Tetrahydrofuran

Other names:	1,4-epoxybutane
	Butane «alpha»,«delta»-oxide
	Butane Â«alphaÂ»,Â«deltaÂ»-oxide
	Butane, 1,4-epoxy-
	Cyclotetramethylene oxide
	Diethylene oxide
	Dynasolve 150
	Furan, tetrahydro-
	Furanidine
	Hydrofuran
	NCI-C60560
	NSC 57858
	Oxacyclopentane
	Oxolane
	Rcra waste number U213
	THF
	THF (tetrahydrofuran)
	Tetrahydrofuraan
	Tetrahydrofurane
	Tetrahydrofuranne
	Tetraidrofurano
	Tetramethylene oxide
	UN 2056
Inchi:	InChI=1S/C4H8O/c1-2-4-5-3-1/h1-4H2
InchiKey:	WYURNTSHIVDZCO-UHFFFAOYSA-N
Formula:	C4H8O
SMILES:	C1CCOC1
Mol. weight [g/mol]:	72.11
CAS:	109-99-9

Physical Properties

Property code	Value	Unit	Source
affp	822.10	kJ/mol	NIST Webbook
basg	794.70	kJ/mol	NIST Webbook
chg	-2533.20 ± 0.67	kJ/mol	NIST Webbook
chl	-2505.80 ± 2.10	kJ/mol	NIST Webbook

chl	-2501.20 ± 0.84	kJ/mol	NIST Webbook
dvisc	0.0004750	Pa×s	A volumetric and viscosity study for the binary mixtures of 1-hexyl-3-methylimidazolium tetrafluoroborate with some molecular solvents
gf	-59.06	kJ/mol	Joback Method
hf	-184.20 ± 0.71	kJ/mol	NIST Webbook
hfus	6.96	kJ/mol	Joback Method
hvap	32.00	kJ/mol	NIST Webbook
hvap	32.90	kJ/mol	NIST Webbook
hvap	32.00	kJ/mol	NIST Webbook
hvap	32.16	kJ/mol	NIST Webbook
ie	9.38 ± 0.05	eV	NIST Webbook
ie	9.38	eV	NIST Webbook
ie	9.41	eV	NIST Webbook
ie	9.42 ± 0.01	eV	NIST Webbook
ie	9.40 ± 0.02	eV	NIST Webbook
ie	9.74	eV	NIST Webbook
ie	9.71	eV	NIST Webbook
ie	9.65	eV	NIST Webbook
ie	9.53	eV	NIST Webbook
ie	9.57 ± 0.02	eV	NIST Webbook
ie	9.54	eV	NIST Webbook
log10ws	0.49		Estimated Solubility Method
logp	0.797		Crippen Method
mcvol	62.230	ml/mol	McGowan Method
pc	5190.00 ± 50.66	kPa	NIST Webbook
rinpol	626.00		NIST Webbook
rinpol	640.00		NIST Webbook
rinpol	618.00		NIST Webbook
rinpol	612.90		NIST Webbook
rinpol	618.00		NIST Webbook
rinpol	631.00		NIST Webbook
rinpol	620.00		NIST Webbook
rinpol	626.00		NIST Webbook
rinpol	631.00		NIST Webbook
rinpol	630.00		NIST Webbook
rinpol	611.00		NIST Webbook
rinpol	609.80		NIST Webbook
rinpol	609.90		NIST Webbook
rinpol	617.00		NIST Webbook
rinpol	621.00		NIST Webbook
rinpol	609.00		NIST Webbook

rinpol	632.00	NIST Webbook
rinpol	619.00	NIST Webbook
rinpol	610.00	NIST Webbook
rinpol	632.00	NIST Webbook
rinpol	629.00	NIST Webbook
rinpol	617.00	NIST Webbook
rinpol	624.00	NIST Webbook
rinpol	622.00	NIST Webbook
rinpol	623.00	NIST Webbook
rinpol	638.00	NIST Webbook
rinpol	627.00	NIST Webbook
rinpol	629.00	NIST Webbook
rinpol	612.00	NIST Webbook
rinpol	622.00	NIST Webbook
rinpol	608.00	NIST Webbook
rinpol	638.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	624.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	636.00	NIST Webbook
rinpol	619.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	615.00	NIST Webbook
rinpol	633.00	NIST Webbook
rinpol	620.00	NIST Webbook
rinpol	618.00	NIST Webbook
rinpol	623.20	NIST Webbook
rinpol	614.00	NIST Webbook
rinpol	630.00	NIST Webbook
rinpol	624.00	NIST Webbook
rinpol	620.00	NIST Webbook
rinpol	620.00	NIST Webbook
rinpol	629.00	NIST Webbook
rinpol	638.00	NIST Webbook
rinpol	640.00	NIST Webbook
ripol	898.00	NIST Webbook
ripol	888.00	NIST Webbook
ripol	907.00	NIST Webbook
ripol	915.00	NIST Webbook
ripol	887.00	NIST Webbook
ripol	829.00	NIST Webbook
ripol	900.00	NIST Webbook
ripol	868.00	NIST Webbook
ripol	872.00	NIST Webbook

ripol	895.00		NIST Webbook
ripol	854.00		NIST Webbook
ripol	829.00		NIST Webbook
ripol	857.00		NIST Webbook
ripol	854.00		NIST Webbook
ripol	867.00		NIST Webbook
ripol	866.00		NIST Webbook
ripol	861.00		NIST Webbook
ripol	868.00		NIST Webbook
ripol	903.00		NIST Webbook
ripol	868.00		NIST Webbook
ripol	900.00		NIST Webbook
sg	301.70 ± 1.70	J/mol×K	NIST Webbook
sl	203.90	J/mol×K	NIST Webbook
sl	203.80	J/mol×K	NIST Webbook
sl	203.90	J/mol×K	NIST Webbook
tb	337.65 ± 2.00	K	NIST Webbook
tb	337.65 ± 2.00	K	NIST Webbook
tb	339.65 ± 0.30	K	NIST Webbook
tb	339.10 ± 0.20	K	NIST Webbook
tb	340.00 ± 2.00	K	NIST Webbook
tb	339.00 ± 2.00	K	NIST Webbook
tb	337.90 ± 0.50	K	NIST Webbook
tb	340.15 ± 2.00	K	NIST Webbook
tb	338.90 ± 0.15	K	NIST Webbook
tb	339.10	K	NIST Webbook
tb	340.20	K	NIST Webbook
tb	339.20 ± 0.50	K	NIST Webbook
tb	339.15 ± 0.10	K	NIST Webbook
tb	338.90	K	NIST Webbook
tb	339.20	K	Experimental Isobaric Vapor-Liquid Equilibrium Data for Binary Mixtures of Cyclic Ethers with (1-Methylethyl)benzene
tb	339.12	K	Isobaric Vapor-Liquid Equilibrium for Binary System of Tetrahydrofuran + 1,4-Butanediol and gamma-Butyrolactone at 50.0 and 70.0 kPa
tb	338.65	K	Measurement and correlation of isobaric vapor-liquid equilibria of methanol + tetrahydrofuran + 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide at 101.3 kPa

tb	339.12	K	Study of isobaric vapour liquid equilibrium of some cyclic ethers with 1-chloropropane: Experimental results and SAFT-VR modelling
tb	339.35 ± 0.50	K	NIST Webbook
tb	330.15 ± 2.00	K	NIST Webbook
tb	339.05 ± 0.30	K	NIST Webbook
tc	541.00 ± 1.50	K	NIST Webbook
tc	540.10 ± 0.30	K	NIST Webbook
tc	540.20	K	NIST Webbook
tf	164.05 ± 0.40	K	NIST Webbook
tf	164.63 ± 0.20	K	NIST Webbook
tf	164.15 ± 1.00	K	NIST Webbook
tf	165.10 ± 0.50	K	NIST Webbook
tt	164.76 ± 0.02	K	NIST Webbook
tt	164.76 ± 0.02	K	NIST Webbook
tt	164.76 ± 0.05	K	NIST Webbook
vc	0.225 ± 0.003	m3/kmol	NIST Webbook
volm	8.17e-05	m3/mol	Excess Gibbs Energies of the Ternary System 2-Methoxyethanol + Tetrahydrofuran + Cyclohexane and Other Relevant Binaries at 298.15 K
volm	8.17e-05	m3/mol	Excess Gibbs energies and volumes of the ternary system chloroform + tetrahydrofuran + cyclohexane at 298.15 K

