15	25000.00	825.3
383.15	30000.00	833.5
383.15	35000.00	841.6
388.15	1000.00	778.2
388.15	2000.00	780.5
388.15	3000.00	782.7
388.15	3800.00	784.4
388.15	5000.00	786.9
388.15	10000.00	796.4
388.15	15000.00	805.1
388.15	20000.00	813.0
388.15	25000.00	820.3
388.15	30000.00	828.8
388.15	35000.00	837.2
393.15	1000.00	771.0
393.15	2000.00	773.5
393.15	3000.00	775.9
393.15	3800.00	777.7
393.15	5000.00	780.4
393.15	10000.00	790.5
393.15	15000.00	799.4
393.15	20000.00	807.6
393.15	25000.00	815.2

393.15	30000.00	823.9
393.15	35000.00	832.4

Reference https://www.doi.org/10.1021/je060339m

Sources

Influence of Different Solvent Properties and Composition for the Solubility of Iopromide:	https://www.doi.org/10.1021/acs.jced.8b00416
Improving the solubility of temozolomide by cosolvent and its correlation with the Jouyban-Acree and CNIBS/R-K models:	https://www.doi.org/10.1016/j.jct.2019.07.017
Solubility of Physalin D in Ethanol, Methanol, Propanone, Trichloromethane, Ethyl Ethanoate, and Water at Temperatures from (283.2 to 313.2) K:	https://www.doi.org/10.1021/je100227k
Liquid-liquid equilibrium data for (n-hexane + ethyl acetate + acetonitrile) ternary system at (298.15, 308.15, and 318.15) K:	https://www.doi.org/10.1016/j.fluid.2016.03.008
Solubilities of 3-Pentadecylphenol in Ethanol, 1-Butanol, Toluene, Acetone, Tetrachloromethane, and Ethyl Acetate:	https://www.doi.org/10.1021/je900346t
Liquid-liquid equilibria and COSMO-SAC modeling of organic solvent/ionic liquid - hydroxyacetone - water mixtures:	https://www.doi.org/10.1016/j.fluid.2018.01.026
Solubility and Thermodynamic Analysis of Tenoxicam in Different Pure Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.5b00382
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.07.004
Thermodynamics of fenofibrate and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.029
Investigation of Surface and Solubility Properties of N-Vinylimidazolium Tetrahalogenidoferrate(III) Magnetic Ionic Liquids Using Density Functional Theory:	https://www.doi.org/10.1021/acs.jced.5b00395
The Research and Measurement about the Solubility of L-Serine in Eight Common Pure Solvents and Four Binary Mixed Solvents for T = (278.15–333.15) K:	https://www.doi.org/10.1021/acs.jced.9b00460
Solubility Modeling and Mixing Thermodynamics of Thiamphenicol in Twelve Neat Organic Solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.7b00542
Solubility of 1,3,2-dioxaphosphorinane-5,5-dimethyl-2-(2,4,6-tribromophenoxy) 2-oxide in selected solvents:	https://www.doi.org/10.1016/j.fluid.2013.09.023
Dissolution behavior and thermodynamic properties of lapatinib ditosylate in pure and mixed organic solvents from T = (283.15–323.15) K:	https://www.doi.org/10.1016/j.fluid.2019.01.005
Activity Coefficients at Infinite Dilution of Volatile Compounds in Water: Effect of Temperature and Salt Concentration:	https://www.doi.org/10.1021/je2013878
Solubility Measurement and Thermodynamic Modeling for o-Toluenesulfonamide in 16 Solvents from T = 273.15 to 323.85 K:	https://www.doi.org/10.1021/acs.jced.9b00445
Measurements and Thermodynamic Models for the Separation of 2-Methoxyethanol and Water Azeotropic Mixtures via Different Solvents:	https://www.doi.org/10.1021/acs.jced.7b00160
The potential of head space gas chromatography for VLE measurements:	https://www.doi.org/10.1016/j.jct.2012.01.020
Solubility and Solution Thermodynamics of 3-Nitrophthalonitrile in 12 Neat Solvents at Temperatures from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b01250
Liquid–Liquid Equilibrium of (Water + Pentane-2,4-dione + Ethyl Ethanoate) and (Water + Pentane-2,4-dione + Cyclohexane) at (298.15 and 313.15) K:	https://www.doi.org/10.1021/je100704j
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Solubility and solution thermodynamics of 2-methyl-6-nitroaniline in ten organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.05.027
Isobaric Vapor-Liquid Equilibria and Excess Quantities for Binary Mixtures of an Ethyl Ester + tert-Butanol and a New Approach to VLE Data Processing:	https://www.doi.org/10.1021/je0202073
Measurement and correlation of the solubility of N-(phosphonomethyl)iminodiacetic acid in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.08.036
Equilibrium solubility of dinitolmide in several neat solvents and binary aqueous co-solvent mixtures: Experimental determination and thermodynamic analysis:	https://www.doi.org/10.1016/j.jct.2019.01.013
Determination and Correlation of Solubilities of 2-Isopropylthioxanthone in Different Solvents from (299.15 to 329.85) K:	https://www.doi.org/10.1021/je501011t
Experimental Measurement and Correlation of Solubility Data and Thermodynamic Properties of Protocatechuic Acid in Four Organic Solvents:	https://www.doi.org/10.1021/je500519y
Activity coefficients at infinite dilution of organic solvents and water in 1-ethyl-3-methylimidazolium dicyanamide. A literature review of hexane/hex-1-ene separation:	https://www.doi.org/10.1016/j.fluid.2016.02.004
Experimental VLE data of Methyl Acetate or Ethyl Acetate with 2-Butanol at 0.3 and 0.6 MPa-Quality assessment and predictions:	https://www.doi.org/10.1021/je5003225
Volumetric and FT-IR Studies of the Binary Liquid Mixtures of Tributylamine and Alkyl Ester (C1-C5):	https://www.doi.org/10.1021/acs.jced.8b01120
Volumetric and Transport Properties of Binary Liquid Mixtures of Phenylacetonitrile with Aliphatic Esters at Temperatures of (303.15 to 313.15) K:	https://www.doi.org/10.1021/je900525f
(Liquid + liquid) equilibria for mixtures of dodecane and ethanol with alkylsulfate-based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.09.009
Determination of the thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide by gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Solubility of N-phenylanthranilic acid in nine organic solvents from T = (283.15 to 318.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.08.017

Thermodynamic models for determination of the solubility of omeprazole in pure and mixture organic solvents from T = (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.jct.2015.11.005
Solubilities of 4,4'-Dichlorodiphenyl Disulfide in Six Organic Solvents between (303.15 and 333.15) K:	https://www.doi.org/10.1021/je301243f
Solid-liquid equilibrium of diphenyl anilinophosphonate in the different organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.036
Separation of binary mixtures based on gamma infinity data using [OMMIM][NTf2] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2017.12.012
Determination and Thermodynamic Modeling of Solid-Liquid Phase Equilibrium for Succinic Acid in the Glutaric Acid + Adipic Acid + Ethyl Acetate Mixture and Adipic Acid in the Succinic Acid + Glutaric Acid + Ethyl Acetate Mixture:	https://www.doi.org/10.1021/acs.jced.8b01127
solubility of L-(+)-ascorbic acid in water, ethanol, methanol, propan-2-ol, acetone, acetonitrile, ethyl acetate, and tetrahydrofuran from (293 to 323) K:	https://www.doi.org/10.1021/je800056h
Solubility Modeling and Mixing Properties for Benzoin in Different Monosolvents and Solvent Mixtures at the Temperature Range from 273.15 to 313.15 K:	https://www.doi.org/10.1021/acs.jced.7b00743
Separation of water/butan-1-ol based on activity coefficients at infinite dilution in 1,3-didecyl-2-methylimidazolium dicyanamide ionic liquid:	https://www.doi.org/10.1016/j.jct.2017.10.003
Apparent molar volume and apparent molar isentropic compressibility for binary systems {methyltrioctylammonium bis(trifluoromethylsulfonyl)imide + methyl acetate or ethanol} at different temperatures under atmospheric pressure:	https://www.doi.org/10.1016/j.tca.2013.05.015
Solubility of itaconic acid in different organic solvents: Experimental measurement and thermodynamic modeling:	https://www.doi.org/10.1016/j.fluid.2011.09.027
Solubilities of Ursolic Acid and Oleanolic Acid in Four Solvents from (283.2 to 329.7) K:	https://www.doi.org/10.1021/je101309a
Density, Speed of Sound, and Refractive Index of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate with Acetone, Methyl Acetate, and Ethyl Acetate at Temperatures from (278.15 to 328.15) K:	https://www.doi.org/10.1021/je9005618
Determination and Correlation of Solubilities of 1,3,5-Trifluoro-2,4,6-triiodobenzene in Different Solvents from 289.25 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00381
Measurement of the thermal conductivity of five aliphatic esters in the liquid phase:	https://www.doi.org/10.1016/j.jct.2019.06.014
Solubility of clopidogrel hydrogen sulfate polymorphs in ethyl acetate + 2-butanol mixtures at 283.15-313.15 K:	https://www.doi.org/10.1016/j.jct.2019.06.021
Solubility of Genistein in Water, Methanol, Ethanol, Propan-2-ol, 1-Butanol, and Ethyl Acetate from (280 to 333) K:	https://www.doi.org/10.1021/je100261w
Liquid-Liquid Equilibria of Water + 2-Butanol + (Methyl Methacrylate or Butyl Methacrylate or Isobutyl Methacrylate) at (288.2 and 318.2) K:	https://www.doi.org/10.1021/je7002572
Solubility Study and Mixing Property of 3,5-Dinitro-2-methylbenzoic Acid in 13 Pure Solvents from 288.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00497
Solubilities of 1,3,2-Dioxaphosphorinane-2-methanol-a,a-5,5-tetramethyl 2-oxide in Selected Solvents at 278.15 K to 343.15 K:	https://www.doi.org/10.1021/je4003114
Solubility Determination and Modeling for 4-Nitrobenzonitrile in Binary Solvent Mixtures of Ethyl Acetate Plus (Methanol, Ethanol, n-Propanol, and Isopropanol):	https://www.doi.org/10.1021/acs.jced.8b00555
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2013.01.007
Solubility of 1-H-Tetrazole-1-acetic Acid in Different Solvents between 283 K and 323 K:	https://www.doi.org/10.1021/je700200b
Solubility of Benzoin in Six Monosolvents and in Some Binary Solvent Mixtures at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00238
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je900704b
Solubility Measurement and Thermodynamic Modeling of N-(4-Methylphenyl)-2-3-chloro-2-(phenylthio)propenamide in 12 Pure Solvents at Temperatures Ranging from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b01011
Solubility of Diflufenican in Pure and Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b00632
Solubility of p-Methylbenzene Sulfonic Acid in Pure and Modified Supercritical Carbon Dioxide:	https://www.doi.org/10.1021/je700639s
Tetramethylammonium chloride + glycerol deep eutectic solvent as separation agent for organic liquid mixtures: Assessment from experimental limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.06.003
Interactions of Volatile Organic Compounds with the Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide:	https://www.doi.org/10.1021/je200822w
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in 1-Ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide ionic liquid:	https://www.doi.org/10.1016/j.jct.2016.06.028
Experimental Determination of Vapor Liquid Equilibria, Binary Systems of Methyl Acetate, Ethyl Acetate, and Propyl Acetate with 1-Propanol at 0.6 MPa:	https://www.doi.org/10.1021/je400622g
Solubility of Dexibuprofen in Different Solvents from (263.15 to 293.15) K:	https://www.doi.org/10.1021/je101165m
Effect of Inorganic Salts on the Isobaric Vapor Liquid Equilibrium of the Ethyl Acetate Ethanol System:	https://www.doi.org/10.1021/je300886k
Measurement and Correlation of Solubility and Thermodynamic Properties of Vinpocetine in Nine Pure Solvents and (Ethanol + Water) Binary Solvent:	https://www.doi.org/10.1021/acs.jced.8b00663
Concentration Dependence of Surface Tension for Very Dilute Aqueous Solutions of Organic Non-Electrolytes:	https://www.doi.org/10.1021/je049955d
Solubility of Acetoguanamine in Twelve Neat Solvents from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00593
Solubility of 2,5-Di-tert-butylhydroquinone and Process Design for Its Purification Using Crystallization:	https://www.doi.org/10.1021/je501049c
Solubility of 2-nitro-p-phenylenediamine in nine pure solvents and mixture of (methanol + N-methyl-2-pyrrolidone) from T = (283.15 to 318.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2017.01.006
Vapor-Liquid Equilibrium of Binary Mixtures Containing Ethyl Acetate + 2-Methyl-1-propanol and Ethyl Acetate + 3-Methyl-1-butanol at 101.3 kPa:	https://www.doi.org/10.1021/je050420+
Solution thermodynamics of simvastatin in pure solvents and binary solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2015.07.055
Thermodynamic Properties of Ionic Liquids in Organic Solvents from (293.15 to 303.15) K:	https://www.doi.org/10.1021/je060497d
Dependence of Refractive Index on Concentration and Temperature in Electrolyte Solution, Polar Solution, Nonpolar Solution, and Protein Solution:	https://www.doi.org/10.1021/acs.jced.5b00018
Solubility of Artemisinin in Different Single and Binary Solvent Mixtures Between (284.15 and 323.15) K and NRTL Interaction Parameters:	https://www.doi.org/10.1021/je100125x

