

# p-(2-Chloroethoxy)benzaldehyde

<b>Inchi:</b>	InChI=1S/C9H9ClO2/c10-5-6-12-9-3-1-8(7-11)2-4-9/h1-4,7H,5-6H2
<b>InchiKey:</b>	HBHHMVNKQWECIS-UHFFFAOYSA-N
<b>Formula:</b>	C9H9ClO2
<b>SMILES:</b>	O=Cc1ccc(OCCCl)cc1
<b>Mol. weight [g/mol]:</b>	184.62
<b>CAS:</b>	54373-15-8

## Physical Properties

Property code	Value	Unit	Source
gf	-88.77	kJ/mol	Joback Method
hf	-237.57	kJ/mol	Joback Method
hfus	20.39	kJ/mol	Joback Method
hvap	52.08	kJ/mol	Joback Method
log10ws	-2.40		Crippen Method
logp	2.117		Crippen Method
mcvol	133.590	ml/mol	McGowan Method
pc	3280.28	kPa	Joback Method
tb	545.49	K	Joback Method
tc	762.26	K	Joback Method
tf	324.28	K	Joback Method
vc	0.515	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	282.18	J/molxK	545.49	Joback Method
cpg	293.39	J/molxK	581.62	Joback Method
cpg	303.95	J/molxK	617.75	Joback Method
cpg	313.88	J/molxK	653.88	Joback Method
cpg	323.19	J/molxK	690.01	Joback Method
cpg	331.89	J/molxK	726.13	Joback Method
cpg	340.01	J/molxK	762.26	Joback Method
dvisc	0.0019031	Paxs	324.28	Joback Method
dvisc	0.0011406	Paxs	361.15	Joback Method

dvisc	0.0007517	Paxs	398.02	Joback Method
dvisc	0.0005316	Paxs	434.88	Joback Method
dvisc	0.0003969	Paxs	471.75	Joback Method
dvisc	0.0003092	Paxs	508.62	Joback Method
dvisc	0.0002491	Paxs	545.49	Joback Method

## Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	419.00 ± 1.00	K	0.30	NIST Webbook

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C54373158&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C54373158&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tbrp:</b>	Boiling point at reduced pressure
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point

**vc:** Critical Volume

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