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	133.68 ± 0.27	J/mol×K	500.15	NIST Webbook
cpg	120.39 ± 0.24	J/mol×K	449.15	NIST Webbook
cpg	106.12 ± 0.21	J/mol×K	399.15	NIST Webbook
cpg	91.36 ± 0.18	J/mol×K	349.15	NIST Webbook
cpg	85.13 ± 0.17	J/mol×K	328.15	NIST Webbook
cpl	106.70	J/mol×K	210.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry

cpl	124.50	J/mol×K	305.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	123.10	J/mol×K	300.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	121.70	J/mol×K	295.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	120.30	J/mol×K	290.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	119.20	J/mol×K	285.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	120.50	J/mol×K	298.00	NIST Webbook
cpl	120.00	J/mol×K	298.15	NIST Webbook
cpl	123.90	J/mol×K	298.15	NIST Webbook
cpl	123.56	J/mol×K	298.15	NIST Webbook
cpl	122.92	J/mol×K	298.15	NIST Webbook
cpl	124.10	J/mol×K	298.15	NIST Webbook
cpl	124.10	J/mol×K	298.15	NIST Webbook
cpl	126.00	J/mol×K	310.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	135.20	J/mol×K	318.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model
cpl	128.80	J/mol×K	320.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry

cpl	133.20	J/mol×K	313.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	131.40	J/mol×K	308.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	128.50	J/mol×K	303.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	136.60	J/mol×K	323.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	126.40	J/mol×K	298.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	123.80	J/mol×K	293.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	
cpl	121.10	J/mol×K	288.15	Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 308.15K Application of an extended cell model	

cpl	115.50	J/mol×K	270.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	130.50	J/mol×K	325.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	105.50	J/mol×K	180.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	105.60	J/mol×K	185.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	105.80	J/mol×K	190.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	106.00	J/mol×K	195.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	106.30	J/mol×K	200.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	117.90	J/mol×K	280.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	116.60	J/mol×K	275.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	107.90	J/mol×K	225.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry

cpl	114.40	J/mol×K	265.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	113.40	J/mol×K	260.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	112.40	J/mol×K	255.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	111.50	J/mol×K	250.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	110.60	J/mol×K	245.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	109.80	J/mol×K	240.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	109.20	J/mol×K	235.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	108.50	J/mol×K	230.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	106.50	J/mol×K	205.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	107.60	J/mol×K	220.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry
cpl	123.90	J/mol×K	298.15	NIST Webbook

cpl	127.30	J/mol×K	315.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry	
cpl	107.20	J/mol×K	215.15	Measurement of heat capacities of ionic liquids by differential scanning calorimetry	
dvisc	0.0004540	Pa×s	298.15	Excess molar volume, viscosity, and refractive index study for the ternary mixture {2-methyl-2-butanol (1) + tetrahydrofuran (2) + propylamine (3)} at different temperatures. Application of the ERAS-model and Peng Robinson Stryjek Vera equation of state	
dvisc	0.0004276	Pa×s	308.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques	
dvisc	0.0004631	Pa×s	298.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques	

dvisc	0.0003902	Pa×s	318.15	Densities, Viscosities, and Sound Speeds of Some Acetate Salts in Binary Mixtures of Tetrahydrofuran and Methanol at (303.15, 313.15, and 323.15) K
dvisc	0.0004277	Pa×s	308.15	Densities, Viscosities, and Sound Speeds of Some Acetate Salts in Binary Mixtures of Tetrahydrofuran and Methanol at (303.15, 313.15, and 323.15) K
dvisc	0.0004630	Pa×s	298.15	Densities, Viscosities, and Sound Speeds of Some Acetate Salts in Binary Mixtures of Tetrahydrofuran and Methanol at (303.15, 313.15, and 323.15) K
dvisc	0.0004631	Pa×s	298.15	Densities and Viscosities of the Binary Mixtures of Tetrahydrofuran with Isomeric Chlorobutanes at 298.15 K and 313.15 K
dvisc	0.0003992	Pa×s	313.15	Excess Enthalpy, Excess Volume, Viscosity Deviation, and Speed of Sound Deviation for the Mixture Tetrahydrofuran + 2,2,2-Trifluoroethanol at (283.15, 298.15, and 313.15) K

dvisc	0.0004645	Pa×s	298.15	Excess Enthalpy, Excess Volume, Viscosity Deviation, and Speed of Sound Deviation for the Mixture Tetrahydrofuran + 2,2,2-Trifluoroethanol at (283.15, 298.15, and 313.15) K
dvisc	0.0005438	Pa×s	283.15	Excess Enthalpy, Excess Volume, Viscosity Deviation, and Speed of Sound Deviation for the Mixture Tetrahydrofuran + 2,2,2-Trifluoroethanol at (283.15, 298.15, and 313.15) K
dvisc	0.0004172	Pa×s	308.15	Excess molar volume, viscosity, and refractive index study for the ternary mixture {2-methyl-2-butanol (1) + tetrahydrofuran (2) + propylamine (3)} at different temperatures. Application of the ERAS-model and Peng Robinson Stryjek Vera equation of state
dvisc	0.0003903	Pa×s	318.15	Studies on Thermodynamic and Transport Properties of Binary Mixtures of Acetonitrile with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques

dvisc	0.0004991	Pa×s	288.15	Excess molar volume, viscosity, and refractive index study for the ternary mixture {2-methyl-2-butanol (1) + tetrahydrofuran (2) + propylamine (3)} at different temperatures. Application of the ERAS-model and Peng Robinson Stryjek Vera equation of state
dvisc	0.0004125	Pa×s	308.15	Ultrasonic and viscometric study of molecular interactions in binary mixtures of aniline with 1-propanol, 2-propanol, 2-methyl-1-propanol, and 2-methyl-2-propanol at different temperatures
dvisc	0.0004580	Pa×s	298.15	Ultrasonic and viscometric study of molecular interactions in binary mixtures of aniline with 1-propanol, 2-propanol, 2-methyl-1-propanol, and 2-methyl-2-propanol at different temperatures
dvisc	0.0003900	Pa×s	313.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0004200	Pa×s	303.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol

dvisc	0.0004700	Pa×s	293.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0004900	Pa×s	293.00	Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol
dvisc	0.0004260	Pa×s	308.15	Viscosity Behavior of Some Oxygen Containing Compounds
dvisc	0.0004620	Pa×s	298.15	Viscosity Behavior of Some Oxygen Containing Compounds
dvisc	0.0005080	Pa×s	288.15	Viscosity Behavior of Some Oxygen Containing Compounds
dvisc	0.0003979	Pa×s	313.15	Densities and Viscosities of the Binary Mixtures of Tetrahydrofuran with Isomeric Chlorobutanes at 298.15 K and 313.15 K
dvisc	0.0003762	Pa×s	318.15	Ultrasonic and viscometric study of molecular interactions in binary mixtures of aniline with 1-propanol, 2-propanol, 2-methyl-1-propanol, and 2-methyl-2-propanol at different temperatures
hfust	8.54	kJ/mol	164.80	NIST Webbook
hfust	8.54	kJ/mol	164.76	NIST Webbook
hfust	8.54	kJ/mol	164.80	NIST Webbook
hfust	8.54	kJ/mol	164.76	NIST Webbook
hvapt	33.10	kJ/mol	306.00	NIST Webbook
hvapt	30.80	kJ/mol	320.50	NIST Webbook
hvapt	32.50 ± 0.20	kJ/mol	287.50	NIST Webbook

hvapt	29.60	kJ/mol	504.00	NIST Webbook
hvapt	29.00	kJ/mol	439.00	NIST Webbook
hvapt	31.90	kJ/mol	334.50	NIST Webbook
hvapt	32.30	kJ/mol	314.50	NIST Webbook
hvapt	29.81	kJ/mol	339.10	NIST Webbook
hvapt	32.80	kJ/mol	290.50	NIST Webbook
hvapt	33.00	kJ/mol	293.00	NIST Webbook
hvapt	31.80	kJ/mol	303.00	NIST Webbook
kvisc	0.0000005	m2/s	313.15	Experimental and predicted viscosities of binary mixtures of cyclic ethers with 1-chloropentane or 1-chlorohexane at 283.15, 298.15, and 313.15K
kvisc	0.0000005	m2/s	298.15	Experimental and predicted viscosities of binary mixtures of cyclic ethers with 1-chloropentane or 1-chlorohexane at 283.15, 298.15, and 313.15K
kvisc	0.0000006	m2/s	283.15	Experimental and predicted viscosities of binary mixtures of cyclic ethers with 1-chloropentane or 1-chlorohexane at 283.15, 298.15, and 313.15K
pvap	11.68	kPa	284.88	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	21.60	kPa	298.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	17.25	kPa	293.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	13.67	kPa	288.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	10.73	kPa	283.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	101.30	kPa	338.65	Measurement and correlation of isobaric vapor-liquid equilibria of methanol + tetrahydrofuran + 1-alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide at 101.3 kPa
pvap	70.61	kPa	328.15	Isothermal vapour-liquid equilibrium for cyclic ethers with 1-chloropentane

pvap	40.62	kPa	313.15	Isothermal vapour-liquid equilibrium for cyclic ethers with 1-chloropentane
pvap	21.61	kPa	298.15	Isothermal vapour-liquid equilibrium for cyclic ethers with 1-chloropentane
pvap	8.61	kPa	279.41	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	26.81	kPa	303.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	32.94	kPa	308.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	40.21	kPa	313.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures

pvap	48.68	kPa	318.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	58.54	kPa	323.15	Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + 1,1,2,2-tetrachloroethane or tetrachloroethene) at nine temperatures
pvap	13.79	kPa	288.50	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	21.61	kPa	298.15	Isothermal (vapour + liquid) equilibrium of (cyclic ethers + chlorohexane) mixtures: Experimental results and SAFT modelling
pvap	14.00	kPa	289.00	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	14.48	kPa	289.66	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	15.65	kPa	291.04	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	16.65	kPa	292.31	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	17.43	kPa	293.30	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	19.72	kPa	296.05	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	20.37	kPa	296.84	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	20.65	kPa	297.09	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	21.50	kPa	298.15	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	22.24	kPa	298.72	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	70.61	kPa	328.15	Isothermal (vapour + liquid) equilibrium of (cyclic ethers + chlorohexane) mixtures: Experimental results and SAFT modelling
pvap	25.28	kPa	301.73	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	26.87	kPa	303.23	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	27.79	kPa	304.04	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	7.10	kPa	275.99	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	7.56	kPa	276.53	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	7.84	kPa	277.76	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	8.33	kPa	278.19	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	12.44	kPa	286.55	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	9.04	kPa	279.86	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	8.93	kPa	280.16	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	9.96	kPa	281.95	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	10.55	kPa	283.07	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	10.32	kPa	283.15	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	10.90	kPa	283.36	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	29.43	kPa	305.35	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	30.77	kPa	306.54	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	32.49	kPa	308.07	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	35.55	kPa	310.20	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	36.65	kPa	310.86	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	38.54	kPa	311.90	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	40.13	kPa	313.12	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	40.09	kPa	313.18	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	44.65	kPa	315.86	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	47.86	kPa	317.71	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	51.31	kPa	319.58	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	55.34	kPa	321.57	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	59.46	kPa	323.59	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