Solubility and Solvent Effect of Acetamiprid in Thirteen Pure Solvents and Aqueous Solutions of Ethanol:	https://www.doi.org/10.1021/acs.jced.9b00294
Solid-Liquid Equilibrium of 5-Chloro-8-hydroxyquinoline and 5,7-Dichloro-8-hydroxyquinoline in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.9b00028
Low pressure carbon dioxide solubility in pure electrolyte solvents for lithium-ion batteries as a function of temperature. Measurement and prediction:	https://www.doi.org/10.1016/j.jct.2012.01.027
Measurement of Activity Coefficients of Solutes at Infinite Dilution in (Dimethyl Sulfoxide + Acetamide, or Formamide, or Urea) Using Gas Liquid Chromatography at the Temperature 298.15 K:	https://www.doi.org/10.1021/je030122h
Viscosity and Density of Water + Ethyl Acetate + Ethanol Mixtures at 298.15 and 318.15 K and Atmospheric Pressure:	https://www.doi.org/10.1021/je600565m
Abraham model correlations for describing the thermodynamic properties of solute transfer into pentyl acetate based on headspace chromatographic and solubility measurements:	https://www.doi.org/10.1016/j.jct.2018.05.003
Solubility of 6-Chloropyridazin-3-amine in Different Solvents:	https://www.doi.org/10.1021/je3001188
Measurement and Correlation of Solubility and Dissolution Thermodynamic Properties of Furan-2-carboxylic Acid in Pure and Binary Solvents:	https://www.doi.org/10.1021/je500012b
Solubility Measurement and Correlation of (+)-Biotin Intermediate Lactone in Different Organic Solvents from (287.15 to 323.75) K:	https://www.doi.org/10.1021/acs.jced.5b00857
Solubilities of Tris(o-phenylenedioxy)cyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je200252c
Solubility determination and thermodynamic functions of 2-chlorophenothiazine in nine organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.029
Solubility determination and thermodynamic modeling of paclobutrazol in nine organic solvents from (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.038
Liquid-Liquid Equilibrium in Ternary Systems Containing Ethylene Glycol, Monofunctional Benzene Derivative, and Ethyl Acetate:	https://www.doi.org/10.1021/je100921h
Determination and correlation of solubility and solution thermodynamics of musk xylene in different pure solvents:	https://www.doi.org/10.1016/j.jct.2019.03.029
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.03.015
Activity coefficients at infinite dilution of organic solutes in the ionic liquid [illegible]	https://www.doi.org/10.1016/j.jct.2014.04.020
Thermodynamic modelling for solubility of 5-chloro-1-methyl-4-nitroimidazole in eleven organic solvents from T = (283.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.10.006
Thermodynamic analysis for solid-liquid equilibrium of mercaptosuccinic acid in binary solvent of methanol and ethyl acetate at temperatures between (278.15 and 333.15) K:	https://www.doi.org/10.1016/j.fluid.2015.02.008
Measurement and Correlation of the Solubility of Febuxostat in Four Organic Solvents at Various Temperatures:	https://www.doi.org/10.1021/je300647k
Vapour liquid equilibrium for the ethyl ethanoate + 1-butene, +cis-2-butene, +trans-2-butene, +2-methylpropene, +n-butane and +2-methylpropane:	https://www.doi.org/10.1016/j.fluid.2004.10.023
Solubilities of rutin in eight solvents at T = 283.15, 298.15, 313.15, 323.15, and 333.15 K:	https://www.doi.org/10.1016/j.fluid.2007.07.030
Solubility and Thermodynamic Behavior of Veramoss in Different Pure Solvents and {Ethanol + Water} Mixtures from T = 283.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.6b00646
Solubility of methylphosphonic acid in selected organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.01.009
Densities and Speeds of Sound of Binary Liquid Mixtures of Some n-Alkoxypropanols with Methyl Acetate, Ethyl Acetate, and n-Butyl Acetate at T = (288.15, 293.15, 298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1021/je300789a
Solubility of Imperatorin in Ethyl Acetate, Ethanol, Methanol, n-Hexane, and Petroleum Ether from (288.2 to 328.2) K:	https://www.doi.org/10.1021/je101303e
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2013.05.008
Solubility of 2-Chlorophenylacetic Acid in 12 Pure Solvents from T = (273.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.8b00139
Measurements of the Excess Properties and Vapor-Liquid Equilibria at 101.32 kPa for Mixtures of Ethyl Ethanoate + Alkanes (from C5 to C10):	https://www.doi.org/10.1021/je100832h
Solubility, correlation and thermodynamic analysis of two forms of Metaxalone in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2015.09.013
Measurement and Correlation of the Solubility of Dienogest in Twelve Pure and Water + Methanol Binary Solvents at Temperatures from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00696
Measurement and Correlation for the Solubility of Dimethyl 1,4-Cyclohexanedione-2,5-dicarboxylate in Different Solvents at Temperatures from (278.15 to 323.15) K:	https://www.doi.org/10.1021/je2000292
Total Vapor Pressure Measurements for 2-Ethoxyethanol with Methyl Acetate, Ethyl Acetate, Propyl Acetate, and Ethyl Propionate at 313.15 K and for 2-Ethoxyethanol with Methyl Formate at 308.15 K:	https://www.doi.org/10.1021/je0255628
Experimental Measurements and Modeling of the Solubility of Aceclofenac in Six Pure Solvents from (293.35 to 338.25) K:	https://www.doi.org/10.1021/je500038u
Solubility Determination and Modeling for 4,4'-Dihydroxydiphenyl Sulfone in Mixed Solvents of (Acetone, Ethyl Acetate, or Acetonitrile) + Methanol and Acetone + Ethanol from (278.15 to 313.15) K:	https://www.doi.org/10.1021/acs.jced.6b00421
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylmorpholinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.08.030
Determination and Correlation of the Solubility of Musk Ketone in Pure and Binary Solvents at 273.15–313.15 K:	https://www.doi.org/10.1021/acs.jced.9b00310
Measurement and correlation of solubility of dodecanedioic acid in different pure solvents from T = (288.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2013.09.012
Experimental solubility and thermodynamic aspects of methylene blue in different solvents:	https://www.doi.org/10.1016/j.tca.2019.03.024
Density, dynamic viscosity, and derived properties of binary mixtures of methanol or ethanol with water, ethyl acetate, and methyl acetate at T = (293.15, 298.15, and 303.15) K:	https://www.doi.org/10.1016/j.jct.2007.05.004
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je1009812
Solid-Liquid Phase Equilibrium and Solubility of Dibenzo[b,d]furan and 9H-Fluorene in Organic Solvents:	https://www.doi.org/10.1021/je201282d
Revision of the volumetric method for measurements of liquid liquid equilibria in binary systems:	https://www.doi.org/10.1016/j.fluid.2004.12.004
Solubility and Data Correlation of Isoniazid in Different Pure and Binary Mixed Solvent Systems from 283.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00785

Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Trifluoromethanesulfonate Using Gas Liquid Chromatography	https://www.doi.org/10.1021/je700560s
Solubility measurement and thermodynamic functions of [illegible] acetate in different solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.01.023
Solubility of Podophyllotoxin in Six Organic Solvents from (283.2 to 323.2)	https://www.doi.org/10.1021/je8007028
Thermostabilities and solubilities of flame retardant bisphenol S-bis (5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy phosphate ester) as a function of temperature:	https://www.doi.org/10.1016/j.fluid.2014.04.006
Determination and correlation of molar excess enthalpies of binary systems 2,4-pentanedione + (1-butanol, + 2-methyl-1-propanol, + 1-pentanol, + 1-heptane, + ethyl acetate, and + water)	https://www.doi.org/10.1016/j.fluid.2007.12.008
Solubility and Dissolution Thermodynamic Data of Mefenamic Acid Crystal Forms in Different Classes of Organic Solvents:	https://www.doi.org/10.1021/je400714f
Solubilities of [illegible] acid in different organic solvents at several temperatures from (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.fluid.2013.06.038
Vapor-Liquid Equilibrium for Binary Mixtures of Acetates in the Direct Esterification of Fusel Oil:	https://www.doi.org/10.1021/acs.jced.6b00221
Solubility of Amorphous Clopidogrel Hydrogen Sulfate in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.5b00306
Solubility of 4-Methylsulfonylacetophenone in Nine Pure Solvents and (Ethyl Acetate + n-Hexane) Binary Solvent Mixtures from 278.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00617
Vapor-liquid equilibrium data for the binary systems in the process of synthesizing diethyl carbonate:	https://www.doi.org/10.1016/j.fluid.2009.12.011
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00490
Solubility and thermodynamic properties of dirithromycin form A and form B in pure solvents and binary solvent mixture:	https://www.doi.org/10.1016/j.jct.2018.12.044
Measurement and Correlation for the Solubilities of Dihydrocapsaicin in n-Heptane, n-Hexane, n-Pentane, Ethyl Acetate, Acetone, Ethanol, and Water:	https://www.doi.org/10.1021/je1011136
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
Solubility of 2-Chloro-4,6-dinitroresorcinol in Ethanol, Methanol, Acetic Acid, Ethyl Acetate, and Water:	https://www.doi.org/10.1021/je1011423
Measurement and correlation for solubility of levofloxacin in six solvents at temperatures from 288.15 to 328.15 K:	https://www.doi.org/10.1016/j.fluid.2012.05.027
Thermodynamic study of molecular interaction-selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Determination and Thermodynamic Modeling of Solid Liquid Phase Equilibrium for Esomeprazole Sodium in Monosolvents and in the (Ethanol + Ethyl Acetate) Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b01044
Determination and correlation thermodynamic models for solid liquid equilibrium of the Nifedipine in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.jct.2016.07.035
Determination of the solubility, dissolution enthalpy and entropy of suberic acid in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.012
Activity coefficients at infinite dilution for organic solutes and water in 1-ethyl-1-methylpyrrolidinium lactate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Solubility determination and thermodynamic modeling of 2,4-dinitroaniline in nine organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.009
Vapor-Liquid Equilibrium of Carbon Dioxide + Ethyl Acetate + Oleic Acid Mixtures at High Pressures:	https://www.doi.org/10.1021/acs.jced.7b00163
Measurement and Correlation of Solubility of 2-Chloro-4-ethylamino-6-isopropylamino-1,3,5-triazine in Different Organic Solvents:	https://www.doi.org/10.1021/je400639m
Solubility Measurement and Thermodynamic Model Correlation and Evaluation of 2-Chloro-5-nitroaniline in 12 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00931
Enhancement of liquid phase splitting of water + ethanol + ethyl acetate mixtures in the presence of a hydrophilic agent or an electrolyte substance:	https://www.doi.org/10.1016/j.fluid.2005.08.009
Solubilities of cinnamic acid esters in binary mixtures of ionic liquids and organic solvents:	https://www.doi.org/10.1016/j.fluid.2010.05.015
Isobaric vapor-liquid equilibrium of a ternary system of ethyl acetate + propyl acetate + dimethyl sulfoxide and binary systems of ethyl acetate + dimethyl sulfoxide and propyl acetate + dimethyl sulfoxide at 101.3 kPa:	https://www.doi.org/10.1016/j.jct.2019.03.036
Density, Viscosity, Refractive Index, and Speed of Sound for Binary Mixtures of 1,4-Dioxane with Different Organic Liquids at (298.15, 303.15, and 308.15) K.	https://www.doi.org/10.1021/je049609w
Solubility of Hydroquinone in Different Solvents from 276.65 K to 345.10 K:	https://www.doi.org/10.1021/je0502748
Solubility Measurement and Modeling for 2-Benzoyl-3-chlorobenzoic Acid and 1-Chloroanthraquinone in Organic Solvents:	https://www.doi.org/10.1021/je400246s
Study of Ether-, Alcohol-, or Cyano-Functionalized Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je900838a
Molar excess enthalpies of ethyl acetate + alkanols at T=298.15K, p=10MPa	https://www.doi.org/10.1016/j.tca.2005.03.009
Thermodynamic Functions for the Solubility of 3-Nitrobenzonitrile in 12 Organic Solvents from T/K = (278.15 to 318.15):	https://www.doi.org/10.1021/acs.jced.7b00615
Measurement and correlation of the solubility of maleic anhydride in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.008
Liquid-liquid equilibrium study for ternary systems of water + propargyl alcohol + solvents at 308.2 K: Measurement and thermodynamic modelling:	https://www.doi.org/10.1016/j.jct.2019.03.033
Vapor-Liquid Equilibrium Data for Binary Mixtures of Dimethyl Carbonate with Methyl Acetate, Ethyl Acetate, n-Propyl Acetate, Isopropyl Acetate, n-Butyl Acetate, and Isoamyl Acetate at 93.13 kPa:	https://www.doi.org/10.1021/acs.jced.7b00704
Activity Coefficients at Infinite Dilution in 1-Alkyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids	https://www.doi.org/10.1021/je050544m
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid trihexyl-tetradecyl-phosphonium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2018.01.003
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2017.03.004
Solubilities of Bis(2,4,6-trimethylbenzoyl)phenylphosphine Oxide in Different Organic Solvents at Several Temperatures:	https://www.doi.org/10.1021/je300343y
Solubility of diosgenin in different solvents:	https://www.doi.org/10.1016/j.jct.2011.11.009

