

# Uroxine

<b>Other names:</b>	5,5'-Bibarbituric acid, 5,5'-dihydroxy- 5,5'-Dihydroxy-5,5'-dibarbituric acid 5,5'-dihydroxy-5,5'-bipyrimidinehexaone Alloxantin Alloxantin, dihydrate Uroxin [5,5'-Bipyrimidine]-2,2',4,4',6,6'(1H,1'H,3H,3'H,5H,5'H)-hexone, 5,5'-dihydroxy-alloxanthin
<b>Inchi:</b>	InChI=1S/C8H6N4O8/c13-1-7(19,2(14)10-5(17)9-1)8(20)3(15)11-6(18)12-4(8)16/h19-20
<b>InchiKey:</b>	IWDDXZKCDHOOSF-UHFFFAOYSA-N
<b>Formula:</b>	C8H6N4O8
<b>SMILES:</b>	O=C1NC(=O)C(O)(C2(O)C(=O)NC(=O)NC2=O)C(=O)N1
<b>Mol. weight [g/mol]:</b>	286.16
<b>CAS:</b>	76-24-4

## Physical Properties

Property code	Value	Unit	Source
gf	-603.94	kJ/mol	Joback Method
hf	-1048.75	kJ/mol	Joback Method
hfus	31.15	kJ/mol	Joback Method
hvap	117.83	kJ/mol	Joback Method
log10ws	-1.99		Estimated Solubility Method
log10ws	-1.99		Aqueous Solubility Prediction Method
logp	-4.819		Crippen Method
mcvol	162.940	ml/mol	McGowan Method
pc	10161.94	kPa	Joback Method
tb	1207.50	K	Joback Method
tc	1497.96	K	Joback Method
tf	1193.56	K	Joback Method
vc	0.574	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	631.77	J/mol×K	1207.50	Joback Method
cpg	641.46	J/mol×K	1255.91	Joback Method
cpg	649.59	J/mol×K	1304.32	Joback Method
cpg	656.09	J/mol×K	1352.73	Joback Method
cpg	660.89	J/mol×K	1401.14	Joback Method
cpg	663.93	J/mol×K	1449.55	Joback Method
cpg	665.16	J/mol×K	1497.96	Joback Method

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>Aqueous Solubility Prediction Method:</b>	<a href="http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDataset002.xlsx">http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDataset002.xlsx</a>
<b>Estimated Solubility Method:</b>	<a href="http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt">http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C76244&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C76244&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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