pvap	6.59	kPa	273.60	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	8.38	kPa	278.30	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	11.81	kPa	285.02	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	13.02	kPa	286.97	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	14.89	kPa	289.75	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	

pvap	17.00	kPa	292.65	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	21.27	kPa	297.57	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	26.67	kPa	302.75	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	32.88	kPa	307.83	Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols	
pvap	26.67	kPa	303.04	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran	
pvap	26.68	kPa	303.05	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran	

pvap	29.09	kPa	305.11	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran
pvap	34.50	kPa	309.28	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran
pvap	40.15	kPa	313.11	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran
pvap	45.51	kPa	316.36	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran
pvap	48.62	kPa	318.10	Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane High-Boiling Diastereomer + Tetrahydrofuran
pvap	23.84	kPa	300.57	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran
pvap	26.62	kPa	302.96	Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran

rfi	1.40720	298.15	Bubble temperature measurements on the binary mixtures formed by decane with a variety of compounds at 95.8 kPa
rfi	1.39510	318.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.39750	313.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.39990	308.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.40230	303.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.40460	298.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.40700	293.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K

rfi	1.40940	288.15	Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K
rfi	1.40500	298.15	Viscosity, Density, Speed of Sound, and Refractive Index of Binary Mixtures of Organic Solvent + Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate at 298.15 K
rfi	1.40520	298.15	Acoustic and Volumetric Properties of Binary Mixtures of Ionic Liquid 1-Butyl-3-methylimidazolium Bis(trifluoromethylsulfonyl)imide with Acetonitrile and Tetrahydrofuran
rfi	1.40610	298.15	Vapor Liquid Equilibrium for Ternary and Binary Mixtures of Tetrahydrofuran, Cyclohexane, and 1,2-Propanediol at 101.3 kPa
rfi	1.40500	293.15	Solubilities of Phosphorus-Containing Compounds in Selected Solvents
rfi	1.40470	298.15	Solubilities of Bis (2,2,6,6-Tetramethyl-4-Piperidinyl) Maleate in Hexane, Heptane, Octane, m-Xylene and Tetrahydrofuran from (253.15 to 310.15) K
rfi	1.40480	303.15	Thermodynamic Properties of Water + Tetrahydrofuran and Water + 1,4-Dioxane Mixtures at (303.15, 313.15, and 323.15) K

rfi	1.40260	303.15	Density, Viscosity, Refractive Index, and Speed of Sound in the Binary Mixtures of Tri-n-butylamine + Triethylamine, + Tetrahydrofuran, + Tetradecane, + Tetrachloroethylene, + Pyridine, or + Trichloroethylene at (298.15, 303.15, and 308.15) K
rfi	1.39980	308.15	Density, Viscosity, Refractive Index, and Speed of Sound in the Binary Mixtures of Tri-n-butylamine + Triethylamine, + Tetrahydrofuran, + Tetradecane, + Tetrachloroethylene, + Pyridine, or + Trichloroethylene at (298.15, 303.15, and 308.15) K
rfi	1.40490	298.15	Density, Viscosity, Refractive Index, and Speed of Sound in the Binary Mixtures of Tri-n-butylamine + Triethylamine, + Tetrahydrofuran, + Tetradecane, + Tetrachloroethylene, + Pyridine, or + Trichloroethylene at (298.15, 303.15, and 308.15) K

rfi	1.40720		293.15	Bubble-Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Amyl Alcohol with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa
rfi	1.40280		303.15	Densities, speeds of sound, isentropic compressibilities, refractive indexes, and viscosities of tetrahydrofuran with haloalkane or alkyl ethanoate at T = 303.15 K
rfi	1.40480		298.15	To the issue of temperature-dependent behavior of standard molar volumes of components in the binary system (water + tetrahydrofuran) at ambient pressure
rhol	882.00	kg/m3	298.15	Vapour liquid equilibrium of cyclic ethers with 1-chlorohexane: Experimental results and UNIFAC predictions
rhol	878.00	kg/m3	303.00	Densities, ultrasonic speeds and refractive indices of phenetole with N-methyl-2-pyrrolidone, N,N-dimethylformamide and tetrahydrofuran binary mixtures at different temperatures

rhol	882.00	kg/m3	298.00	Densities, ultrasonic speeds and refractive indices of phenetole with N-methyl-2-pyrrolidone, N,N-dimethylformamide and tetrahydrofuran binary mixtures at different temperatures
rhol	855.46	kg/m3	323.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF
rhol	874.00	kg/m3	308.00	Densities, ultrasonic speeds and refractive indices of phenetole with N-methyl-2-pyrrolidone, N,N-dimethylformamide and tetrahydrofuran binary mixtures at different temperatures
rhol	887.01	kg/m3	293.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures
rhol	882.05	kg/m3	298.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures

rhol	872.14	kg/m3	308.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures
rhol	867.19	kg/m3	313.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures
rhol	877.10	kg/m3	303.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures
rhol	882.00	kg/m3	298.15	Volumetric Properties of the Ionic Liquid, 1-Butyl-3-methylimidazolium Tetrafluoroborate, in Organic Solvents at T = 298.15 K
rhol	882.09	kg/m3	298.15	Excess Enthalpies of the Ternary Mixtures: Tetrahydrofuran + (Hexane or Cyclohexane) + Decane at 298.15K

rhol	882.06	kg/m3	298.15	Excess Molar volumes and Surface Tensions of Trimethylbenzene with Tetrahydrofuran Tetrachloromethane and Dimethylsulfoxide at 298.15 K
rhol	882.10	kg/m3	298.15	Study of the Surface Tension of Chlorocyclohexane or Bromocyclohexane with Some Cyclic Ethers
rhol	887.41	kg/m3	293.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	881.98	kg/m3	298.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	876.47	kg/m3	303.13	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	870.96	kg/m3	308.15	Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents
rhol	882.08	kg/m3	298.15	Excess Molar Enthalpies of Dipropyl Ether + Dibutyl Ether + (1-Hexene or Tetrahydrofuran) at 298.15 K

rhol	881.50	kg/m3	298.15	Densities, Viscosities, and Sound Speed of Binary Mixtures of Hexyl Acetate with Tetrahydrofuran, 1,4-Dioxane, Anisole, and Butyl Vinyl Ether	
rhol	878.70	kg/m3	303.15	Densities, Viscosities, and Sound Speed of Binary Mixtures of Hexyl Acetate with Tetrahydrofuran, 1,4-Dioxane, Anisole, and Butyl Vinyl Ether	
rhol	873.00	kg/m3	308.15	Densities, Viscosities, and Sound Speed of Binary Mixtures of Hexyl Acetate with Tetrahydrofuran, 1,4-Dioxane, Anisole, and Butyl Vinyl Ether	
rhol	866.90	kg/m3	313.15	Densities, Viscosities, and Sound Speed of Binary Mixtures of Hexyl Acetate with Tetrahydrofuran, 1,4-Dioxane, Anisole, and Butyl Vinyl Ether	
rhol	887.52	kg/m3	293.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory	
rhol	861.11	kg/m3	318.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF	

rhol	866.71	kg/m3	313.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF
rhol	872.28	kg/m3	308.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF
rhol	877.81	kg/m3	303.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF
rhol	883.30	kg/m3	298.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF
rhol	888.77	kg/m3	293.15	Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III [Cnmim][PF6] (n = 4, 6, 8) + THF

rhol	859.96	kg/m3	318.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	871.16	kg/m3	308.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	882.21	kg/m3	298.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	855.29	kg/m3	323.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers

rhol	860.96	kg/m3	318.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	866.58	kg/m3	313.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	872.16	kg/m3	308.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers
rhol	877.70	kg/m3	303.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers

rhol	882.21	kg/m3	298.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers	
rhol	888.69	kg/m3	293.15	Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents I. 1-Hexyl-3-methylimidazlouim tetraflouroborate and 1-octyl-3-methylimidazlouim tetraflouroborate with cyclic ethers	
rhol	882.15	kg/m3	298.15	Densities, speeds of sound and viscosities of binary mixtures of tetrahydrofuran with 1-hexanol, 1-octanol and 1-decanol at T = (298.15 to 313.15) K	
rhol	866.23	kg/m3	313.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K	
rhol	872.02	kg/m3	308.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K	

rhol	877.60	kg/m3	303.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	883.24	kg/m3	298.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	888.73	kg/m3	293.15	Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K
rhol	859.84	kg/m3	318.15	Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it? Discussion on the paper "Volumetric properties on the (tetrahydrofuran + water) and (tetra-n-butylammonium bromide + water) systems: Experimental measurements and correlations" by Veronica Belandria, Ammir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382 1386]