Solubility modelling, solution thermodynamics and preferential solvation for nitroxoline in solvent mixtures of ethyl acetate + (methanol, ethanol, n-propanol and isopropanol):	https://www.doi.org/10.1016/j.jct.2017.05.037
Liquid liquid equilibria of 1,3-dimethylimidazolium methyl sulfate with ketones, dialkyl carbonates and acetates:	https://www.doi.org/10.1016/j.fluid.2007.02.027
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in the ionic liquid 1-dodecyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Solubilities of Triphenylphosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/je800842z
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-hexyl-2,3-dimethylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.07.021
Solubilities of phosphoric acid, P,P0-1,4-phenylene P,P,P0,P0-tetraphenyl ester in the organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.12.034
Solubility of 3-Chlorophthalic Anhydride and 4-Chlorophthalic Anhydride in Organic Solvents and Solubility of 3-Chlorophthalic Acid and 4-Chlorophthalic Acid in Water from (283.15 to 333.15) K:	https://www.doi.org/10.1021/je800869g
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Solubilities of succinimide in different pure solvents and binary methanol+ethyl acetate solvent mixtures:	https://www.doi.org/10.1016/j.tca.2012.03.007
Measurements for the solubilities of m-phthalic acid in acetone, ethanol and acetic ether:	https://www.doi.org/10.1016/j.fluid.2008.01.014
Solvation parameter model and thermodynamic parameters in a dicationic ionic liquid based on pyrrolidinium:	https://www.doi.org/10.1016/j.jct.2014.12.023
Thermodynamic analysis and enthalpy-entropy compensation for the solubility of indomethacin in aqueous and non-aqueous mixtures:	https://www.doi.org/10.1016/j.fluid.2011.06.016
Thermodynamic study of (alkyl esters + a,x-alkyl dihalides) II: HE m and V E m for 25 binary mixtures {xCu-1H2u-1CO2C2H5 + (1-x)a,x-BrCH2(CH2)v-2CH2Br}, where u = 1 to 3, and v = 2 to 6:	https://www.doi.org/10.1016/j.jct.2005.07.009
Vapor Liquid Equilibrium at 350 K, Excess Molar Enthalpies at 298 K, and Excess Molar Volumes at 298 K of Binary Mixtures Containing Ethyl Acetate, Butyl Acetate, and 2-Butanol:	https://www.doi.org/10.1021/je400036b
Solubilities of cyadiamine in twelve organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2012.11.021
Thermodynamic study of solubility for 2-amino-4-chloro-6-methoxypyrimidine in twelve organic solvents at temperatures from 273.15 K to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.10.020
Isothermal (vapour + liquid) equilibrium (VLE) for binary mixtures containing diethyl carbonate, phenyl acetate, diphenyl carbonate, or ethyl acetate:	https://www.doi.org/10.1016/j.jct.2015.07.024
Effect of Ionic Liquids on the Binary Vapor Liquid Equilibrium of Ethyl Acetate + Methanol System at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.8b00424
o-Nitrophenylacetonitrile Solubility in Several Pure Solvents: Measurement, Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00243
Solubility measurement and correlation of the form A of ibrutinib in organic solvents from 278.15 to 323.15 K:	https://www.doi.org/10.1016/j.jct.2016.08.034
Experiment and Application of the Solid-Liquid Phase Equilibrium for Ternary 2-Nitro-p-phenylenediamine + Ethyl Acetate + Isopropanol and 2-Nitro-p-phenylenediamine + 4-Nitro-o-phenylenediamine + N-Methyl-2-pyrrolidone Systems:	https://www.doi.org/10.1021/acs.jced.8b00952
Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je200531k
Solubility and mixing thermodynamic properties of (2,4,6-trimethylbenzoyl) diphenylphosphine oxide in pure and binary solvents:	https://www.doi.org/10.1016/j.fluid.2018.01.015
Solubility Determination and Thermodynamic Modeling of Buprofezin in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01099
Experimental determination and calculation of the critical curves for the binary systems of CO2 containing ketone, alkane, ester and alcohol, respectively:	https://www.doi.org/10.1016/j.fluid.2011.05.005
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds - Part II: C18 fatty acids:	https://www.doi.org/10.1016/j.jct.2012.12.009
Phase Equilibria of the Azeotropic Mixture Hexane + Ethyl Acetate with Ionic Liquids at 298.15 K:	https://www.doi.org/10.1021/je800077r
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.08.016
Measurement, Correlation, and Thermodynamics of Solubility of Isatin in Nine Different Green Solvents at (298.15 to 338.15) K:	https://www.doi.org/10.1021/je501036r
Solubility Determination of 2-Chlorobenzoic Acid and Analysis of Solvent Effect:	https://www.doi.org/10.1021/acs.jced.9b00661
Solubility of (2Z)-N-Cyclohexyl-2-(3-hydroxybenzylidine) Hydrazine Carbothioamide in Different Pure Solvents at (298.15 to 338.15) K:	https://www.doi.org/10.1021/je5003708
Solubility Measurement and Correlation for epsilon-2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Different Alkanes/Aromatic Hydrocarbon + Ethyl Acetate Binary Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00353
Volumetric and transport properties of ternary mixtures containing 1-alkanol + ethyl acetate + cyclohexane at 298.15 K: Experimental data, correlation and prediction by ERAS model:	https://www.doi.org/10.1016/j.tca.2009.02.022
Solubilities of Dihydroxylammonium 5,5'-Bistetrazole-1,1'-diolate in Various Pure Solvents at Temperatures between 293.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00009
The isobaric vapor liquid equilibria of ethyl acetate + acetonitrile with bis(trifluoromethylsulfonyl)imide-based ionic liquids at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2016.06.022
An odd-even effect on solubility of dicarboxylic acids in organic solvents:	https://www.doi.org/10.1016/j.jct.2014.05.009
Determination and thermodynamic modelling for 2-methyl-6-nitroaniline solubility in binary solvent mixtures of ethyl acetate + (methanol, ethanol, n-propanol and isopropanol):	https://www.doi.org/10.1016/j.jct.2016.10.043
Solubilities of Betulin in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/je700069g
Solubility of sodium diclofenac in different solvents:	https://www.doi.org/10.1016/j.fluid.2007.07.020
Measurement and correlation of the solubility of 2,3,4,5-tetrabromothiophene in different solvents:	https://www.doi.org/10.1016/j.jct.2012.06.005
Solid-liquid phase equilibrium and dissolution properties of ethyl vanillin in pure solvents:	https://www.doi.org/10.1016/j.jct.2016.10.029
Solubility of 4-methyl-2-nitroaniline in fourteen organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.008
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1059

Solid Liquid Equilibrium and Phase Diagram for the Ternary 4-Chlorophthalic Anhydride + 3-Chlorophthalic Anhydride + Ethyl Acetate System:	https://www.doi.org/10.1021/je3013588
Thermodynamic Models for Determination of the Solubility of Boc-(R)-3-Amino-4-(2,4,5-trifluorophenyl)butanoic acid in Different Pure Solvents and (Tetrahydrofuran + n-Butanol) Binary Mixtures with Temperatures from (283.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.5b00758
Solubility of Desmosterol in Five Organic Solvents:	https://www.doi.org/10.1021/je8006088
Solubility of 1H-1,2,4-Triazole in Ethanol, 1-Propanol, 2-Propanol, 1,2-Propanediol, Ethyl Formate, Methyl Acetate, Ethyl Acetate, and Butyl Acetate at (283 to 363) K	https://www.doi.org/10.1021/je060452c
Measurement and Correlation of Solubility of [illegible] in Different Pure Solvents and Binary Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00721
Densities, Excess Molar Volumes, Viscosities, Speeds of Sound, Excess Isentropic Compressibilities, and Relative Permittivities for Alkyl (Methyl, Ethyl, Butyl, and Isoamyl) Acetates + Glycols at Different Temperatures:	https://www.doi.org/10.1021/je0340248
Solubility of Tiamulin Hydrogen Fumarate in Acetone, Acetonitrile, Ethyl Acetate, Ethyl Formate, and Butyl Acetate from (288.2 to 318.2) K:	https://www.doi.org/10.1021/je800838w
Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.09.028
Solid Liquid Phase Equilibria of N,N'-[1,3-Phenylenebis(methylene)]bis(phosphoramidic acid) P,P,P',P'-Tetraphenyl Ester in 10 Pure Solvents:	https://www.doi.org/10.1021/je4010917
Measurement and Correlation of the Solubility of p-Coumaric Acid in Nine Pure and Water + Ethanol Mixed Solvents at Temperatures from 293.15 to 333.15 K:	https://www.doi.org/10.1021/acs.jced.6b00361
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
Volumetric Behavior and Saturated Pressure for Carbon Dioxide + Ethyl Acetate at a Temperature of 313.15 K:	https://www.doi.org/10.1021/je050514j
Isobaric Vapor Liquid Equilibria for Binary Mixtures of Isoamyl Acetate + Ethyl Acetate at 50 and 100 kPa:	https://www.doi.org/10.1021/acs.jced.8b01053
Density, refraction index and vapor-liquid equilibria of N-methyl-2-hydroxyethylammonium butyrate plus (methyl acetate or ethyl acetate or propyl acetate) at several temperatures:	https://www.doi.org/10.1016/j.jct.2013.02.022
Activity coefficient at infinite dilution measurements for organic solutes (polar and non-polar) in fatty compounds: Saturated fatty acids:	https://www.doi.org/10.1016/j.jct.2012.06.009
Ternary Liquid-Liquid Equilibria for Six Systems Containing Ethylacetate + Ethanol or Acetic Acid + an Imidazolium-Based Ionic Liquid with a Hydrogen Sulfate Anion at 313.2 K:	https://www.doi.org/10.1021/je700342k
Solubility determination and thermodynamic modelling of 2,4-dihydro-5-methyl-2-(4-methylphenyl)-3H-pyrazol-3-one in twelve organic solvents from T = (283.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.08.007
Solubility and polymorphic forms of antibiotic lasalocid sodium in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.04.008
Solubilities of Proxetine Hydrochloride Hemihydrate between 286 K and 363 K:	https://www.doi.org/10.1021/je0498762
Vapour-liquid and chemical equilibria in the ethyl ethanoate + ethanol + propyl ethanoate + propanol system accompanied with transesterification reaction:	https://www.doi.org/10.1016/j.fluid.2012.05.016
Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in N-Butylpyridinium Nitrate/N-Octylpyridinium Nitrate by Blend Inverse Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00244
Thermodynamic study of solubility for pyrazinamide in ten solvents from T = (283.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2017.04.014
Vapor Liquid Equilibrium for Binary System of 2-Propanol + 2-Ethoxy-2-methylpropane and Ethyl Ethanoate + 2-Ethoxy-2-methylpropane at 330 K and 2-Propanol + 2-Ethoxy-2-methylpropane at 328 K:	https://www.doi.org/10.1021/je0342421
Activity Coefficients at Infinite Dilution of Various Solutes in Tetrapropylammonium Bromide + 1,6-Hexanediol Deep Eutectic Solvent:	https://www.doi.org/10.1021/acs.jced.8b00600
Solubilities of Organic Semiconductors and Nonsteroidal Anti-inflammatory Drugs in Pure and Mixed Organic Solvents: Measurement and Modeling with Hansen Solubility Parameter:	https://www.doi.org/10.1021/acs.jced.8b00536
Physical Properties of the Reacting Mixture in the Esterification of Acetic Acid with Ethanol:	https://www.doi.org/10.1021/je049771z
Studies of viscosities of dilute solutions of alkylamine in non-electrolyte solvents. II. Haloalkanes and other polar solvents:	https://www.doi.org/10.1016/j.tca.2004.08.013
Solubilities of 1,3,2-Dioxaphosphorinane-2-methanol-a,5,5-trimethyl-a-phenyl-2-oxide in Selected Solvents between 278.15 K and 343.15 K:	https://www.doi.org/10.1021/je500806u
Solubility of Triethylenediamine in Methanol, Ethanol, Isopropanol, 1-Butanol, and Ethyl Acetate between 276.15 K and 363.15 K:	https://www.doi.org/10.1021/je700186v
Solubilities of 2,6-Dimethylnaphthalene in Six Pure Solvents and Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b01209
Solid-liquid phase equilibrium and ternary phase diagrams of ethenzamide-saccharin cocrystals in different solvents:	https://www.doi.org/10.1016/j.fluid.2016.02.047
Liquid-liquid equilibrium for ternary systems of ethyl acetate/isopropyl acetate + 2,2,3,3-tetrafluoro-1-propanol + water at 298.15, 318.15 K:	https://www.doi.org/10.1016/j.jct.2016.12.006
Henry's constants and activity coefficients of some organic solutes in 1-butyl-3-methylimidazolium hydrogen sulfate and in 1-methyl-3-trimethylsilylmethylimidazolium	https://www.doi.org/10.1016/j.jct.2012.08.022
Ternary Excess Molar Volumes of {Methyltrioctylammonium Bis(trifluoromethylsulfonyl)imide + Methanol + Methyl Acetate or Ethyl Acetate} Systems at (298.15, 303.15, and 313.15) K:	https://www.doi.org/10.1021/je900906j
Solubilities of Methyldiphenylphosphine Oxide in Selected Solvents:	https://www.doi.org/10.1021/je900277m
Solubility Determination and Correlation of Glibenclamide in 11 Monosolvents and (Acetone + Acetonitrile) Binary Solvents from 283.15 K to 323.15 K:	https://www.doi.org/10.1021/acs.jced.8b00717
Solubility of 5-(Dithiolan-3-yl)pentanoic Acid in the Mixed Solvents of Cyclohexane + Ethyl Acetate, Heptane + Ethyl Acetate, and Hexane + Ethyl Acetate:	https://www.doi.org/10.1021/je100048g
Measurement and Correlation of Solubility of 3-Hydroxy-2-naphthoic Acid in Pure and Binary Mixed Organic Solvents from T = (293.15 to 333.15 K):	https://www.doi.org/10.1021/acs.jced.8b00578
Measurement and Correlation of Solubility on Reactive Crystallization of Methyl D-(-)-4-Hydroxyphenylglycinate:	https://www.doi.org/10.1021/acs.jced.6b00415
Solubility modelling and dissolution properties of 5-phenyltetrazole in thirteen mono-solvents and liquid mixtures of (methanol + ethyl acetate) at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2017.04.019
Measurement of the Effect of Salt on Vapor-Liquid Equilibria by Using Headspace Gas Chromatography:	https://www.doi.org/10.1021/je025561f
Henry's Law Constants of Organic Compounds in Water and n-Octane at T = 293.2 K:	https://www.doi.org/10.1021/je900711h
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in ionic liquid [HMMIM][BF4]:	https://www.doi.org/10.1016/j.jct.2018.05.017

