

# 4,8-diethylidene-2-thiaadamantane

<b>Inchi:</b>	InChI=1S/C13H18S/c1-3-10-8-5-9-7-12(10)14-13(6-8)11(9)4-2/h3-4,8-9,12-13H,5-7H2,1-
<b>InchiKey:</b>	DQYPQUPLZVFUDV-HULALXFYSA-N
<b>Formula:</b>	C13H18S
<b>SMILES:</b>	CC=C1C2CC3CC1SC(C2)C3=CC
<b>Mol. weight [g/mol]:</b>	206.35

## Physical Properties

Property code	Value	Unit	Source
gf	351.80	kJ/mol	Joback Method
hf	77.57	kJ/mol	Joback Method
hfus	27.11	kJ/mol	Joback Method
hvap	51.52	kJ/mol	Joback Method
log10ws	-4.28		Crippen Method
logp	3.793		Crippen Method
mcvol	169.200	ml/mol	McGowan Method
pc	2421.88	kPa	Joback Method
rinpol	1679.00		NIST Webbook
rinpol	1677.00		NIST Webbook
rinpol	1663.00		NIST Webbook
rinpol	1689.00		NIST Webbook
rinpol	1645.00		NIST Webbook
rinpol	1645.00		NIST Webbook
rinpol	1670.00		NIST Webbook
rinpol	1654.00		NIST Webbook
rinpol	1667.00		NIST Webbook
rinpol	1658.00		NIST Webbook
rinpol	1679.00		NIST Webbook
tb	577.77	K	Joback Method
tc	808.16	K	Joback Method
tf	386.50	K	Joback Method
vc	0.637	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	436.51	J/mol×K	577.77	Joback Method
cpg	455.93	J/mol×K	616.17	Joback Method
cpg	473.97	J/mol×K	654.57	Joback Method
cpg	490.76	J/mol×K	692.97	Joback Method
cpg	506.43	J/mol×K	731.37	Joback Method
cpg	521.12	J/mol×K	769.76	Joback Method
cpg	534.96	J/mol×K	808.16	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R207995&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R207995&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307l">http://pubs.acs.org/doi/abs/10.1021/ci990307l</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpola:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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