rhol	871.01	kg/m3	308.15	Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it? Discussion on the paper "Volumetric properties on the (tetrahydrofuran + water) and (tetra-n-butylammonium bromide + water) systems: Experimental measurements and correlations" by Veronica Belandria, Ammir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382 1386]
rhol	882.04	kg/m3	298.15	Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it? Discussion on the paper "Volumetric properties on the (tetrahydrofuran + water) and (tetra-n-butylammonium bromide + water) systems: Experimental measurements and correlations" by Veronica Belandria, Ammir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382 1386]

rhol	892.94	kg/m3	288.15	Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it? Discussion on the paper "Volumetric properties on the (tetrahydrofuran + water) and (tetra-n-butylammonium bromide + water) systems: Experimental measurements and correlations" by Veronica Belandria, Ammir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382 1386]
rhol	903.71	kg/m3	278.15	Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it? Discussion on the paper "Volumetric properties on the (tetrahydrofuran + water) and (tetra-n-butylammonium bromide + water) systems: Experimental measurements and correlations" by Veronica Belandria, Ammir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382 1386]
rhol	876.54	kg/m3	303.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory

rhol	882.04	kg/m3	298.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory
rhol	881.95	kg/m3	298.15	Surface Tension of Mixtures of Tetrahydrofuran or Tetrahydropyran with Isomeric Chlorobutanes
rhol	892.86	kg/m3	288.15	Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine Flory Patterson theory
rhol	882.09	kg/m3	298.15	Excess molar enthalpies of the ternary mixtures (1-hexene + tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether) at the temperature 298.15K.
rhol	882.00	kg/m3	298.15	Volumetric and compressibility behaviour of ionic liquid, 1-n-butyl-3-methylimidazolium hexafluorophosphate and tetrabutylammonium hexafluorophosphate in organic solvents at T = 298.15 K
rhol	882.09	kg/m3	298.15	Excess molar enthalpies of the ternary mixtures: (tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether + n-octane) at the temperature 298.15 K

rhol	882.09	kg/m3	298.15	Excess molar enthalpies of the ternary mixtures: {tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether + n-dodecane} at the temperature 298.15 K
rhol	882.01	kg/m3	298.15	Standard partial molar volumes of some electrolytes in ethylene carbonate based mixtures
rhol	876.50	kg/m3	303.15	Unravelling various types of non-covalent interactions of benzyl amine with ethers in n-hexane at 303.15 K by ultrasonic and DFT methods
rhol	884.99	kg/m3	298.15	Quaternary isothermal vapor-liquid equilibrium of the model biofuel 2-butanone + n-heptane + tetrahydrofuran + cyclohexane using Raman spectroscopic characterization
rhol	854.56	kg/m3	323.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination

rhol	860.21	kg/m3	318.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination
rhol	865.82	kg/m3	313.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination
rhol	871.38	kg/m3	308.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination

rhol	876.91	kg/m3	303.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination
rhol	882.40	kg/m3	298.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination
rhol	887.87	kg/m3	293.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination

rhol	893.29	kg/m3	288.15	Volumetric and viscometric behavior of the binary systems ethyl lactate + 1,2- propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination
rhol	862.23	kg/m3	318.15	Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures
sfust	51.80	J/mol×K	164.76	NIST Webbook
sfust	51.83	J/mol×K	164.76	NIST Webbook
speedsl	1207.00	m/s	313.15	Thermodynamic and acoustic properties of binary mixtures of ethers. V. Diisopropyl ether or oxolane with 2- or 3-chloroanilines at 303.15, 313.15 and 323.15 K
speedsl	1229.61	m/s	308.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1277.37	m/s	298.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures

speedsl	1325.78	m/s	288.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1374.98	m/s	278.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1280.20	m/s	298.15	Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T) 298.15 K
speedsl	1165.00	m/s	323.15	Thermodynamic and acoustic properties of binary mixtures of ethers. V. Diisopropyl ether or oxolane with 2- or 3-chloroanilines at 303.15, 313.15 and 323.15 K
speedsl	1182.37	m/s	318.15	Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures
speedsl	1254.00	m/s	303.15	Thermodynamic and acoustic properties of binary mixtures of ethers. V. Diisopropyl ether or oxolane with 2- or 3-chloroanilines at 303.15, 313.15 and 323.15 K

speedsl	1165.00	m/s	323.15	Thermodynamic and acoustic properties of binary mixtures of ethers. IV. Diisopropyl ether or oxolane with N,N-dimethylaniline or N,Ndiethylaniline at 303.15, 313.15 and 323.15 K
speedsl	1207.00	m/s	313.15	Thermodynamic and acoustic properties of binary mixtures of ethers. IV. Diisopropyl ether or oxolane with N,N-dimethylaniline or N,Ndiethylaniline at 303.15, 313.15 and 323.15 K
speedsl	1254.00	m/s	303.15	Thermodynamic and acoustic properties of binary mixtures of ethers. IV. Diisopropyl ether or oxolane with N,N-dimethylaniline or N,Ndiethylaniline at 303.15, 313.15 and 323.15 K
speedsl	1165.00	m/s	323.15	Thermodynamic and acoustic properties of binary mixtures of ethers. III. Diisopropyl ether or oxolane with o- or m-toluidines at 303.15, 313.15 and 323.15 K
speedsl	1256.00	m/s	303.15	Thermodynamic and acoustic properties of binary mixtures of oxolane with aniline and substituted anilines at 303.15, 313.15 and 323.15 K

speedsl	1254.00	m/s	303.15	Thermodynamic and acoustic properties of binary mixtures of ethers. III. Diisopropyl ether or oxolane with o- or m-toluidines at 303.15, 313.15 and 323.15 K	
speedsl	1171.00	m/s	323.15	Thermodynamic and acoustic properties of binary mixtures of oxolane with aniline and substituted anilines at 303.15, 313.15 and 323.15 K	
speedsl	1212.00	m/s	313.15	Thermodynamic and acoustic properties of binary mixtures of oxolane with aniline and substituted anilines at 303.15, 313.15 and 323.15 K	
speedsl	1207.00	m/s	313.15	Thermodynamic and acoustic properties of binary mixtures of ethers. III. Diisopropyl ether or oxolane with o- or m-toluidines at 303.15, 313.15 and 323.15 K	
srf	0.03	N/m	293.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide	
srf	0.03	N/m	308.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane	

srf	0.03	N/m	298.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane
srf	0.03	N/m	288.15	Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydofuran, 2-Propanol, and 2,2,4-Trimethylpentane
srf	0.03	N/m	308.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
srf	0.03	N/m	298.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile
srf	0.03	N/m	293.15	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile

srf	0.03	N/m	308.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	303.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	298.15	Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide
srf	0.03	N/m	303.13	Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile

Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbp	339.15	K	96.60	Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/T + C*ln(T) + D*T^2
Coeff. A	7.83577e+01
Coeff. B	-6.33998e+03
Coeff. C	-9.60880e+00
Coeff. D	8.18348e-06
Temperature range (K), min.	164.65
Temperature range (K), max.	540.15

Datasets

Mass density, kg/m3

Pressure, kPa - Liquid	Temperature, K - Liquid	Mass density, kg/m3 - Liquid
85.90	298.15	882.02

Reference *https://www.doi.org/10.1016/j.fluid.2013.05.001*

Viscosity, Pa*s

Temperature, K - Liquid	Pressure, kPa - Liquid	Viscosity, Pa*s - Liquid
333.20	10200.00	0.0003650