Separation of binary mixtures based on limiting activity coefficients data using specific ammonium-based ionic liquid and modelling of thermodynamic functions.	https://www.doi.org/10.1016/j.fluid.2017.12.029
Investigation of Surface Properties and Solubility of 1-Alkyl/Esterimidazolium Halide Ionic Liquids by Density Functional Methods.	https://www.doi.org/10.1021/je500205z
Determination and thermodynamic modeling of solid-liquid phase equilibrium for 3,5-dichloroaniline in pure solvents and ternary 3,5-dichloroaniline + 1,3,5-trichlorobenzene + toluene system:	https://www.doi.org/10.1016/j.jct.2016.03.007
Solubility of 3,5-Dinitrobenzoic Acid in Supercritical Carbon Dioxide with Cosolvent at Temperatures from (308 to 328) K and Pressures from (10 to 21) MPa.	https://www.doi.org/10.1021/je100331h
Solubility of daidzin in different organic solvents and (ethyl alcohol + water) mixed solvents.	https://www.doi.org/10.1016/j.jct.2015.04.025
Solubility and Data Correlation of b-Arbutin in Different Monosolvents from 283.15 to 323.15 K.	https://www.doi.org/10.1021/acs.jced.9b00703
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je1005517
Studies on liquid-liquid interactions of some ternary mixtures by density, viscosity, ultrasonic speed and refractive index measurements.	https://www.doi.org/10.1016/j.tca.2009.07.011
Solid-Liquid Phase Equilibria of the Ternary System (2-Naphthaldehyde + 4-Methylphthalic Anhydride + Ethyl Acetate) at (288.15, 298.15, and 308.15) K.	https://www.doi.org/10.1021/acs.jced.8b00170
Experimental and theoretically study of interaction between organic compounds and tricyanomethanide based ionic liquids.	https://www.doi.org/10.1016/j.jct.2014.12.027
Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the ionic liquid 1-butyl-3-methylimidazolium [illegible]	https://www.doi.org/10.1016/j.jct.2017.07.012
Determination and Correlation of Pyridazin-3-amine Solubility in Eight Organic Solvents at Temperatures Ranging from (288.05 to 333.35) K.	https://www.doi.org/10.1021/je300517q
Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution.	https://www.doi.org/10.1016/j.fluid.2018.09.024
Determination and correlation of vapor liquid equilibrium for binary systems consisting of close-boiling components	https://www.doi.org/10.1016/j.fluid.2007.01.042
Solubility of 1,6,6-Trimethyl-6,7,8,9-tetrahydrophenanthro[1,2-b]furan-10,11-dione in Four Organic Solvents from (283.2 to 323.3) K.	https://www.doi.org/10.1021/je800238y
Experimental and calculated liquid-liquid equilibrium data for water + furfural + solvents:	https://www.doi.org/10.1016/j.fluid.2012.07.032
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of the tris(pentafluoroethyl) trifluorophosphate ([FAP]) anion and a functionalized cation.	https://www.doi.org/10.1016/j.jct.2013.05.035
Solubility of Isoniazid in Various Organic Solvents from (301 to 313) K:	https://www.doi.org/10.1021/je800156m
Solubilities of Naringin and Naringenin in Different Solvents and Dissociation Constants of Naringenin:	https://www.doi.org/10.1021/je501004g
Solubility, Density, and Metastable Zone Width of the (1,4-Dioxan-2-one + Ethyl Acetate) System:	https://www.doi.org/10.1021/je050149w
Measurement of the Solubility of Diclazuril and Discussion of the Effect of the Cosolvent:	https://www.doi.org/10.1021/acs.jced.9b00353
Solubility, Solubility Modeling, and Antisolvent Precipitation of 1,3-Bis(9-carbazolyl)benzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00432
Measurement and correlation of the solubility of maleic acid in acetone-ethyl acetate mixtures:	https://www.doi.org/10.1016/j.tca.2012.03.023
High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid:	https://www.doi.org/10.1016/j.fluid.2017.06.001
Determination and Thermodynamic Analysis of the Solubility of Limonin in Eight Organic Solvents and Ethyl Acetate + 2-Propanol Binary Solvents from 283.2 to 323.2 K.	https://www.doi.org/10.1021/acs.jced.7b00851
Measurement and modeling of high pressure VLE for methyl acetate or ethyl acetate + 2-Butanol. Isobaric data at 1.5 MPa.	https://www.doi.org/10.1021/acs.jced.5b00773
Isothermal Vapor-Liquid Equilibria of ethyl acetate + dibromomethane, or + bromochloromethane or + 1,2-dichloroethane or +1-bromo-2-chloroethane at T = 313.15 K.	https://www.doi.org/10.1021/je034259j
Liquid-liquid equilibrium data for water + formic acid + solvent (butyl acetate, ethyl acetate, and isobutyl alcohol) at T = 291.15 K:	https://www.doi.org/10.1016/j.fluid.2016.01.045
Vapor-Liquid Equilibrium and Excess Enthalpy Data for Systems Containing N,N-Dimethylacetamide:	https://www.doi.org/10.1021/acs.jced.7b00119
Solubility, Model Correlation, and Solvent Effect of 2-Amino-3-methylbenzoic Acid in 12 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01226
(Liquid + liquid) equilibria of (water + 1-propanol + solvent) at T = 298.2K:	https://www.doi.org/10.1016/j.fluid.2005.11.024
Measurement and Correlation of the Solubility of Enrofloxacin in Different Solvents from (303.15 to 321.05) K:	https://www.doi.org/10.1021/je5002158
Solubilities of p-coumaric and caffeic acid in ionic liquids and organic solvents:	https://www.doi.org/10.1016/j.jct.2013.02.013
Solubility Measurement and Phase Equilibrium Modeling of 2-Aminobenzamide in 15 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Solubility of p-Aminobenzoic Acid Potassium in Organic Solvents and Binary (Water + Isopropyl Alcohol) Mixture at Temperatures from (283.15 to 318.15) K:	https://www.doi.org/10.1021/acs.jced.7b00978
Solubility of 2-Ethylanthraquinone in Binary Mixtures of Oligooxymethylene Dimethyl Ethers with Different Number of CH2O Groups (n = 2, 3, and 4 from 293.15 to 343.15 K:	https://www.doi.org/10.1021/acs.jced.6b00334
Binary Solid-Liquid Solubility Determination and Model Correlation of [illegible] in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01181
Measurement and Correlation of Solubilities of 5,5'-Dinitramino-3,3'-bi[1,2,4-triazolate] Carbohydrazide Salt (CBNT) in Various Pure Solvents and a Binary Mixture (Dimethyl Sulfoxide + Water) from 298.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00275
Thermodynamic models for determination of [illegible] solubility in binary solvent mixtures of (acetone, ethyl acetate and 1,4-dioxane + methanol)	https://www.doi.org/10.1016/j.jct.2016.04.007
Solubilities of Cinnamic Acid Esters in Organic Solvents:	https://www.doi.org/10.1021/je9004382
Vapor-Liquid Equilibrium for Binary System of 2-Propanol + 1,1-Diethoxyethane at 353 K and Ethyl ethanoate + 1,1-Diethoxyethane at 348 K and 2-Propanone + 1,1-Diethoxyethane at 328 K:	https://www.doi.org/10.1021/je049816w
Solubility determination and thermodynamic modeling of [illegible]benzoic acid in twelve organic solvents from T = (283.15-318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.07.023
Lamotrigine Solubility in Some Nonaqueous Solvent Mixtures at 298.2 K:	https://www.doi.org/10.1021/acs.jced.5b00355

Excess molar enthalpies of diethyl malonate+ (1-butanol, 2-methyl-1-propanol, 1-pentanol, n-heptane, and ethyl acetate) at T = (288.2, 298.2, 313.2, 328.2, 338.2, and 348.2) K and p = 101.3 kPa	https://www.doi.org/10.1016/j.fluid.2009.11.030
Density, Refractive Index, and Speed of Sound at 298.15 K and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Ethyl Acetate + 1-Pentanol and Ethanol + 2-Methyl-1-propanol.	https://www.doi.org/10.1021/je030218b
Solubilities of 4-Amino-4?-nitrodiphenyl Sulfide in Six Organic Solvents from 278.15 K to 363.55 K:	https://www.doi.org/10.1021/je4000982
Dynamic Viscosities, Densities, and Speed of Sound and Derived Properties of the Binary Systems Acetic Acid with Water, Methanol, Ethanol, Ethyl Acetate and Methyl Acetate at T = (293.15, 298.15, and 303.15) K at Atmospheric Pressure.	https://www.doi.org/10.1021/je0342825
PVT Property Measurements for Some Aliphatic Esters from (298 to 393) K and up to 35 MPa	https://www.doi.org/10.1021/je060339m
4-Amino-3,6-dichloropyridazine Solubility Measurement and Correlation in Seven Pure Organic Solvents from (278.15 to 333.15) K.	https://www.doi.org/10.1021/je500286x
Study on thermo-physical and excess molar properties of binary systems of ionic liquids. I: [Cnmim][PF6] (n = 6, 8) and alkyl acetates:	https://www.doi.org/10.1016/j.jct.2014.01.012
The dissolution behaviour and apparent thermodynamic analysis of diprophylline in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2019.06.007
Solubility, liquid-liquid equilibrium and critical states for the quaternary system acetic acid-ethanol-ethyl acetate-water at 303.15 K and 313.15 K:	https://www.doi.org/10.1016/j.fluid.2014.04.013
Thermal Stability and Vapor-Liquid Equilibrium for Imidazolium Ionic Liquids as Alternative Reaction Media:	https://www.doi.org/10.1021/je5009558
Solubility of cinnamic acid in supercritical carbon dioxide and subcritical 1,1,1,2-tetrafluoroethane: Experimental data and modelling:	https://www.doi.org/10.1016/j.fluid.2018.10.009
Liquid-Liquid Equilibrium Diagrams of Ethanol + Water + (Ethyl Acetate or 1-Pentanol) at Several Temperatures:	https://www.doi.org/10.1021/je060054+
Densities, speeds of sound, isentropic compressibilities, refractive indexes, and viscosities of tetrahydrofuran with haloalkane or alkyl ethanoate at T = 303.15 K:	https://www.doi.org/10.1016/j.tca.2004.08.001
Solubility of Canthaxanthin in Pure Solvents from (293.15 to 343.15) K:	https://www.doi.org/10.1021/je800801x
Solubility Behavior of 2-Chloro-3-(trifluoromethyl)pyridine in (Ethyl Acetate + n-Butanol, DMF + n-Butanol, DMF + Ethanol) Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00989
Investigation of the Solubility of Octahydro-1,3,5,7-tetranitro-1,3,5,7-tetrazocine and 1,3,5-Triacetylhexahydro-s-triazine:	https://www.doi.org/10.1021/je301237t
Solubilities of 2,4-Dinitro-L-phenylalanine in Monosolvents at (273.15 to 368.15) K:	https://www.doi.org/10.1021/je100352r
Experimental and Modeling Studies on the Solubility of 2-Chloro-N-(4-methylphenyl)propanamide (S1) in Binary Ethyl Acetate + Hexane, Toluene + Hexane, Acetone + Hexane, and Butanone + Hexane Solvent Mixtures Using Polythermal Method:	https://www.doi.org/10.1021/acs.jced.7b00288
Solubility of o-Nitrobenzoic Acid in Modified Supercritical Carbon Dioxide at (308 to 328) K:	https://www.doi.org/10.1021/je300023v
Solubility and Dissolution Thermodynamics for 1,3,5-Trichlorobenzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.5b00607
Measurement and correlation solubility and mixing properties of L-malic acid in pure and mixture organic solvents:	https://www.doi.org/10.1016/j.fluid.2013.10.011
Measurement and correlation of the solubilities of tetra(5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy) neopentane in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2014.01.043
Solubility of naproxen in ethyl acetate + ethanol mixtures at several temperatures and correlation with the Jouyban-Acree model:	https://www.doi.org/10.1016/j.fluid.2012.02.009
Measurement and Correlation of the Solubility of e-CL-20 in 12 Organic Solvents at Temperatures Ranging from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.6b00761
Liquid-Liquid Equilibrium for the Quaternary System Toluene + Water + Propionic Acid + Ethyl Acetate at 25 deg C and Atmospheric Pressure:	https://www.doi.org/10.1021/je0496488
Solubility and thermodynamic function of a bioactive compound bergenin in various pharmaceutically acceptable neat solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.05.003
Solubility of Baicalin in Methanol, Ethanol, and Ethyl Acetate from (298.2 to 328.2) K:	https://www.doi.org/10.1021/je900171g
Solubility of Sulfachloropyridazine in Pure and Binary Solvent Mixtures and Investigation of Intermolecular Interactions:	https://www.doi.org/10.1021/acs.jced.7b01134
KDB:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1059
Solubility measurement and modelling of 1,8-dinitronaphthalene in nine organic solvents from T = (273.15 to 308.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2015.07.010
Partition Coefficients of Organic Compounds in New Imidazolium and Tetralkylammonium Based Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je9003178
Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate:	https://www.doi.org/10.1016/j.jct.2012.09.017
Determination and correlation of solubility and solution thermodynamics of ethenzamide in different pure solvents:	https://www.doi.org/10.1016/j.fluid.2016.08.019
Thermodynamic functions for solubility of 3-nitro-o-toluic acid in nine organic solvents from T = (283.15 to 318.15) K and apparent thermodynamic properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.017
Thermodynamic models for determination of the solubility of 4-(4-aminophenyl)-3-morpholinone in different pure solvents and (1,4-dioxane + ethyl acetate) binary mixtures with temperatures from (278.15 to 333.15) K	https://www.doi.org/10.1016/j.jct.2016.01.022
4-Nitrobenzyl Cyanide Solubility in Several Nonaqueous Solvent Mixtures: Determination and Modeling:	https://www.doi.org/10.1021/acs.jced.9b00005
Densities, Viscosities, and Refractive Indices for Binary and Ternary Mixtures of N,N-Dimethylacetamide (1) + 2-Methylbutan-2-ol (2) + Ethyl Acetate (3) at 298.15 K for the Liquid Region and at Ambient Pressure:	https://www.doi.org/10.1021/je050538q
Aqueous Solubility Prediction Method:	http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDa
Solubilities of Rutaecarpine in Twelve Organic Solvents from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je4001334
Solubility of a-Sitosteryl Maleate and Stigmasteryl Maleate in Acetone and Ethyl Acetate:	https://www.doi.org/10.1021/je7000396
Measurement of VLE Data by Using an Experimental Installation with Automatic Control: Modeling of Binary Systems of Methyl Acetate or Ethyl Acetate with n-Heptane or 2,2,4-Trimethylpentane at Both 0.1 and 1.5 MPa:	https://www.doi.org/10.1021/acs.jced.9b00668
Solubility Determination, Correlation and Solute-Solvent Molecular Interactions of 5-Aminotetrazole in Various Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00385
Dortmund Data Bank Vapor-Liquid Equilibrium Data:	http://www.ddbst.com/en/EED/VLE/VLE%20Acetonitrile%3BEthyl acetate.php

