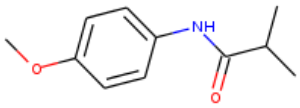
Propanamide, n-(4-methoxyphenyl)-2-methyl-

InChI: InChI=1S/C11H15NO2/c1-8(2)11(13)12-9-4-6-10(14-3)7-5-9/h4-8H,1-3H3,(H,12,13)
InChI Key: VYJJVIXNXXKTFM-UHFFFAOYSA-N
Formula: C11H15NO2
SMILES: COc1ccc(NC(=O)C(C)C)cc1
Molecular Weight: 193.24



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-2.45	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-241.92	kJ/mol	Joback Method
$\Delta_{fus} H°$	22.26	kJ/mol	Joback Method
$\Delta_{vap} H°$	58.22	kJ/mol	Joback Method
$logP_{oct/wat}$	2.29		Crippen Method
P_c	2796.51	kPa	Joback Method
T_{boil}	608.76	K	Joback Method
T_c	823.24	K	Joback Method
T_{fus}	362.49	K	Joback Method
V_c	0.60	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	396.17	J/mol×K	608.76	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method

NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C11H15NO2/c1-8(2)11(13)12-9-4-6-10(14-3)7-5-9/h4-8H,1-3H3,(H,12,13)

Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).

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