Reference *https://www.doi.org/10.1007/s10765-008-0542-6*

Temperature, K	Pressure, kPa	Viscosity, Pa*s
303.15	101.33	0.0003260

Reference *https://www.doi.org/10.1016/j.tca.2009.07.008*

Sources

Thermodynamics and activity coefficients at infinite dilution for organic compounds and water in the ionic liquid 1-butyl-3-methylimidazolium perchlorate:	https://www.doi.org/10.1016/j.jct.2017.07.012
Liquid liquid equilibria and partitioning in inorganic-aqueous systems:	https://www.doi.org/10.1016/j.fluid.2007.01.003
Measurement and correlation of the solubility of 2,3,4,5-tetrabromothiophene in	https://www.doi.org/10.1016/j.jct.2012.06.005
Measurements of activity coefficients at infinite dilution for organic solutes in two quaternary ammonium-based ionic liquids [DDA][...] and [DDA][BF4]:	https://www.doi.org/10.1016/j.fluid.2018.11.011
Vapor Pressure and Its Temperature Dependence of 28 Organic Compounds: Cyclic Amines, Cyclic Ethers, and Cyclic and Open Chain Secondary Alcohols:	https://www.doi.org/10.1021/acs.jced.6b00576
... determination of the solubility of ... mixture ... from T = (278.15 to 323.15) K:	https://www.doi.org/10.1016/j.jct.2015.11.005
Solubility and solution thermodynamics of 2,3,4,5-tetrabromothiophene in (ethanol + tetrahydrofuran) binary solvent mixtures:	https://www.doi.org/10.1016/j.jct.2013.05.023
Isothermal (vapour + liquid) equilibrium of (cyclic ethers + chlorohexane):	https://www.doi.org/10.1016/j.jct.2008.03.017
Solubility of Organic Systems Containing 1,4-Dioxan-2-one:	https://www.doi.org/10.1021/je050406x
High selective water/butan-1-ol separation on investigation of limiting activity coefficients with [P8,8,8,8][NTf2] ionic liquid:	https://www.doi.org/10.1016/j.fluid.2017.06.001
Joback Method:	https://en.wikipedia.org/wiki/Joback_method
Activity coefficients at infinite dilution of organic solutes in 1-hexyl-3-methylimidazolium trifluoroacetate and ... using gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2013.10.038
Solubility of oxcarbazepine in eight solvents within the temperature range ...:	https://www.doi.org/10.1016/j.jct.2016.09.011
Solubility of Florfenicol in Different Solvents at Temperatures from (278 to 318) K:	https://www.doi.org/10.1021/je1008284
Ultrasonic and Volumetric Properties of 1-Ethyl-3-methylimidazolium Trifluoromethanesulfonate Ionic Liquid with 2-Propanol or Tetrahydrofuran at Several Temperatures:	https://www.doi.org/10.1021/je200609g
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid trihexyltetradecylphosphonium ...:	https://www.doi.org/10.1016/j.jct.2018.01.003
Standard partial molar volumes of some electrolytes in ethylene ...:	https://www.doi.org/10.1016/j.jct.2003.09.005
Densities and ultrasonic studies for binary mixtures of tetrahydrofuran with chlorobenzenes, chlorotoluenes and nitrotoluenes at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2011.07.018
Acoustic and Volumetric Properties of Binary Mixtures of Ionic Liquid 1-Butyl-3-methylimidazolium ...:	https://www.doi.org/10.1021/je400781b
Isothermal (vapour + liquid) equilibrium for binary mixtures of (tetrahydrofuran + ...) and tetrachloroethylene at nine temperatures:	https://www.doi.org/10.1016/j.jct.2005.06.003
Solubility of Chloroquine Diphosphate and 4,7-Dichloroquinoline in Water, Ethanol, Tetrahydrofuran, Acetonitrile, and Acetone from (298.2 to 333.2) K:	https://www.doi.org/10.1021/je8007099
Abraham model correlations for describing the thermodynamic properties of solute transfer into pentyl acetate ...:	https://www.doi.org/10.1016/j.jct.2018.05.003
Thermodynamic and acoustic properties of binary mixtures of ethers. ...:	https://www.doi.org/10.1016/j.tca.2011.01.019
Experimental and theoretical excess molar properties of imidazolium based ionic liquids with molecular organic solvents – I. 1-Hexyl-3-methylimidazolium ...:	https://www.doi.org/10.1016/j.jct.2013.11.027
Thermodynamic and acoustic properties of binary mixtures of ethers. ... 303.15, 313.15 and 323.15 K:	https://www.doi.org/10.1016/j.tca.2011.07.005
Solubility of Trimethoprim (TMP) in Different Organic Solvents from (278 to 333) K:	https://www.doi.org/10.1021/je700497h
... properties of the (tetrahydrofuran + water) and (... tetra-n-butyl ammonium bromide + water) systems: Experimental measurements and correlations:	https://www.doi.org/10.1016/j.jct.2009.06.014
Solubilities of Phosphoramidic Acid, N-(phenylmethyl)-, Diphenyl Ester in ...:	https://www.doi.org/10.1021/acs.jced.5b00007
Solubility of Naphthalene in Isobutyl Acetate, n-Butyric Acid, Ethyl Acetate, N-Methyl Pyrrolidone, N,N-Dimethylformamide and ...:	https://www.doi.org/10.1021/je800428r
Studies on Thermodynamic and Transport Properties of Binary Mixtures of ... with Some Cyclic Ethers at Different Temperatures by Volumetric, Viscometric, and Interferometric Techniques:	https://www.doi.org/10.1021/je060311a
Vapour-liquid equilibria ... tetrahydrofuran + ethanol system:	https://www.doi.org/10.1016/j.fluid.2014.03.027
Separation of aliphatic from aromatic hydrocarbons and sulphur compounds from fuel based on measurements of activity coefficients at infinite dilution ...:	https://www.doi.org/10.1016/j.jct.2016.07.017
... determination of solid liquid equilibrium ...:	https://www.doi.org/10.1016/j.jct.2016.12.002
... binary mixtures based on limiting activity coefficients data using ...:	https://www.doi.org/10.1016/j.fluid.2017.12.029
Thermodynamics and activity coefficients at infinite dilution ... N,N-diethyl-N-methyl-N-(2-methoxy-ethyl)ammonium ...:	https://www.doi.org/10.1016/j.fluid.2018.06.013
... binary mixtures containing ... and ethers at atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2004.11.006
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2011.11.021
Solubility Measurement, Correlation, and Molecular Interactions of 3-Methyl-6-nitroindazole in Different Neat Solvents and Mixed Solvents from T = 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.8b01256

Activity coefficients at infinite dilution of organic solutes in 1-butyl-3-methylimidazolium nitrate using gas-liquid chromatography:	https://www.doi.org/10.1016/j.fluid.2012.05.006
Liquid-Liquid Equilibria for the Ternary Systems of Perfluorohexane or Perfluoroamine + Hydrofluoroether + Tetrahydrofuran at 298.15 K or 273.15 K:	https://www.doi.org/10.1021/je400227m
Solubility of 4-(3,4-Dichlorophenyl)-1-tetralone in Some Organic Solvents:	https://www.doi.org/10.1021/je060329l
New ionic liquid [P4,4,4,4][NTf2] in bio-butanol extraction on investigation of limiting activity coefficients:	https://www.doi.org/10.1016/j.fluid.2018.07.028
Vapour liquid equilibrium of cyclic ethers with 1-chlorohexane: Experimental results and UNIFAC predictions:	https://www.doi.org/10.1016/j.fluid.2007.05.013
Densities and Viscosities of the Binary Mixtures of Tetrahydrofuran with Isomeric Chlorobutanes at 298.15 K and 313.15 K:	https://www.doi.org/10.1021/je0600653
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(2-methoxyethyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.05.017
Systems with ionic liquids: Measurement of VLE and c1 data and prediction of their thermodynamic behavior using original UNIFAC, mod. UNIFAC(Do) and COSMO-RS(Ol):	https://www.doi.org/10.1016/j.jct.2005.04.010
Surface Tension of Mixtures of Tetrahydrofuran or Tetrahydropyran with Isomeric Chlorobutanes:	https://www.doi.org/10.1007/s10765-007-0169-z
High-pressure vapor-liquid equilibria for carbon dioxide + tetrahydrofuran mixtures:	https://www.doi.org/10.1016/j.fluid.2010.04.018
Determination and correlation of solubility and solution thermodynamics of musk xylene in different pure solvents:	https://www.doi.org/10.1016/j.jct.2019.03.029
Solubility, Solubility Modeling, and Antisolvent Precipitation of 1,3-Bis(9-carbazolyl)benzene in Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00432
Phase equilibrium and thermophysical properties of mixtures containing a cyclic ether and 1-chloropropane:	https://www.doi.org/10.1016/j.fluid.2010.04.017
Solid-Liquid Equilibrium of Azacyclotridecan-2-one in 15 Pure Solvents from T = 273.15 to 323.15 K: Experimental Determination and Thermodynamic Modeling:	https://www.doi.org/10.1021/acs.jced.8b01193
Excess molar volume and viscosity study for the ternary system tetrahydrofuran (1) + 1-chlorobutane (2) + 2-butanol (3) at 283.15, 298.15 and 313.15 K:	https://www.doi.org/10.1016/j.fluid.2005.11.018
Thermodynamic and Spectroscopic Study of the Ternary System Dimethyladipate + Tetrahydrofuran + 1-Butanol at T = (288.15 to 323.15) K:	https://www.doi.org/10.1021/je4003916
Solubilities of 4',5,7-Triacetoxyflavanone in Fourteen Organic Solvents at Different Temperatures:	https://www.doi.org/10.1021/acs.jced.6b00929
Excess Molar Volumes and Surface Tensions of Trimethylbenzene with Tetrahydrofuran Tetrachloromethane and Dimethylsulfoxide at 298.15 K:	https://www.doi.org/10.1021/je0497294
Interactions of volatile organic compounds with the ionic liquids 1-butyl-1-methylpyrrolidinium tetracyanoborate and 1-butyl-1-methylpyrrolidinium bis(oxalato)borate:	https://www.doi.org/10.1016/j.jct.2012.09.017
Phase Equilibria of (1-Ethyl-3-methylimidazolium Ethylsulfate + Hydrocarbon, Ketone, and Ether) Binary Systems:	https://www.doi.org/10.1021/je700591h
Determination of the thermodynamic parameters of ionic liquid 1-propyl-3-methylimidazolium bromide by gas-liquid chromatography:	https://www.doi.org/10.1016/j.jct.2018.09.023
Solubility Measurement and Correlation of 2-Oxindole in 12 Pure Organic Solvents:	https://www.doi.org/10.1021/acs.jced.9b00308
Phenomenon of "negative partial molar expansibility" of water in tetrahydrofuran: How plausible is it?	https://www.doi.org/10.1016/j.jct.2010.08.006
The study of activity coefficients at infinite dilution for organic solutes and water in 1-butyl-4-methylpyridinium [B4MPy][DCA] using GLC	https://www.doi.org/10.1016/j.jct.2013.09.007
Discussion on the paper "Volumetric properties of the (tetrahydrofuran + water and tetra-n-butylammonium chloride + water) systems: Experimental measurements and correlations" by Vanessa W. Belandria, Amir H. Mohammadi and Dominique Richon [J. Chem. Thermodyn. 41 (2009) 1382-1386]	https://www.doi.org/10.1016/j.jct.2017.11.017
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on gamma infinity data measurements	https://www.doi.org/10.1016/j.jct.2013.08.030
Thermodynamics and activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylmorpholinium tricyanomethanide	https://www.doi.org/10.1021/je700289p
Tetrachloromethane + Toluene + 1-Butanol Mixture at 25 deg C and Atmospheric Pressure:	https://www.doi.org/10.1021/je800056h
Clarification of the volumetric properties of the (tetrahydrofuran + water) systems [J. Chem. Thermodyn. 41 (2009) 1382-1386]: Author's statement	https://www.doi.org/10.1016/j.jct.2012.08.033
Determination of the solubility, dissolution enthalpy and entropy of ... in different solvents:	https://www.doi.org/10.1016/j.fluid.2012.01.019
Apparent molar volumes and apparent molar heat capacities of aqueous tetrahydrofuran, dimethyl sulfoxide, 1,4-dioxane, and 1,2-dimethoxyethane at temperatures from 278.15 K to 393.15 K	https://www.doi.org/10.1016/j.jct.2007.06.010
Densities, Viscosities, and Sound Speeds of Some Acetate Salts in Binary Mixtures of Tetrahydrofuran and Methanol at (303.15, 313.15 and 323.15) K	https://www.doi.org/10.1021/je060113j
Isothermal vapour-liquid equilibrium for cyclic ethers with 1-chloropentane:	https://www.doi.org/10.1016/j.fluid.2006.10.024
Volumetric Properties of the Ionic Liquid, 1-Butyl-3-methylimidazolium Tetrafluoroborate, in Organic Solvents at T = 298.15 K:	https://www.doi.org/10.1007/s10765-008-0395-z
Refractive Indices of Binary Mixtures of Tetrahydrofuran with Aromatic Hydrocarbon at Temperatures from (288.15 to 318.15) K:	https://www.doi.org/10.1021/je700564c
Solid-Liquid Phase Equilibrium of N,N'-Piperazinebis(neopentylglycol)phosphoramidate in Selected Solvents from T = (293.15 to 333.15) K:	https://www.doi.org/10.1021/acs.jced.7b00206
Activity coefficients at infinite dilution measurements for organic solutes and water in the 1-hexyloxymethyl-3-methyl-imidazolium and 1,3-dihexyloxymethyl-imidazolium bis(trifluoromethylsulfonyl)-imide ionic liquids: The cation influence.	https://www.doi.org/10.1016/j.fluid.2009.08.017
Compressibility Studies of Binary Solutions Involving Water as a Solute in Nonaqueous Solvents at T = 298.15 K:	https://www.doi.org/10.1021/je0601098
Activity coefficients at infinite dilution for organic solutes and water in 1-ethyl-1-methylpyrrolidinium lactate:	https://www.doi.org/10.1016/j.jct.2015.05.014
Measurement and Correlation of Solubility of Marbofloxacin in 12 Pure Solvents from 283.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00490
McGowan Method:	http://link.springer.com/article/10.1007/BF02311772
Solubility of 5-Amino Salicylic Acid in Different Solvents at Various Temperatures:	https://www.doi.org/10.1021/je900646u
Thermodynamics and activity coefficients at infinite dilution for organic compounds in the ionic liquid 1-hexyl-3-methylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2018.08.028