Effect of Cosolvents on the Solubility of Lenalidomide and Thermodynamic Model Correlation of Data:	https://www.doi.org/10.1021/acs.jced.9b00331
Liquid-liquid equilibria of water + 3-hydroxy-2-butanone + ethyl ethanoate:	https://www.doi.org/10.1016/j.fluid.2011.03.004
Solubilities of Some Phosphaspirocyclic Compounds in Selected Solvents:	https://www.doi.org/10.1021/je060138i
Solubility Measurement and Correlation for e-2,4,6,8,10,12-Hexanitro-2,4,6,8,10,12-hexaazaisowurtzitane in Five Organic Solvents at Temperatures between 283.15 and 323.15 K:	https://www.doi.org/10.1021/acs.jced.6b00664
Determination, Correlation, and Application of Sodium L-Ascorbate Solubility in [illegible] Solvents and Two Binary [illegible]	https://www.doi.org/10.1021/acs.jced.7b00846
Effects of the presence of ethylacetate or benzene on the densities and [illegible] mixtures	https://www.doi.org/10.1016/j.jct.2007.02.007
Solubility of [illegible] in Ethyl Acetate [illegible] Supercritical Carbon Dioxide with and without Cosolvents:	https://www.doi.org/10.1021/je401082x
Densities and Viscosities of Binary Liquid Mixtures of Trichloroethylene and Tetrachloroethylene with Some Polar and Nonpolar Solvents:	https://www.doi.org/10.1021/je034203p
Determination and Correlation of Vapor Liquid Equilibrium Data for the Ethyl Acetate + Hexamethyl Disiloxane System at 101.3 kPa:	https://www.doi.org/10.1021/je2008445
Measurement and Correlation of the Solubility of Azoxystrobin in Seven Monosolvents and Two Different Binary Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.7b00669
Determination of Solubility and Thermodynamic Properties of Benzophenone in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00196
(Solid + liquid) phase equilibria of tetraphenyl piperazine-1,4-diyldiphosphonate in pure solvents:	https://www.doi.org/10.1016/j.jct.2014.06.021
Determination of Activity Coefficients at Infinite Dilution of Solutes in N,N'-Di(2-ethylhexyl)isobutyramide Using Inverse Gas-Liquid Chromatography:	https://www.doi.org/10.1021/acs.jced.8b00635
Solubility of 16.alpha.,17.alpha.-Epoxyprogesterone in Six Different Solvents:	https://www.doi.org/10.1021/je050195w
Liquid-liquid equilibrium data for ternary systems of water + acetic acid + acetate esters at 293.2 K and 303.2 K and ~ 95 kPa:	https://www.doi.org/10.1016/j.fluid.2018.02.003
Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Acetonitrile + 1-Butyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.5b00995
A model to calculate concentration dependent surface tension of binary systems:	https://www.doi.org/10.1016/j.fluid.2016.04.007
Solubility and solution thermodynamics of 2-methyl-4-nitroaniline in eleven organic solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.037
Measurement and correlation solubility and mixing properties of dimethyl [illegible]ate in pure and mixture organic solvents from (278.15 to 323.15) K:	https://www.doi.org/10.1016/j.fluid.2013.06.037
Thermodynamic models for determination of solid liquid equilibrium of the 6-benzyladenine in pure and binary organic solvents:	https://www.doi.org/10.1016/j.jct.2016.12.002
Solubility Study of (2E)-1-(3-Pyridyl)-3-(dimethylamino)-2-propen-1-one in Different Pure Solvents and Binary Solvent Mixtures from 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00047
A novel static analytical apparatus for phase equilibrium measurements:	https://www.doi.org/10.1016/j.fluid.2012.11.008
Vapor liquid equilibria for the quaternary reactive system ethyl acetate + ethanol + water + acetic acid and some of the constituent binary systems at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2005.07.010
Solubility modelling and thermodynamic dissolution functions of phthalimide in ten organic solvents:	https://www.doi.org/10.1016/j.jct.2015.10.024
Solubility of (+)-(S)-2-(6-Methoxynaphthalen-2-yl) Propanoic Acid in Acetone, Methanol, Ethanol, Propan-2-ol, and Ethyl Ethanoate at Temperatures between (278 and 320) K:	https://www.doi.org/10.1021/je8008039
Mixing properties of binary mixtures presenting azeotropes at several temperatures:	https://www.doi.org/10.1016/j.jct.2007.02.008
Measurement and Correlation of Solubility of Gatifloxacin in 12 Pure Solvents from 273.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00902
Determination and Correlation of Solubility Data and Dissolution Thermodynamic Data of L-Lactide in Different Pure Solvents:	https://www.doi.org/10.1021/je301014d
Measurement and Correlation of Solubility of Resorcinol Bis(cyclic 2,2-dimethyl-1,3-propanediol phosphate) in Selected Solvents from T = (293.15 to 353.15) K:	https://www.doi.org/10.1021/acs.jced.7b00695
Differential Scanning Calorimetry Data and Solubility of Rosmarinic Acid in Different Pure Solvents and in Binary Mixtures (Methyl Acetate + Water) and (Ethyl Acetate + Water) from 293.2 to 323.2 K:	https://www.doi.org/10.1021/acs.jced.6b00008
Solubility determination and thermodynamic modeling of 5-nitro-8-hydroxyquinoline in ten organic solvents from T = (278.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.04.018
Solubility of D-phenylglycine methyl ester hydrochloride in water and in organic individual or mixed solvents: Experimental data and results of thermodynamic modeling:	https://www.doi.org/10.1016/j.fluid.2016.02.020
Volumetric and transport properties of ternary mixtures containing 1-propanol + ethyl ethanoate + cyclohexane or benzene at 303.15 K: Experimental data, correlation and prediction by ERAS model:	https://www.doi.org/10.1016/j.tca.2008.11.007
Determination and Modeling of Solid-Liquid Equilibrium for Ternary Systems [illegible] + 4-Hydroxybenzaldehyde + Ethyl Acetate/Acetone:	https://www.doi.org/10.1021/acs.jced.8b00801
Vapor-Liquid Equilibrium of the Carbon Dioxide + Ethyl Propanoate and Carbon Dioxide + Ethyl Acetate Systems at Pressure from 2.96 MPa to 11.79 MPa and Temperature from 313 K to 393 K:	https://www.doi.org/10.1021/je034224j
Solubility of Baicalein in Different Solvents from (287 to 323) K:	https://www.doi.org/10.1007/s10765-014-1740-z
Solubility of 3-chloro-N-phenylphthalimide in ten organic solvents from T = (283.15 to 323.15) K: Determination and modelling:	https://www.doi.org/10.1016/j.jct.2016.01.003
Vapor-Liquid Equilibrium for Several Compounds Relevant to the Biofuels Industry Modeled with the Wilson Equation:	https://www.doi.org/10.1021/je400885z
Thermodynamic properties of enantiotropic polymorphs of glycolide:	https://www.doi.org/10.1016/j.jct.2017.03.011
Solubility determination and modelling for phthalimide in mixed solvents of (acetone, ethyl acetate or acetonitrile + methanol) from (278.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2016.11.023
The Yaws Handbook of Vapor Pressure:	https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure
Correlation and thermodynamic analysis of solubility of diphenhydramine hydrochloride in pure and binary solvents:	https://www.doi.org/10.1016/j.jct.2015.09.032
Determination and Modeling of the Solubility of 2,4-Dimethoxybenzoic Acid in Six Pure and Isopropanol + Ethyl Acetate Mixed Organic Solvents at Temperatures From (288.15 to 323.15) K:	https://www.doi.org/10.1021/je501045s

Isobaric Vapor-Liquid Equilibria for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate at 100 kPa:	https://www.doi.org/10.1021/je700322p
Solubility of Atractylenolide III in Hexane, Ethyl Acetate, Diethyl Ether, and Ethanol from (283.2 to 323.2) K:	https://www.doi.org/10.1021/je6005482
Solubilities of Ibuprofen in Different Pure Solvents:	https://www.doi.org/10.1021/je100255z
Density, Refractive Index, Speed of Sound at 298.15 K, and Vapor-Liquid Equilibria at 101.3 kPa for Binary Mixtures of Ethyl Acetate + Ethyl Lactate and Methyl Acetate + Ethyl Lactate:	https://www.doi.org/10.1021/je060052p
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Separation of ethyl acetate and 2-propanol azeotropic mixture using ionic liquids as entrainers:	https://www.doi.org/10.1016/j.fluid.2016.08.011
Isothermal Vapor-Liquid Equilibria for Binary Mixtures of Methyl Nonafluorobutyl Ether + Acetone, Cyclopentyl Methyl Ether, Ethyl Acetate, n-Heptane, Methanol, and Toluene:	https://www.doi.org/10.1021/acs.jced.7b00599
New high pressure vapor-liquid equilibrium data and density predictions for carbon dioxide + ethyl acetate system:	https://www.doi.org/10.1016/j.fluid.2012.03.028
Measurement and correlation of solubility of boscalid with thermodynamic analysis in pure and binary solvents from 288.15 K to 313.15 K:	https://www.doi.org/10.1016/j.jct.2017.04.017
Density and Speed of Sound of Binary Mixtures of N-Methylacetamide with Ethyl Acetate, Ethyl Chloroacetate, and Ethyl Cyanoacetate in the Temperature Interval (303.15 to 318.15) K:	https://www.doi.org/10.1021/je060343y
Measurement and Correlation of Solubility of Azithromycin Monohydrate in Five Pure Solvents:	https://www.doi.org/10.1021/je400899e
Solubility of Naphthalene in Isobutyl Acetate, n-Butyric Acid, Ethyl Acetate, N-Methyl Pyrrolidone, N,N-Dimethylformamide, and Tetrahydrofuran:	https://www.doi.org/10.1021/je800428r
Solubilities of 4-(Hydroxymethyl)-1-oxido-2,6,7-trioxa-1-phosphabicyclo[2.2.2]octane in Organic Solvents:	https://www.doi.org/10.1021/je4009816
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C141786&Units=SI
Experimental Determination of Densities and Isobaric Vapor Liquid Equilibria of Methyl Acetate and Ethyl Acetate with Alcohols (C3 and C4) at 0.3 MPa:	https://www.doi.org/10.1007/s10765-013-1483-2
Solubility determination and thermodynamic dissolution functions of 1,3-diphenylguanidine in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2016.03.011
Solubilities of 4',5,7-Triacetoxyflavanone in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00929
A volumetric and viscosity study for the binary mixtures of 1-hexyl-3-methylimidazolium tetrafluoroborate with some molecular solvents:	https://www.doi.org/10.1016/j.jct.2010.12.027
Thermodynamics of risperidone and solubility in pure organic solvents:	https://www.doi.org/10.1016/j.fluid.2014.04.028
Co-solvence phenomenon and thermodynamic properties of edaravone in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2019.06.018
Solubility of Decabromodiphenyl Ether in Different Solvents at (283.0 to 323.0) K:	https://www.doi.org/10.1021/je100316k
Phase equilibria for reactive distillation of diethyl succinate. Part II: Systems diethyl succinate + ethyl acetate + water and diethyl succinate + acetic acid + water:	https://www.doi.org/10.1016/j.fluid.2009.11.013
Solubility Measurement and Modeling of 2-Amino-4,6-dichloropyrimidine in Ten Pure Solvents and (Ethyl Acetate + Ethanol) Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.8b00292
Measurement and Thermodynamic Modeling of the Solubility of Lamotrigine, Deferiprone, Cefixime Trihydrate, and Cephalexin Monohydrate in Different Pure Solvents from (283.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.6b00163
Activity coefficient at infinite dilution, azeotropic data, excess enthalpies and solid-liquid-equilibria for binary systems of alkanes and aromatics with esters:	https://www.doi.org/10.1016/j.fluid.2004.12.005
Solubility determination and thermodynamic modelling of 3,5-dimethylpyrazole in nine organic solvents from T = (283.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.02.011
Isobaric Vapor-Liquid Equilibrium for Three Binary Systems of Acetaldehyde + Ethanol, Ethyl Acetate, and 1-Butanol at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.7b00231
Determination and Correlation of Solubility and Thermodynamic Properties of trans-Cinnamyl Alcohol in Pure and Binary Solvents from 253.15 K to 298.15 K:	https://www.doi.org/10.1021/acs.jced.7b00665
Effect of Solvent Properties and Composition on the Solubility of [illegible] Form I:	https://www.doi.org/10.1021/acs.jced.8b01080
Measurements of activity coefficient at infinite dilution for organic solutes in tetramethylammonium chloride + ethylene glycol deep eutectic solvent using gas-liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2018.01.019
Solubility Determination of 1,5-Naphthalenediamine and 1,8-Naphthalenediamine in Different Solvents and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01214
Measurement and correlation of solubility of trans-resveratrol in 11 solvents at T = (278.2, 288.2, 298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1016/j.jct.2007.10.006
Solubility determination and thermodynamic modelling for 2-amino-4-chlorobenzoic acid in eleven organic solvents from T = (278.15 to 318.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.019
Measurement and correlation of solubilities of apigenin and apigenin 7-O-rhamnosylglucoside in seven solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2010.09.013
Determination of the solubility, dissolution enthalpy and entropy of 3-pentadecylphenyl acrylate in different organic solvents:	https://www.doi.org/10.1016/j.tca.2012.06.025
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Butyl-3-methylimidazolium Nitrate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je200050q
Cosolvency and Mathematical Modeling Analysis of Chloroxine in Some Binary Solvent System:	https://www.doi.org/10.1021/acs.jced.8b00257
Solubility in Different Solvents, Correlation, and Solvent Effect in the Solvent Crystallization Process of Iohexol:	https://www.doi.org/10.1021/acs.jced.8b01101
Vapor-liquid equilibria for the binary mixtures of 2,3-butanediol with n-butanol, n-butyl acetate, and ethyl acetate at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2007.09.001
Determination and correlation of solubility and solution thermodynamics of saccharin in different pure solvents:	https://www.doi.org/10.1016/j.jct.2019.02.005
Solid-Liquid Phase Equilibrium and Thermodynamic Analysis of N,N'-Diethylthiourea in Different Solvent Systems:	https://www.doi.org/10.1021/acs.jced.9b00854
Measurement and Correlation of Solubility of Amorphous Cefmetazole Sodium in Pure Solvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00741
Solubility Determination, Modeling, and Thermodynamic Dissolution Properties of Benzenesulfonamide in 16 Neat Solvents from 273.15 to 324.45 K:	https://www.doi.org/10.1021/acs.jced.9b00360

Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and influence of interfacial adsorption using gas liquid chromatography	https://www.doi.org/10.1016/j.jct.2013.10.038
Phase equilibrium measurements and thermodynamic modeling of {CO_2 + ethyl acetate + ethanol} systems.	https://www.doi.org/10.1016/j.fluid.2019.112285
Solubility of citalopram hydrobromide in (ethanol + toluene) and (ethanol + ethyl acetate) at 283.15–318.15 K and its correlation with the Jouyban-Acree and CNIBS/R-K models.	https://www.doi.org/10.1016/j.fluid.2018.09.018
Temperature dependence of limiting activity coefficients, Henry's law constants, and related infinite dilution properties of C4–C6 isomeric n-alkyl ethanoates/ethyl n-alkanoates in water. Measurement, critical compilation, correlation, and recommended data.	https://www.doi.org/10.1016/j.fluid.2014.05.023
Determination and correlation of terephthaldialdehyde solubility in (ethanol + n-propanol/ethyl acetate/N,N-dimethylformamide) mixed solvents at 278.15 K to 318.15 K.	https://www.doi.org/10.1016/j.jct.2016.10.044
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(3-hydroxypropyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)amide:	https://www.doi.org/10.1016/j.jct.2015.02.024
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide.	https://www.doi.org/10.1016/j.jct.2012.01.019
Solubility of Dichloronitrobenzene in Eight Organic Solvents from T = (278.15 to 303.15) K: Measurement and Correlation	https://www.doi.org/10.1021/je401044h
Solubility parameter of lenalidomide for predicting the type of solubility profile and application of thermodynamic model	https://www.doi.org/10.1016/j.jct.2018.12.035
Solid–liquid phase equilibrium and mixing properties of 2-Cyano-4'-methylbiphenyl in pure solvents:	https://www.doi.org/10.1016/j.jct.2016.07.050
Measurement and Correlation of the Solubility of Aminocaproic Acid in Some Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.9b00543
Excess molar enthalpies for mixtures of supercritical CO_2 and ethyl acetate and their role in supercritical fluid applications:	https://www.doi.org/10.1016/j.jct.2012.02.034
Solubility Determination and Modeling of Sodium Mononitrobenzenesulphonate in Different Solvents at Temperatures Ranging from (288.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00783
Solubilities of benzene sulfonamide in supercritical CO_2 in the absence and presence of cosolvent	https://www.doi.org/10.1016/j.tca.2011.10.023
The dissolution behavior and apparent thermodynamic analysis of temozolomide in pure and mixed solvents:	https://www.doi.org/10.1016/j.jct.2018.11.026
Measurement and correlation of solubility of ciclesonide in seven pure organic solvents:	https://www.doi.org/10.1016/j.jct.2016.10.014
Solubility of propanoic acid 3-(hydroxyphenylphosphinyl)-sodium salts in different solvent:	https://www.doi.org/10.1016/j.fluid.2013.02.001
Solubilities of 2-(6-Oxido-6H-dibenz[c,e][1,2]oxaphosphorin-6-yl)-methanol in the Selected Solvents:	https://www.doi.org/10.1021/je050488p
Thermodynamic properties of (an ester + an alkane). XVI. Experimental HEm and VEm values and a new correlation method for (an alkyl ethanoate + an n-alkane) at 318.15 K:	https://www.doi.org/10.1016/j.jct.2003.11.008
Solubility measurement and thermodynamic model correlation of baricitinib in twelve pure solvents:	https://www.doi.org/10.1021/acs.jced.9b00693
Vapor liquid equilibria for the ternary system of carbon dioxide + ethanol + ethyl acetate at elevated pressures:	https://www.doi.org/10.1016/j.fluid.2006.01.017
Solubilities of Naringin Dihydrochalcone in Pure Solvents and Mixed Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00543
Solubility of Butyl Paraben in Methanol, Ethanol, Propanol, Ethyl Acetate, Acetone, and Acetonitrile:	https://www.doi.org/10.1021/je1006289
Thermodynamic equilibrium of hydrocortisone form I in different solvent systems:	https://www.doi.org/10.1016/j.fluid.2017.09.007
Thermodynamics and activity coefficients at infinite dilution for organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028
Solubility of 3-Aminopyridine in Acetone, Chloroform, and Ethyl Acetate:	https://www.doi.org/10.1021/je800768x
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028
Activity coefficients at infinite dilution of organic solutes in the ionic liquid ethyl(2-hydroxyethyl)dimethyl-ammonium diethylphosphate using gas liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2012.04.008
Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethyl acetate from (298.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2012.02.033
Properties of ionic liquid HMIMPF6 with carbonates, ketones and alkyl acetates:	https://www.doi.org/10.1016/j.jct.2005.07.020
Separation of ethyl acetate and ethanol azeotrope mixture using dialkylphosphates-based ionic liquids as entrainers:	https://www.doi.org/10.1016/j.fluid.2017.09.016
Solubility of Diamantane, Trimantane, Tetramantane, and Their Derivatives in Organic Solvents:	https://www.doi.org/10.1021/je800277a
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00425
Liquid liquid equilibria of water + 2,3-butanediol + ethyl acetate at several temperatures:	https://www.doi.org/10.1016/j.fluid.2008.01.009
High-Pressure Vapor-Liquid Equilibrium Data for Systems Involving Carbon Dioxide + Organic Solvent + β-Carotene:	https://www.doi.org/10.1021/je700125v
Solubility of Artemisinin in Seven Different Pure Solvents from (283.15 to 323.15) K:	https://www.doi.org/10.1021/je800515w
Isobaric Vapor - Liquid Equilibrium for Ethyl Acetate + Acetonitrile + 1 - Butyl-3-methylimidazolium Hexa fl uorophosphate at 101.3 kPa:	https://www.doi.org/10.1021/je301204w
Solubility and Thermodynamic Behavior of Syringic Acid in Eight Pure and Water + Methanol Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.7b00333
Isobaric Vapor–Liquid Equilibrium Data for Binary Systems of Anisole with Methyl Acetate, Ethyl Acetate, n-Propyl Acetate, and Isopropyl Acetate at 93.9 kPa:	https://www.doi.org/10.1021/acs.jced.8b00063
Solubilities of Adefovir Dipivoxil in Different Binary Solvents at 298.15 K:	https://www.doi.org/10.1021/je700664g
Solubility Determination and Modeling of p-Nitrobenzamide Dissolved in Twelve Neat Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00065
Solubility evaluation and thermodynamic modeling of beta-lapachone in water and ten organic solvents at different temperatures:	https://www.doi.org/10.1016/j.fluid.2018.05.005
Measurement and Correlation of Solubilities of Indole-2-carboxylic Acid in Ten Different Pure Solvents from (278.15 to 360.15) K:	https://www.doi.org/10.1021/je400813d

Solubility of Veratric Acid in Eight Monosolvents and Ethanol + 1-Butanol at Various Temperatures:	https://www.doi.org/10.1021/je400029t
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2017.11.017
Thermodynamic Parameters of a New Synthesized Tricationic Ionic Liquid Stationary Phase by Inverse Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.8b00601
Thermodynamics and activity coefficients at infinite dilution for organic solutes, water and diols in the ionic liquid choline bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2014.04.024
Solubilities of Phosphoramidic Acid, N-(phenylmethyl)-, Diphenyl Ester in Selected Solvents: Thermodynamic Functions for [illegible]	https://www.doi.org/10.1021/acs.jced.5b00007
Solubility of 1-Hydroxybenzotriazole in Sixteen Solvents at Temperatures from (278.15 to 313.15) K and Mixing Property of Mixtures:	https://www.doi.org/10.1021/acs.jced.7b00316
Solubility Determination and Thermodynamic Mixing Properties of 5-Methyl-2-pyrazinecarboxylic Acid in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00406
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1,4-diaza[2.2.2]bicyclooctanium bis(trifluoromethylsulfonyl)imide using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.026
Solubility and thermodynamic functions of tebuconazole in nine organic solvents from T = (278.15 to [illegible]	https://www.doi.org/10.1016/j.jct.2016.11.032
Solid–Liquid Equilibrium of Azacyclotridecan-2-one in 15 Pure Solvents from T = 273.15 to 323.15 K: Experimental Determination and Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.8b01193
Solubility of 1-Fluoro-4-(methylsulfonyl)benzene in Five Pure Organic Solvents at Temperatures from (288.40 to 331.50) K:	https://www.doi.org/10.1021/je400965y
Solubility of lansoprazole in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.06.011
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties of organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium tris(pentafluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2015.05.022
Measurement and Correlation of the Solubility of Mifepristone in Eight Pure and Water + Ethanol Mixed Solvents at Temperatures from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b01062
Surface Tension Data of Aqueous Binary Mixtures of Methyl, Ethyl, Propyl, and Butyl Acetates at 298.15 K:	https://www.doi.org/10.1021/je900962v
Solubilities of [illegible] in Water, Methanol, Ethanol, Acetone, and Ethyl Acetate from (273.15 to 343.15) K:	https://www.doi.org/10.1021/je1000928
Experimental and Modeling Studies on the Solubility of d-Pantolactone in Four Pure Solvents and Ethanol-Water Mixtures:	https://www.doi.org/10.1021/je500996n
Liquid-Liquid Equilibrium in Mixtures Containing Propane-1,2,3-triol and Mixtures Containing Vegetable Oils at Atmospheric Pressure:	https://www.doi.org/10.1021/je500545s
Solubility Measurement and the Correlation of 1-Naphthaleneacetic Acid in Pure and Methanol + Water Binary Solvents from T = (278.25 to 323.55) K:	https://www.doi.org/10.1021/acs.jced.6b00816
Thermodynamic study of (alkyl esters + α,ω-alkyl dihalides) I: HE and VE for 25 binary mixtures {xCu-1H2u-1CO2C2H5 + (1-x)α,ω-ClCH2(CH2)v-2CH2Cl}, where u = 1 to 5 and v = 2 to 6:	https://www.doi.org/10.1016/j.jct.2005.03.020
Thermodynamics of solubility of thiomalic acid in different organic solvents from 278.15 K to 333.15 K:	https://www.doi.org/10.1016/j.fluid.2013.05.002
Viscosities and Interfacial Properties of 1-Methyl-3-butylimidazolium Hexafluorophosphate and 1-Isobutenyl-3-methylimidazolium Tetrafluoroborate Ionic Liquids:	https://www.doi.org/10.1021/je0604277
Reactive Extraction of Gallic Acid Using Tributyl Phosphate in Different Classes of Diluents:	https://www.doi.org/10.1021/acs.jced.9b00192
Solubility determination and modelling of 4-nitro-1,2-phenylenediamine in eleven organic solvents from T = (283.15 to 318.15) K and thermodynamic properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.11.014
Dissociation Constants and Solubilities of Dalbergin and Nordalbergin in Different Solvents:	https://www.doi.org/10.1021/acs.jced.5b00098
Solubility Measurement and Modeling of 1-(3-nitrophenyl)Ethanone and 1-(4-nitrophenyl)Ethenone in Nine Pure Organic Solvents from T = (278.15 to 318.15) K and Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00192
Solubility Determination and Thermodynamic Modeling of 2-[illegible] Solvents from T = 278.15 K to T = 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00190
Solubility Measurement and Modeling of 4,4′-Dihydroxydiphenyl Sulfone in Nine Organic Solvents from T = (278.15 to 313.15) K and Thermodynamic Property of Dissolution:	https://www.doi.org/10.1021/acs.jced.5b00714
Solubility of d-Tryptophan and L-Tyrosine in Several Organic Solvents: Determination and Solvent Effect:	https://www.doi.org/10.1021/acs.jced.9b00258
Estimated Solubility Method	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Densities, Viscosities, and Speeds of Sound of Binary Liquid Mixtures of Sulfolane with Ethyl Acetate, n-Propyl Acetate, and n-Butyl Acetate at Temperature of (303.15, 308.15, and 313.15) K:	https://www.doi.org/10.1021/je800653d
Solubility of Anthracene in Binary and Ternary Mixtures of Cyclohexanone, Ethyl Acetate and Methanol at 298.2 K:	https://www.doi.org/10.1021/je900852t
Measurement of Activity Coefficients at Infinite Dilution of Benzene, Toluene, Ethanol, Esters, Ketones, and Ethers at Various Temperatures in Water Using the Dilutor Technique:	https://www.doi.org/10.1021/je049875+
Solubility of Deflazacort in Binary Solvent Mixtures:	https://www.doi.org/10.1021/je800790g
Volumetric properties of ternary (IL + 2-propanol or 1-butanol or 2-butanol + ethyl acetate) systems and binary (IL + 2-propanol or 1-butanol or 2-butanol) and (1-butanol or 2-butanol + ethyl acetate) systems:	https://www.doi.org/10.1016/j.jct.2012.01.002
Solubility determination and thermodynamic models for [illegible] in different solvents and three binary solvent systems: (ethyl acetate + methanol), (ethyl acetate + ethanol) and (ethyl acetate + isopropanol):	https://www.doi.org/10.1016/j.jct.2016.06.014
Determination of the solubility, dissolution enthalpy and entropy of [illegible] acetate in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.01.019
Excess Volumes and Excess Isentropic Compressibilities of Binary Liquid Mixtures of Trichloroethylene with Esters at 303.15 K:	https://www.doi.org/10.1007/s10765-015-1927-y
Solubilities of Phosphorus-Containing Compounds in Selected Solvents:	https://www.doi.org/10.1021/je100341q
Temperature and Pressure Dependence of the Volumetric Properties of Binary Liquid Mixtures Containing Dihaloalkanes:	https://www.doi.org/10.1007/s10765-005-5570-x
Density and Viscosity Correlation for Several Common Fragrance and Flavor Esters:	https://www.doi.org/10.1021/je050001c
Solubility of Thiotriazinone in Ethanol, 1-Propanol, 2-Propanol, 1-Butanol, Acetonitrile, Acetone, Ethyl Acetate, and Water from 293 K to 343 K:	https://www.doi.org/10.1021/je900177h
Solubility of Salicylic Acid in Water, Ethanol, Carbon Tetrachloride, Ethyl Acetate, and Xylene:	https://www.doi.org/10.1021/je7004962

