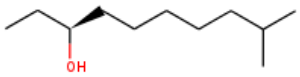
(S)-9-methyldecan-3-ol

InChI: InChI=1S/C11H24O/c1-4-11(12)9-7-5-6-8-10(2)3/h10-12H,4-9H2,1-3H3/t11-/m1/s1
InChI Key: GVNZAPCZQSDLGP-LLVKDONJSA-N
Formula: C11H24O
SMILES: CCC(O)CCCCCC(C)C
Molecular Weight: 172.31



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-99.96	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-433.16	kJ/mol	Joback Method
$\Delta_{fus} H°$	21.29	kJ/mol	Joback Method
$\Delta_{vap} H°$	55.98	kJ/mol	Joback Method
$logP_{oct/wat}$	3.36		Crippen Method
P_c	2131.49	kPa	Joback Method
T_{boil}	542.38	K	Joback Method
T_c	705.81	K	Joback Method
T_{fus}	244.55	K	Joback Method
V_c	0.66	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	429.18	J/mol×K	542.38	Joback Method
η	0.00	Pa×s	542.38	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method
NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C11H24O/c1-4-11(12)9-7-5-6-8-10(2)3/h10-12H,4-9H2,1-3H3/t11-/m1/s1
Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
η: Dynamic viscosity (Pa×s).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).