Separation of binary mixtures based on gamma infinity data using [OMMIM][BF4] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2018.09.003
Volumetric properties of binary mixtures of ethers and acetonitrile: Experimental results and application of the Prigogine-Flory-Patterson theory:	https://www.doi.org/10.1016/j.jct.2007.09.007
Measurements of activity coefficients at infinite dilution for organic solutes and water in N-hexylisoquinolinium thiocyanate, [HiQuin][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2013.02.004
Excess Enthalpy, Excess Volume, Viscosity Deviation, and Speed of Sound Deviation for the Mixture Tetrahydrofuran + 2,2,2-Trifluoroethanol at (283.15, 298.15, and 313.15) K:	https://www.doi.org/10.1021/je034076x
Thermodynamic and acoustic properties of binary mixtures of ethers. V. Dipropyl ether or 1,3-dioxolane with N,N-dimethylaniline or N-methylaniline at 303.15, 318.15 and 333.15 K:	https://www.doi.org/10.1016/j.tca.2011.02.004
Synthesis and Solubility of 5,5-Dimethyl-2-(phenyl(phenylamino)methyl)-1,3,2-dioxaphosphinane 2-oxide in Selected Solvents between 278.15 K and 347.15 K:	https://www.doi.org/10.1021/acs.jced.7b00585
Dortmund Data Bank Vapor-Liquid Equilibrium Data:	http://www.ddbst.com/en/EED/VLE/VLE%20Acetonitrile%3BTetrahydrofuran.php
Measurements of Activity Coefficients at Infinite Dilution for Organic Solutes and Water in the Ionic Liquid 1-Butyl-1-methylpiperidinium Thiocyanate:	https://www.doi.org/10.1021/je101008y
Ultrasonic and viscometric study of molecular interactions in binary mixtures of aniline with 1-propanol, 2-propanol, 2-methyl-1-propanol, and 2-methyl-2-propanol at different temperatures:	https://www.doi.org/10.1016/j.fluid.2007.07.016
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-hydroxyethyl)-3-methylimidazolium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2013.05.008
Activity coefficients at infinite dilution of organic solvents and water in 1-butyl-3-methylimidazolium dicyanamide. A literature review of hexane/hex-1-ene separation:	https://www.doi.org/10.1016/j.fluid.2016.02.004
Activity Coefficients at Infinite Dilution of Polar Solutes in [illegible] Trifluoromethanesulfonate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je700560s
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.jct.2012.03.005
Isothermal compressibility and internal pressure studies of some nonelectrolytes in aqueous solutions at low temperatures:	https://www.doi.org/10.1016/j.jct.2006.08.007
Study of Ether-, Alcohol-, or Cyano-Functionalized Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je900838a
Measurements of activity coefficients at infinite dilution of organic compounds and water in isoquinolinium-based ionic liquid [C8iQuin][NTf2] using GLC:	https://www.doi.org/10.1016/j.jct.2010.10.026
Thermodynamics and activity coefficients at infinite dilution for organic solutes, water and diols in the ionic liquid choline bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2014.04.024
Solubility of 2-(4-Ethylbenzoyl)benzoic Acid in Eleven Organic Solvents between 279.85 K and 343.05 K:	https://www.doi.org/10.1021/je700426k
Viscosity, Density, Speed of Sound, and Refractive Index of Binary Mixtures of Organic Solvent + Ionic Liquid, 1-Butyl-3-methylimidazolium Hexafluorophosphate at 298.15 K:	https://www.doi.org/10.1021/je700338t
Temperature dependence of limiting activity coefficients and Henry s law constants for cyclic and open-chain ethers in water:	https://www.doi.org/10.1016/j.fluid.2007.08.013
Activity Coefficients at Infinite Dilution of Organic Solutes in 1-Ethyl-3-methylimidazolium Tetrafluoroborate Using Gas Liquid Chromatography:	https://www.doi.org/10.1021/je800218g
Experimental and theoretical study on infinite dilution activity coefficients of various solutes in piperidinium ionic liquids:	https://www.doi.org/10.1016/j.jct.2013.01.005
Volumetric and viscosimetric behavior of the binary systems ethyl lactate + 1,2-propanediol, + 1,3-propanediol, + tetrahydrofuran and + tetraethylene glycol dimethyl ether. New UNIFAC-VISCO and ASOG-VISCO parameters determination:	https://www.doi.org/10.1016/j.fluid.2014.04.002
Activity coefficients at infinite dilution of organic solutes in the ionic liquid [illegible] methyl sulfate:	https://www.doi.org/10.1016/j.jct.2014.04.020
Thermophysical properties of dimethyl sulfoxide + cyclic and linear ethers at 298.15 K. Application of an extended cell model:	https://www.doi.org/10.1016/j.tca.2006.05.010
Novel phase diagrams of aqueous two-phase systems based on tetrahydrofuran + carbohydrates + water: equilibrium data and partitioning experiments:	https://www.doi.org/10.1016/j.fluid.2016.11.001
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium trifluorotris(perfluoroethyl)phosphate:	https://www.doi.org/10.1016/j.jct.2012.08.016
Determination and Correlation of Solubilities of 1,3,5-Trifluoro-2,4,6-triiodobenzene in Different Solvents from 289.75 to 332.15 K:	https://www.doi.org/10.1021/acs.jced.9b00381
Guaiacol and its mixtures: new data and predictive models. Part 2: Gibbs energy of solvation:	https://www.doi.org/10.1016/j.fluid.2018.04.004
Solubility Measurement and Modeling of 3-Hydroxy-2-nitropyridine in Ten Pure Solvents and Two Binary Mixed Solvents for T = (278.15–318.15) K:	https://www.doi.org/10.1021/acs.jced.9b00564
Thermodynamic Study of Solubility for Imatinib Mesylate in Nine Monosolvents and Two Binary Solvent Mixtures from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b00551
Solubility of 3,7,9,11-Tetraoxo-2,4,6,8,10-pentaaza[3.3.3]Propellane (TOPAP) in Different Pure Solvents at Temperatures between 273.15 and 318.15 K:	https://www.doi.org/10.1021/acs.jced.6b00349
(Liquid + liquid) equilibria in the (water + tetrahydrofuran + [illegible]) ternary system at temperatures between (278 and 343) K and atmospheric pressure:	https://www.doi.org/10.1016/j.fluid.2014.10.041
Solubilities of succinimide in different pure solvents and binary methanol + ethyl acetate solvent mixtures: Experimental data and correlation:	https://www.doi.org/10.1016/j.tca.2012.03.007
Excess molar enthalpies of the ternary mixtures (1-hexene + tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether) at the temperature 298.15 K:	https://www.doi.org/10.1016/j.jct.2006.03.020
Solubilities of Phenylphosphinic Acid, Methylphenylphosphinic Acid, and N-[(4-Bromo-3,5-difluorine)phenyl]maleimide in Selected Solvents:	https://www.doi.org/10.1021/je1009812
Solubilities of Hexachlorocyclotriphosphazene, and Hexaphenoxycyclotriphosphazene in Ethyl Acetate, Tetrahydrofuran, Ethanol, Trichloromethane, and Acetonitrile at Temperature between [illegible] K:	https://www.doi.org/10.1021/je100819f
Densities, Viscosities, and Sound Speed of Binary Mixtures of Hexyl Acetate with Tetrahydrofuran, 1,4-Dioxane, Anisole, and Butyl Vinyl Ether:	https://www.doi.org/10.1021/je400539h
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.03.015
Determination and correlation of solubilities of 1,2-bis(5,5-dimethyl-1,3,2-dioxyphosphacyclohexyl-2-imino) [illegible]	https://www.doi.org/10.1016/j.fluid.2013.09.058
Solid-Liquid Phase Equilibrium and Thermodynamic Properties of Olaparib in Selected Organic Solvents and (Tetrahydrofuran + MTBE, Acetonitrile + Isopropyl Alcohol) Binary Solvent Mixtures:	https://www.doi.org/10.1021/acs.jced.7b00489
Measurement of Activity Coefficients of Solutes at Infinite Dilution in (Dimethyl sulfoxide + [illegible]) Using Gas Liquid Chromatography at the Temperature 298.15 K:	https://www.doi.org/10.1021/je030122h
Surface Tensions and the Gibbs Excess Surface Concentration of Binary Mixtures of the Ionic Liquid 1-Ethyl-3-methylimidazolium Bis[(trifluoromethyl)sulfonyl]imide with Tetrahydrofuran and Acetonitrile:	https://www.doi.org/10.1021/je101146f