Solid Liquid Phase Equilibrium of Phosphoramidic acid, N,N'-1,2-Ethanediylbis-P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je5010565
Measurement and Correlation of Solubility of Hydrochlorothiazide in Monosolvents and Binary Solvent Mixtures from 283.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00220
Excess Molar Enthalpies of Five Binary Systems Containing Ethyl Acetoacetate at Different Temperatures:	https://www.doi.org/10.1021/je800889y
Measurement, correlation and dissolution thermodynamics of biological active chalcone in organic solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2014.03.026
Thermodynamic Properties of Mixtures Containing Precursors of Vitamin B5:	https://www.doi.org/10.1021/je020060f
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium tris(perfluoroethyl)trifluorophosphate:	https://www.doi.org/10.1016/j.jct.2012.05.017
Solubility of 11r-Hydroxy-16r,17r-Epoxyprogesterone in Different Solvents between 283 K and 323 K:	https://www.doi.org/10.1021/je0495435
Solubility of 2-isopropylimidazole in nine pure organic solvents and liquid mixture of (methanol + ethyl acetate) from T = (278.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2016.12.028
Solubility of dodecanedioic acid in pure solvent systems: An experimental and Experimental measurement and thermodynamic modelling:	https://www.doi.org/10.1016/j.jct.2015.05.026
Isothermal vapor liquid equilibria for different binary mixtures involved in the alcoholic distillation:	https://www.doi.org/10.1016/j.fluid.2008.03.002
Solubility of Tetranitrodiglycoluril (TNDGU) in Different Solvents at Temperatures between 293.15 K and 313.15 K:	https://www.doi.org/10.1021/je500392g
Solubilities of Phloretin in 12 Solvents at Different Temperatures:	https://www.doi.org/10.1021/je101168w
Solubility and Mixing Thermodynamics Properties of p-Toluenesulfonamide and o-Toluenesulfonamide in Seven Monosolvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00714
Densities and Viscosities of Ternary Mixtures of Cyclohexane + Cyclohexanone + Some Alkyl Acetates at 298.15 K:	https://www.doi.org/10.1021/je8007217
Solid liquid phase equilibrium of glyphosate in selected solvents:	https://www.doi.org/10.1016/j.fluid.2012.05.003
Solubility of Gastrodin in Pure and Mixed Solvents at 273.15-313.15 K and its Correlation with Different Thermodynamic Models:	https://www.doi.org/10.1021/acs.jced.9b00286
Solubility of Gallic Acid in Methanol, Ethanol, Water, and Ethyl Acetate:	https://www.doi.org/10.1021/je700633w
Isobaric Vapor - Liquid Equilibrium for Ethyl acetate + Methanol + Ionic Liquids Ternary Systems at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.5b00546
Effects of temperature and concentration on interactions in (methanol + ethyl acetate) and (ethanol + methyl acetate or ethyl acetate) systems: Insights from apparent molar volume and apparent molar isentropic compressibility study:	https://www.doi.org/10.1016/j.tca.2013.12.016
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid [illegible]	https://www.doi.org/10.1016/j.jct.2011.11.025
Equilibrium Solubility, Model Correlation, and Solvent Effect of Indole-3-acetic Acid in Twelve Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01265
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Ethyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je800218g
Determination and Correlation of Solubility of Phenylbutazone in Monosolvents and Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.6b00911
Saturated Solubility and Thermodynamic Evaluation of l-Tryptophan in Eight Pure Solvents and Three Groups of Binary Mixed Solvents by the Gravimetric Method at T = 278.15-333.15 K:	https://www.doi.org/10.1021/acs.jced.9b00562
Measurement and correlation of solubility and solution thermodynamics of 1,3-dimethylurea in different solvents from T = (288.15 to 328.15) K:	https://www.doi.org/10.1016/j.jct.2015.12.032
Determination and Correlation of Solubility of Quetiapine Fumarate in Nine Pure Solvents and Two Aqueous Binary Solvents:	https://www.doi.org/10.1021/acs.jced.7b00518
Solubility of Benzyl Disulfide in Five Organic Solvents between (283.45 and 333.15) K:	https://www.doi.org/10.1021/je400326r
Solubility and Molecular Interactions of Trimetazidine Hydrochloride in 12 Monosolvents and Solvent Mixtures of Methanol + (Ethanol, N,N-Dimethylformamide or Ethyl Acetate):	https://www.doi.org/10.1021/acs.jced.8b00235
Solubility and Thermodynamic Modeling of Sulfanilamide in 12 Mono Solvents and 4 Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00411
Solubility and tie-line data for ternary aqueous mixtures of cyclopentanol with organic solvents at T = 298.2 K: Experiments and NRTL model:	https://www.doi.org/10.1016/j.fluid.2014.07.011
Thermodynamic functions for solubility of 2-mercaptobenzothiazole in eleven pure organic solvents at temperatures from 273.15 K to 318.15 K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2017.05.013
Solubility of cefoxitin acid in different solvent systems:	https://www.doi.org/10.1016/j.jct.2016.07.049
The study of excess molar volumes and related properties for binary mixtures containing benzyl alcohol and 1,3-dichloro-2-propanol with vinyl acetate, ethyl acetate and t-butyl acetate at T = 293.15 to 318.15 K and P = 0.087 MPa:	https://www.doi.org/10.1016/j.tca.2016.04.001
Solubility measurement and prediction of clopidogrel hydrogen sulfate polymorphs in isopropanol and ethyl acetate:	https://www.doi.org/10.1016/j.jct.2015.06.019
Solubility of 1,1-diamino-2,2-dinitroethylene in different pure solvents and binary mixtures (dimethyl sulfoxide p water) and (N,N-dimethylformamide p water) at different temperatures:	https://www.doi.org/10.1016/j.fluid.2017.12.035
Solubility of Acephate in Different Solvents from (292.90 to 327.60) K:	https://www.doi.org/10.1021/je0603630
Measurement and Correlation for Solubilities of 16-Dehydropregnenolone Acetate in Different Solvents:	https://www.doi.org/10.1021/je1013262
Measurement and Correlation of Solubility of Two Isomers of Cyanopyridine in Eight Pure Solvents from 268.15 K to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b00301
Thermodynamic Modeling and Solubility Measurement of Cetirizine Hydrochloride and Deferiprone in Pure Solvents of Acetonitrile, Ethanol, Acetic Acid, Sulfolane, and Ethyl Acetate and Their Mixtures:	https://www.doi.org/10.1021/acs.jced.9b00620
Solubility of Valsartan in Ethyl Acetate + Hexane Binary Mixtures from (278.15 to 313.15) K:	https://www.doi.org/10.1021/je900159a
Solubility Determination of Nicotinamide and Its Application for the Cocrystallization with Benzoic Acid:	https://www.doi.org/10.1021/acs.jced.8b00560
Determination and Modeling of Solubility of 4-Aminobenzamide in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b01144
Thermodynamic functions of 1-methyl-4-(methylsulfonyl)benzene solubility in nine organic solvents from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2016.08.021

Thermodynamic equilibrium of 4-hydroxy-2,5-dimethyl-3(2H)-furanone in different solvent systems:	https://www.doi.org/10.1016/j.jct.2015.08.035
Solubilities of Salicylic Acid in Pure Solvents and Binary Mixtures Containing Cosolvent:	https://www.doi.org/10.1021/je800475d
Correlation of Solubility of Bioactive Compound Reserpine in Eight Green Solvents at (298.15 to 338.15) K:	https://www.doi.org/10.1021/je500893g
Solubility of Phenacetin in Methanol, Ethanol, 1-Propanol, 1-Butanol, 1-Pentanol, Tetrahydrofuran, Ethyl Acetate, and Benzene between 282.65 K and 338.20 K:	https://www.doi.org/10.1021/je700209v
Solubility determination and thermodynamic modelling of [illegible] in ten organic solvents from T = 283.15 K to T = 318.15 K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.033
Solubility of Trimethoprim in Selected Pure Solvents and (Water + 2-Propanol) Mixed-Solvent Systems:	https://www.doi.org/10.1021/acs.jced.5b00616
Solubility of Captopril in 2-Propanol, Acetone, Acetonitrile, Methyl Acetate, Ethyl Acetate, and Butyl Acetate:	https://www.doi.org/10.1021/je900401z
Solution Thermodynamics of Benzotriazole in Different Pure Solvents:	https://www.doi.org/10.1021/acs.jced.7b01085
Correlation and prediction of mixing thermodynamic properties of ester-containing systems: Ester + alkane and ester + ester binary systems and the ternary dodecane + ethyl pentanoate + ethyl ethanoate:	https://www.doi.org/10.1016/j.jct.2012.03.011
Measurement and Correlation of the Solubility of Tetramethylpyrazine in Nine Monosolvents and Two Binary Solvent Systems:	https://www.doi.org/10.1021/acs.jced.8b00888
Solubility of 2,5-Furandicarboxylic Acid in Eight Pure Solvents and Two Binary Solvent Systems at 313.15–363.15 K:	https://www.doi.org/10.1021/acs.jced.7b00927
Vapor–Liquid Equilibria for Ethyl Acetate + Methanol at (0.1, 0.5, and 0.7) MPa: Measurements with a New Ebulliometer:	https://www.doi.org/10.1021/je100614r
Phase equilibria and excess properties for binary systems in reactive distillation processes Part II. Ethyl acetate synthesis:	https://www.doi.org/10.1016/j.fluid.2014.05.031
Solubility Measurement, Correlation, and Molecular Interactions of 3-Methyl-6-nitroindazole in Different Neat Solvents and Mixed Solvents from T = 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b01256
Liquid–Liquid Equilibria for Water + Benzonitrile + Ethyl Acetate or + Butyl Acetate:	https://www.doi.org/10.1021/je050051h
Solubility of Penicillin Sulfoxide in Different Solvents:	https://www.doi.org/10.1021/je900326e
Synthesis and Solubility of 5,5-Dimethyl-2-(phenyl(phenylamino)methyl)-1,3,2-dioxaphosphinane 2-oxide in Selected Solvents between 278.15 K and 347.15 K:	https://www.doi.org/10.1021/acs.jced.7b00585
Experimental Measurement and Correlation of the Solubilities of 2,4-Dichloro-5-methoxypyrimidine in Ethyl Ethanoate, Methanol, Ethanol, Acetone, Tetrachloromethane, and Heptane at Temperatures between (295 and 320) K:	https://www.doi.org/10.1021/je9005689
The use of ionic liquids for separation of binary hydrocarbons mixtures based on COSMO-RS analyses and experimental measurements:	https://www.doi.org/10.1016/j.jct.2018.07.024
Experimental measurements of vapor–liquid equilibria at low pressure: Systems containing alcohols, esters and organic acids:	https://www.doi.org/10.1016/j.fluid.2009.10.002
Solubilities of 3-Methoxy-N-phenylaniline and 3-(Methylthio)-N-phenylaniline in Five Organic Solvents (285 K to 333.75 K):	https://www.doi.org/10.1021/je3001816
Measurement and correlation of ternary phase equilibrium of (hexane + ethyl acetate) with four ILs:	https://www.doi.org/10.1016/j.jct.2017.08.022
Separation of Ethyl Acetate and Ethanol by Room Temperature Ionic Liquids with the Tetrafluoroborate Anion:	https://www.doi.org/10.1021/je700516t
The solubility of maleic acid in various inorganic salts and organic solvents: Thermodynamics model and parameters:	https://www.doi.org/10.1016/j.fluid.2017.07.012
Solubility Measurement and Correlation of Probenecid in 12 Pure Organic Solvents and Thermodynamic Properties of Mixing of Solutions:	https://www.doi.org/10.1021/acs.jced.8b00863
Solubilities of Phosphonic Acid, P,P'-(1,4-Piperazinediyl)bis-, P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je400412z
Ternary (liquid + liquid) equilibria of the azeotrope (ethyl acetate + 2-propanol) with different ionic liquids at T = 298.15 K:	https://www.doi.org/10.1016/j.jct.2007.04.010
Equilibrium study and diversified models of drug Norfloxacin in eight pure organic and binary solvents at T = (278.15 to 328.15) K:	https://www.doi.org/10.1016/j.fluid.2016.12.012
Solubility and Melting Properties of Salicylic Acid:	https://www.doi.org/10.1021/je060134d
Thermodynamic models for determination of the solubility of omeprazole sulfide in (ethanol + ethyl acetate) binary solvent mixtures:	https://www.doi.org/10.1016/j.fluid.2015.07.038
Solubility Measurement and Thermodynamic Model Correlation of 5-Nitrosalicylaldehyde in Different Solvents:	https://www.doi.org/10.1021/acs.jced.9b00232
Solubility and solution thermodynamics of sorbic acid in eight pure organic solvents:	https://www.doi.org/10.1016/j.jct.2015.02.004
Measurement and correlation of the solubility of rivaroxaban (form I) in binary mixtures of ethyl acetate with tetrahydrofuran, N,N-dimethylformamide, and N,N-dimethylacetamide from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2015.10.016
Solubility in Different Solvents, Crystal Polymorph and Morphology, and Optimization of Crystallization Process of AIBN:	https://www.doi.org/10.1021/acs.jced.7b00538
Liquid–Liquid Equilibria for the Ternary Systems Water + 2-Methyl-2-cyclopentenone with Ethyl Acetate or Methyl tert-Butyl Ether at 293.2 K:	https://www.doi.org/10.1021/je0502492
Density and Surface Tension of Binary Mixtures of 2,2,4-Trimethylpentane + n-Heptane, 2,2,4-Trimethylpentane + n-Octane, Ethyl Acetate + Benzene, and Butanenitrile + Benzene from (293.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.5b00009
Solubility of 1-methyl-4-nitropyrazole in seventeen pure solvents at temperatures from 283.15 K to 323.15 K:	https://www.doi.org/10.1016/j.fluid.2018.05.023
Binary Equilibrium Solubility of Amidinothiourea in Monosolvents: Experimental Determination, Model Correlation, and Solvent Effect Analysis:	https://www.doi.org/10.1021/acs.jced.9b00458
Thermodynamic equilibrium of hydroxyacetic acid in pure and binary solvent systems:	https://www.doi.org/10.1016/j.jct.2017.01.004
(Solid + liquid) phase diagram for (indomethacin + nicotinamide)-methanol or methanol/ethyl acetate mixture and solubility behavior of 1:1 (indomethacin + nicotinamide) cocrystal at T = (298.15 and 313.15) K:	https://www.doi.org/10.1016/j.jct.2015.01.015
Solubility and limiting activity coefficient of simvastatin in different organic solvents:	https://www.doi.org/10.1016/j.fluid.2009.03.006
Thermodynamic study of the solubility of 2,4'-dihydroxydiphenyl sulfone in nine organic solvents from T = (278.15 to 313.15) K and thermodynamic properties of dissolution:	https://www.doi.org/10.1016/j.jct.2016.07.001
Determination of the excess molar enthalpies of the binary carbon dioxide + ethyl acetate system at 298.15 to 308.15 K and 5.0 to 7.0 MPa:	https://www.doi.org/10.1016/j.fluid.2013.10.038

Determination and Correlation of Excess Molar Enthalpies of Eight Binary Systems Containing Acetophenone at Different Temperatures:	https://www.doi.org/10.1021/je800193a
Thermodynamics of fenoxycarb in solution:	https://www.doi.org/10.1016/j.jct.2013.06.007
Solubility of 3-Nitrophthalic Acid in Different Solvents between 278 K and 353 K:	https://www.doi.org/10.1021/je0604737
Ternary excess molar volumes of {methyltrioctylammonium bis(trifluoromethylsulfonyl)imide + ethanol + methyl acetate, or ethyl acetate} systems at T = (298.15, 303.15, and 313.15) K:	https://www.doi.org/10.1016/j.jct.2010.01.007
Activity Coefficients at Infinite Dilution by GLC in Alkanediamines as Stationary Phases:	https://www.doi.org/10.1021/je200628n
Determination and correlation of solubility and mixing properties of isonicotinamide (form II) in some pure solvents:	https://www.doi.org/10.1016/j.tca.2016.01.012
Dissociation Constants and Solubilities of Daidzein and Genistein in Different Solvents:	https://www.doi.org/10.1021/je4010905
Solubility of 3,4-Bis(3-nitrofurazan-4-yl)furoxan in Common Solvents at Temperatures between 293.15 K and 313.15 K:	https://www.doi.org/10.1021/je4005992
Equilibrium Solubility of Biapenem in Different Neat and Binary Solvents: Experimental Determination and Model Correlation:	https://www.doi.org/10.1021/acs.jced.8b01051
Measurement and Correlation of Solubility of Theobromine, Theophylline, and Caffeine in Water and Organic Solvents at Various Temperatures:	https://www.doi.org/10.1021/acs.jced.7b00065
Solubility of Losartan Potassium in Different Pure Solvents from (293.15 to 343.15) K:	https://www.doi.org/10.1021/je7007236
Solubility and thermodynamic properties of vanillyl alcohol in some pure solvents:	https://www.doi.org/10.1016/j.jct.2016.11.030
Solubilities of Diphenylphosphinic Acid in Selected Solvents:	https://www.doi.org/10.1021/je800049b
Solubility of 3,7,9,11-Tetraoxo-2,4,6,8,10-pentaaza[3.3.3]Propellane (TOPAP) in Different Pure Solvents at Temperatures between 273.15 and 318.15 K:	https://www.doi.org/10.1021/acs.jced.6b00349
Solubility and Solution Thermodynamics of Gibberellin A4 in Different Organic Solvents from 278.15 K to 333.15 K:	https://www.doi.org/10.1021/acs.jced.5b00190
Solubility of 3,7-Dinitro-1,3,5,7-tetraazabicyclo [3.3.1] Nonane in Ethanenitrile, Methanol, 1,1-Dichloroethane, Dimethyl Sulfoxide, Acetone, and Mixed Solvents [illegible]	https://www.doi.org/10.1021/acs.jced.5b00033
Solubility Measurement and Thermodynamic Modeling of 4-Nitrophthalimide in Twelve Pure Solvents at Elevated Temperatures Ranging from (273.15 to 323.15) K:	https://www.doi.org/10.1021/acs.jced.6b00230
Measurement and correlation of the solubility of genistin in eleven organic solvents [illegible]	https://www.doi.org/10.1016/j.jct.2015.05.019
Experimental Determination and Thermodynamic Models for Aniracetam Solubility in Pure and Binary Solvents: Effect of Molecular Interaction and Analysis of Crystallized Products:	https://www.doi.org/10.1021/acs.jced.8b00165
Solid–Liquid Equilibrium of 4-(4-Aminophenoxy)-N-methylpyridine-2-carboxamide in Pure and Binary Solvent Mixtures for T = (278.15–328.15) K:	https://www.doi.org/10.1021/acs.jced.8b00023
Thermodynamics of 4'-bromomethyl-2-cyanobiphenyl in different solvents:	https://www.doi.org/10.1016/j.jct.2014.12.002
Measurement and modelling of econazole nitrate in twelve pure organic solvents at temperatures from 278.15 K to 318.15 K:	https://www.doi.org/10.1016/j.jct.2016.07.043
Thermodynamic Models for Determination of Solid-Liquid Equilibrium of the Buprofezin in Pure and Binary Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00172
Determination and Correlation of Solubility of N-tertbutylacrylamide in Seven Different Solvents at Temperatures between (279.15 and 353.15) K:	https://www.doi.org/10.1021/acs.jced.5b00135
Isobaric Vapor-Liquid Equilibrium Data for the Acetone + Hexamethyl Disiloxane + Ethyl Acetate Ternary System at 101.3 kPa: Determination and Correlation:	https://www.doi.org/10.1021/acs.jced.8b00488
Solubility Measurement and Thermodynamic Properties of Levetiracetam in Pure and Mixed Solvents:	https://www.doi.org/10.1021/acs.jced.8b00720
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-butyl-1-methylpyrrolidinium dicyanamide:	https://www.doi.org/10.1016/j.fluid.2018.06.013
Phase equilibria in binary mixtures with monoethyl succinate:	https://www.doi.org/10.1016/j.fluid.2011.06.020
Molecular interactions in (2,4,6-trimethyl-1,3,5-trioxane + n-alkyl acetates) at T = (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1016/j.jct.2006.03.014
Measurement and Correlation of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01205
Liquid-Liquid Equilibria for the Ternary Systems of FC3283 + HFE7300 + Hexane, FC3283 + HFE7300 + Octane, and FC72 + HFE7100 + (Acetonitrile or Ethyl Acetate) at 273.15 K, 298.15 K, and 313.15 K:	https://www.doi.org/10.1021/je500107y
Solubility of Benzoic Acid and Aspirin in Pure Solvents Using Focused Beam Reflective Measurement:	https://www.doi.org/10.1021/acs.jced.5b00746
Solubilities of some organic solutes in 1-ethyl-3-methylimidazolium acetate. Chromatographic measurements and predictions from COSMO-RS:	https://www.doi.org/10.1016/j.jct.2011.11.005
Development of a Predictive Equation of State for CO2 + Ethyl Ester Mixtures Based on Critical Points Measurements:	https://www.doi.org/10.1021/je5002494
Evaluation of Excess Molar Enthalpies for a Quaternary Mixture System of Acetic Anhydride, Ethanol, Acetic Acid, and Ethyl Acetate:	https://www.doi.org/10.1021/je200595b
Solubility of Toltrazuril and the Effect of a Cosolvent:	https://www.doi.org/10.1021/acs.jced.8b00581
Solubilities of the Gaseous and Liquid Solutes and Their Thermodynamics of Solubilization in the Novel Room-Temperature Ionic Liquids at Infinite Dilution by Gas Chromatography:	https://www.doi.org/10.1021/je700297c
Standard partial molar volumes of some electrolytes in ethylene carbonate based mixtures:	https://www.doi.org/10.1016/j.jct.2003.09.005
Thermodynamic modelling for solubility of 4-nitrobenzaldehyde in different solvents at temperature range from (273.15 to 313.15) K and mixing properties of solutions:	https://www.doi.org/10.1016/j.jct.2016.09.012
Thermodynamic models for determination of the solid-liquid equilibrium of [illegible] in ethyl acetate plus (isopropanol, tetrahydrofuran, acetone) binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2017.03.015
Measurement and Correlation of the Dissolution Equilibria of o-Iodoaniline in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.7b00840
Solubility and Thermodynamic Functions of Isatin in Pure Solvents:	https://www.doi.org/10.1021/je500396b
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.11.021

Measurement and Correlation of the Solubility of Prednisone Acetate in Different Solvents:	https://www.doi.org/10.1021/je101201w
Activity Coefficients at Infinite Dilution in Methylimidazolium Nitrate Ionic Liquids:	https://www.doi.org/10.1021/je100998r
Physical Properties of Binary and Ternary Mixtures of Ethyl Acetate, Ethanol, and 1-Octyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide at 298.15 K:	https://www.doi.org/10.1021/je800899w
Solubility and preferential solvation of 2-methyl-4-nitroaniline in mixed solvents of ethyl acetate + (methanol, ethanol, n-propanol and isopropanol):	https://www.doi.org/10.1016/j.jct.2017.05.044
Separation of binary mixtures based on gamma infinity data using [OMMIM][BF4] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2018.09.003
Solid–liquid Equilibria of the trans-1,2-Cyclohexanediol + Ethyl Acetate + Water Ternary System:	https://www.doi.org/10.1021/je050338e
Solubility and Solution Thermodynamics of 2,6-Dichloro-4-nitroaniline in 12 Pure Solvents at Temperatures from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00778
Solubility and solution thermodynamics of ethyl 5-amino-4-cyano-3-(2-ethoxy-2-oxoethyl)-2-thiophenecarboxylate in nine organic solvents at evaluated temperatures:	https://www.doi.org/10.1016/j.jct.2015.11.025
β-Sitosterol Solubility in Selected Organic Solvents:	https://www.doi.org/10.1021/je9009909
Estimation and confirmation of the thermodynamic stability relationships of the enantiotropic polymorphs of glycolide:	https://www.doi.org/10.1016/j.jct.2017.10.011
Solubility of Cloxacillin Sodium in Different Binary Solvents:	https://www.doi.org/10.1021/je800730k
Activity coefficients at infinite dilution of organic solutes in 1-octyl-3-methylimidazolium nitrate using gas-liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2012.05.006
Solubility of Lovastatin in Acetone, Methanol, Ethanol, Ethyl Acetate, and Butyl Acetate between 283 K and 323 K:	https://www.doi.org/10.1021/je0500781
Solubility measurement and thermodynamic functions of 3-nitrobenzaldehyde in different solvents at elevated temperatures:	https://www.doi.org/10.1016/j.jct.2016.10.040
Solubility Measurement and Modeling of 3-Hydroxy-2-nitropyridine in Ten Pure Solvents and Two Binary Mixed Solvents for T = (278.15–318.15) K:	https://www.doi.org/10.1021/acs.jced.9b00564
Thermodynamic Solubility and Mixing Properties of Phenformin in 14 Pure Solvents at Temperatures Ranging from 278.15 to 323.15 K:	https://www.doi.org/10.1021/acs.jced.9b00844
Solubility and Thermodynamic Modeling of Dimethyl Terephthalate in Pure Solvents and the Evaluation of the Mixing Properties of the Solutions:	https://www.doi.org/10.1021/acs.jced.9b00658
Isobaric Vapor-Liquid Phase Equilibrium Measurements for Allyl Alcohol with Chloroform, Ethyl Acetate, and Methyl Propionate at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.8b00908
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Butyl-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je060305e
Solubility of Vitamin D2 in Six Organic Solvents at Temperatures from (248.2 to 273.2) K:	https://www.doi.org/10.1021/je300401c
Measurement and correlation of solubility of D-camphor-10-sulfonic acid in pure solvents:	https://www.doi.org/10.1016/j.fluid.2015.04.014
Solubility of Chlorpheniramine Maleate in Ethanol, 1-Propanol, 1-Butanol, Benzene, Ethyl Acetate, Ethyl Formate, and Butyl Acetate between 283 K and 333 K:	https://www.doi.org/10.1021/je7001682
Isobaric Vapor-Liquid Equilibrium for Three Binary Systems of Ethyl Acetate + Propyl Acetate, Ethyl Acetate + Propylene Carbonate, and Propyl Acetate + Propylene Carbonate at 101.3 kPa:	https://www.doi.org/10.1021/acs.jced.7b01107
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Isobaric Vapor-Liquid Equilibrium for Ethyl Acetate + Ethanol + 1-Ethyl-3-methylimidazolium Tetrafluoroborate:	https://www.doi.org/10.1021/je800175s
Determination and Correlation of Solubility Data and Dissolution Thermodynamic Data of Cefixime Trihydrate in Seven Pure Solvents:	https://www.doi.org/10.1021/je5000605
Solubility and Melting Properties of Salicylamide:	https://www.doi.org/10.1021/je060178m
Solubility of Lovastatin in Ethyl Acetate, Propyl Acetate, Isopropyl Acetate, Butyl Acetate, sec-Butyl Acetate, Isobutyl Acetate, tert-Butyl Acetate, and 2-Butanone, between (285 and 313) K:	https://www.doi.org/10.1021/je800120p
Experiment and Computation of Solubility and Dissolution Properties for Enalapril Maleate and Its Intermediate in Pure Solvents:	https://www.doi.org/10.1021/acs.jced.8b00333
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Propyl-2,3-dimethylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je700640r
Determination and correlation of solubility of spironolactone form II in pure solvents and binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2014.07.011

Legend

af:	Acentric Factor
affp:	Proton affinity
aigt:	Autoignition Temperature
basg:	Gas basicity
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity

dm:	Dipole Moment
dvisc:	Dynamic viscosity
fll:	Lower Flammability Limit
flu:	Upper Flammability Limit
fpc:	Flash Point (Closed Cup Method)
fpo:	Flash Point (Open Cup Method)
gf:	Standard Gibbs free energy of formation
gyrad:	Radius of Gyration
hf:	Enthalpy of formation at standard conditions
hfl:	Liquid phase enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
nfpaf:	NFPA Fire Rating
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhoc:	Critical density
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature
tbp:	Boiling point at given pressure
tc:	Critical Temperature
tcondl:	Liquid thermal conductivity
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
zc:	Critical Compressibility
zra:	Rackett Parameter