Quaternary isothermal vapor-liquid equilibrium of the model biofuel 2-butanone + n-heptane + tetrahydrofuran + cyclohexane using Raman spectroscopic characterization:	https://www.doi.org/10.1016/j.fluid.2018.04.009
Excess Gibbs energies and volumes of the ternary system ethanol + dimethyl carbonate + tetrahydrofuran at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2013.07.034
Phase equilibria for ternary liquid systems of (water + tetrahydrofuran + nonprotic aromatic solvent) at T = 298.2 K:	https://www.doi.org/10.1016/j.jct.2005.07.010
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 4-methyl-N-butyl-pyridinium bis(trifluoromethylsulfonyl)-imide:	https://www.doi.org/10.1016/j.jct.2009.06.011
Bubble temperature measurements on the binary mixtures formed by decane with various compounds at 95.8 kPa:	https://www.doi.org/10.1016/j.fluid.2004.10.027
Conductivity is a contrivance to explore ion-pair and triple-ion structure of ethanoates in tetrahydrofuran, dimethyl sulfoxide and their binaries:	https://www.doi.org/10.1016/j.fluid.2012.02.013
Volumetric Properties of Binary Mixtures Containing Ionic Liquids and Some Aprotic Solvents:	https://www.doi.org/10.1021/je100315g
Influence of Ionic Liquids on the Phase Behavior of Aqueous Azeotropic Systems:	https://www.doi.org/10.1021/je034183r
Thermodynamic models for determination of the solubility of 4-(4-aminophenyl)-3-morpholinone in different pure solvents and (1,4-dioxane + ethyl acetate) binary mixtures with temperatures from (278.15 to 333.15) K:	https://www.doi.org/10.1016/j.jct.2016.01.022
Assessment of Pyrrolidinium-Based Ionic Liquid for the Separation of Binary Mixtures Based on Activity Coefficients at Infinite Dilution:	https://www.doi.org/10.1021/acs.jced.9b00341
Measurement and Correlation of the Solubility of 5-Fluorouracil in Pure and Binary Solvents:	https://www.doi.org/10.1021/acs.jced.8b00425
Determination of Infinite Dilution Activity Coefficients of Several Organic Solutes in N-Butylpyridinium Nitrate/N-Octylpyridinium Nitrate by Blend Inverse Gas Chromatography:	https://www.doi.org/10.1021/acs.jced.7b00244
Thermostabilities and solubilities of flame retardant bisphenol S bis (5,5-dimethyl-1,3-dioxaphosphorinanyl-2-oxy phosphate ester) as a function of temperature:	https://www.doi.org/10.1016/j.fluid.2014.04.006
Isobaric Vapor-Liquid Equilibrium for Binary System of Tetrahydrofuran + 1-Butanol and gamma-Butyrolactone at 50.0 and 70.0 kPa:	https://www.doi.org/10.1021/acs.jced.7b00584
Measurement and correlation of solubility of trans-resveratrol in 11 solvents at T = (278.2, 288.2, 298.2, 308.2, and 318.2) K:	https://www.doi.org/10.1016/j.jct.2007.10.006
Solubility Measurement and Correlation of (+)-Biotin Intermediate Lactone in Different Organic Solvents from (287.15 to 323.75) K:	https://www.doi.org/10.1021/acs.jced.5b00857
Excess Gibbs Energies of the Ternary System 2-Methoxyethanol + Tetrahydrofuran + Cyclohexane and Other Relevant Binaries at 298.15 K:	https://www.doi.org/10.1021/je100556t
Partition Coefficients of Organic Compounds in New Imidazolium and Tetralkylammonium Based Ionic Liquids Using Inverse Gas Chromatography:	https://www.doi.org/10.1021/je9003178
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid triethylsulphonium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2008.12.005
Bubble Temperature Measurements on Some Binary Mixtures Formed by Tetrahydrofuran or Anisole with Hydrocarbons, Chlorohydrocarbons, or Butanols at (94.6 or 95.8) kPa:	https://www.doi.org/10.1021/je030141r
Measurement and Correlation of the Solubility of Dienogest in Twelve Pure and Water + Methanol Binary Solvents at Temperatures from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.9b00696
Measurement and correlation of the solubility of genistin in eleven organic solvents from T = (283.2 to 323.2) K:	https://www.doi.org/10.1016/j.jct.2015.05.019
Henry's constants and activity coefficients of some organic solutes in 1-butyl,3-methylimidazolium hydrogen sulfate and in 1-methyl,3-trimethylsilylmethylimidazolium chloride:	https://www.doi.org/10.1016/j.jct.2012.08.022
Separation of (water/butan-1-ol) binary systems based on activity coefficients at infinite dilution with phosphonium ionic liquid:	https://www.doi.org/10.1016/j.jct.2019.05.011
Solid-Liquid Phase Equilibrium of Phosphoramidic acid, N,N'-1,2-Ethanediylbis-P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je5010565
NIST Webbook:	http://webbook.nist.gov/cgi/cbook.cgi?ID=C109999&Units=SI
Thermodynamic surface properties of [BMIm][NTf2] or [EMIm][NTf2] binary mixtures with tetrahydrofuran, acetonitrile or dimethylsulfoxide:	https://www.doi.org/10.1016/j.jct.2013.02.021
Solid-liquid equilibrium of diphenyl anilinophosphonate in the different organic solvents:	https://www.doi.org/10.1016/j.fluid.2015.03.036
Solubility of daidzin in different organic solvents and (ethyl alcohol + water) mixed solvents:	https://www.doi.org/10.1016/j.jct.2015.04.025
Excess molar volume, viscosity, and refractive index study for the ternary mixture {2-methyl-2-butanol (1) + tetrahydrofuran (2) + propylamine (3)} at different temperatures: Application of the ERAS model and Peng-Robinson Stryjek-Vera equation of state:	https://www.doi.org/10.1016/j.jct.2010.05.009
Solubility Measurement and Thermodynamic Modeling of N-(4-Methylphenyl)-3-chloro-2-(phenylthio)propenamide in 12 Pure Solvents at Temperatures Ranging from 278.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.7b01011
A 1-alkylcyanopyridinium-based ionic liquid in the separation processes:	https://www.doi.org/10.1016/j.jct.2016.01.017
Thermodynamic models for determination of the solubility of 2, 5-bis(2-furylmethylidene) cyclopentan-1-one in different solvents at temperatures ranging from 308.15 to 403.15 K:	https://www.doi.org/10.1016/j.fluid.2014.01.039
Double Azeotropy in Binary Mixtures 1,1,1,2,3,4,4,5,5,5-Decafluoropentane [illegible] Tetrahydrofuran:	https://www.doi.org/10.1021/je034157t
Measurements of activity coefficients at infinite dilution of organic solutes and water in 1-propyl-1-methylpiperidinium bis{(trifluoromethyl)sulfonyl}imide ionic liquid using g.l.c.:	https://www.doi.org/10.1016/j.jct.2010.05.017
Solubilities of Bis (2,2,6,6-Tetramethyl-4-Piperidinyl) Maleate in Hexane, Heptane, Octane, m-Xylene, and Tetrahydrofuran from (253.15 to 310.15) K:	https://www.doi.org/10.1021/je034182z
Henry's Law Constants of Organic Compounds in Water and n-Octane at T = 293.15 K:	https://www.doi.org/10.1021/je900711h
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-(2-Hydroxyethyl)-3-methylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je900704b
Solubility of 3-Nitrophthalic Acid in Different Solvents between 278 K and 353 K:	https://www.doi.org/10.1021/je0604737
Solubility of Carbamazepine (Form III) in Different Solvents from (275 to 343) K:	https://www.doi.org/10.1021/je8002157
To the issue of temperature-dependent behavior of standard molar volumes of components in the binary system (water + tetrahydrofuran) at ambient pressure:	https://www.doi.org/10.1016/j.jct.2013.12.028
Thermodynamics of Fluconazole Solubility in Various Solvents at Different Temperatures:	https://www.doi.org/10.1021/je4010257
Infinite dilution activity coefficients of volatile organic compounds in two ionic liquids composed of the tris(pentafluoroethyl) trifluorophosphate ([FAP]) anion and a functionalized cation:	https://www.doi.org/10.1016/j.jct.2013.05.035

Determination and correlation thermodynamic models for solid liquid equilibrium of the Nifedipine in pure and mixture organic solvents.	https://www.doi.org/10.1016/j.jct.2016.07.035
Densities, speeds of sound and viscosities of binary mixtures of tetrahydrofuran with 1-hexanol, 1-octanol and 1-decanol at T = (298.15 to 313.15) K.	https://www.doi.org/10.1016/j.jct.2013.10.025
Determination and Thermodynamic Analysis of the Solubility of Limonin in Four Organic Solvents and Ethyl Acetate + 2-Propanol Binary Solvents from 283.2 to 323.2 K.	https://www.doi.org/10.1021/acs.jced.7b00851
Tailor-made ionic liquids.	https://www.doi.org/10.1016/j.jct.2005.04.013
Volumetric and compressibility behaviour of ionic liquid, 1-n-butyl-3-methylimidazolium hexafluorophosphate and tetrabutylammonium hexafluorophosphate in organic solvents at T = 298.15 K.	https://www.doi.org/10.1016/j.jct.2005.07.018
Solubility of Methane in Tetrahydrofuran + Ethanol at High Pressure	https://www.doi.org/10.1021/je800354p
Activity Coefficients at Infinite Dilution of Polar Solutes in 1-Propyl-2,3-dimethylimidazolium Tetrafluoroborate Using Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je700640r
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpyrrolidinium [illegible]	https://www.doi.org/10.1016/j.jct.2013.01.007
Thermodynamic models for determination of the solid-liquid equilibrium of istradefylline in ethyl acetate plus ([illegible] tetrahydrofuran, acetone) binary solvent mixtures	https://www.doi.org/10.1016/j.jct.2017.03.015
Application of trihexyltetradecylphosphonium [illegible] ionic liquid for various separation problems: Activity coefficients at infinite dilution measurements [illegible]	https://www.doi.org/10.1016/j.fluid.2019.03.023
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(3-hydroxypropyl)pyridinium bis(trifluoromethylsulfonyl)-amide.	https://www.doi.org/10.1016/j.jct.2011.04.018
Measurement and Correlation of Solubility of Pimelic Acid in Ether, Tetrahydrofuran, Ethanol, and Methanol.	https://www.doi.org/10.1021/je900629v
[illegible] Oxygen Containing Compounds:	https://www.doi.org/10.1007/s10765-010-0860-3
The use of ionic liquids for separation of binary hydrocarbons mixtures based on gamma infinity data measurements:	https://www.doi.org/10.1016/j.jct.2018.07.024
Determination of the solubility, dissolution enthalpy and entropy of suberic acid in different solvents.	https://www.doi.org/10.1016/j.fluid.2012.06.012
Excess molar enthalpies of ternary mixtures (1-hexene + tetrahydrofuran + 2,2-dimethylbutane or 2,3-dimethylbutane) at the temperature 298.15 K.	https://www.doi.org/10.1016/j.jct.2010.10.002
Vapor pressures and activity coefficients of binary mixtures of 1-ethyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide with acetonitrile and tetrahydrofuran.	https://www.doi.org/10.1016/j.jct.2011.09.022
Measurement of activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-hexyl-1,4-diaza[2.2.2]bicyclooctanium bis(trifluoromethylsulfonyl)imide using gas-liquid chromatography.	https://www.doi.org/10.1016/j.jct.2013.10.026
Activity coefficients at infinite dilution, physicochemical and thermodynamic properties for organic solutes and water in the ionic liquid ethyl-dimethyl-(2-methoxyethyl)ammonium [illegible]	https://www.doi.org/10.1016/j.jct.2015.05.022
Studies of viscosities of dilute solutions of alkylamines in [illegible] 1,4-dioxane and oxolane at 303.15 K.	https://www.doi.org/10.1016/j.tca.2009.07.008
Study of the Surface Tension of Chlorocyclohexane or Bromocyclohexane with Some Cyclic Ethers.	https://www.doi.org/10.1021/je0500577
Activity coefficients at infinite dilution of organic solutes in the ionic liquid 1-ethyl-3-methylimidazolium methanesulfonate:	https://www.doi.org/10.1016/j.fluid.2010.10.008
Solubility of Sulpiride in Pure Organic Solvents between (278 and 333) K:	https://www.doi.org/10.1021/je1002016
Thermodynamics and selectivity of separation based on activity coefficients at infinite dilution of various solutes in 1-allyl-3-methylimidazolium bis{(trifluoromethyl)sulfonyl}imide ionic liquid.	https://www.doi.org/10.1016/j.jct.2016.06.028
Solubilities of Betulinic Acid in Thirteen Organic Solvents at Different Temperatures.	https://www.doi.org/10.1021/je200531k
Separation of water/butan-1-ol based on activity coefficients at infinite dilution in 1,3-didecyl-2-methylimidazolium dicyanamide ionic liquid.	https://www.doi.org/10.1016/j.jct.2017.10.003
Interactions of Volatile Organic Compounds with the Ionic Liquid 1-Butyl-1-methylpyrrolidinium Dicyanamide.	https://www.doi.org/10.1021/je200822w
Temperature and concentration dependence of volumetric properties of (tetrahydrofuran + methyl acrylate, or + ethyl acrylate, or + n-butyl acrylate, or + tert-butyl acrylate) binary mixtures.	https://www.doi.org/10.1016/j.jct.2016.10.042
Solubilities of Proxetine Hydrochloride Hemihydrate between 286 K and 363 K.	https://www.doi.org/10.1021/je0498762
Solubilities of Terephthalic Acid, Phthalic Acid, and Isophthalic Acid in Tetrahydrofuran, Cyclohexanone, 1,2-Diethoxyethane, and Acetophenone.	https://www.doi.org/10.1021/je9001976
Solubility Measurements at 296 and 310 K and Physicochemical Characterization of Abiraterone and Abiraterone Acetate.	https://www.doi.org/10.1021/acs.jced.8b00566
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-hexyl-3-methylimidazolium tetracyanoborate.	https://www.doi.org/10.1016/j.jct.2011.11.025
Solubilities of the Gaseous and Liquid Solutes and Their Thermodynamics of Solubilization in the Novel Room-Temperature Ionic Liquids at Infinite Dilution by Gas Chromatography.	https://www.doi.org/10.1021/je700297c
Abraham model linear free energy relationships for describing the partitioning and solubility behavior of nonelectrolyte organic solutes dissolved in pyridine at 298.15 K.	https://www.doi.org/10.1016/j.fluid.2016.10.009
Partitioning of diuron in a novel aqueous two-phase system based on [illegible]	https://www.doi.org/10.1016/j.fluid.2016.08.012
Solubility Determination and Modeling of Sodium Mononitrobenzenesulphonate in Different Solvents at Temperatures Ranging From (283.15 to 323.15) K.	https://www.doi.org/10.1021/acs.jced.5b00783
Excess Enthalpies of Dibromomethane with Acetone, 1,4-Dioxane, Pyridine, Diethyl Ether, Ethyl Methyl Ketone, and Tetrahydrofuran at 303.15 K.	https://www.doi.org/10.1021/je900547w
Activity Coefficients at Infinite Dilution for Hydrocarbons in Fatty Alcohols Determined by Gas-Liquid Chromatography.	https://www.doi.org/10.1021/je1005517
Measurement and correlation of solubility of succinic anhydride in pure solvents and binary solvent mixtures.	https://www.doi.org/10.1016/j.jct.2016.09.025
Gas liquid chromatography measurements of activity coefficients at infinite dilution of various organic solutes and water in tri-iso-butylmethylphosphonium tosylate ionic liquid.	https://www.doi.org/10.1016/j.jct.2010.01.004
Determination of Activity Coefficients at Infinite Dilution of Organic Solutes in [illegible]-3-methylimidazolium Nonafluoro-1-butanesulfonate Using Gas-Liquid Chromatography.	https://www.doi.org/10.1021/acs.jced.8b00080
Phase Behavior of Binary Mixtures Containing Succinic Acid or Its Esters.	https://www.doi.org/10.1021/acs.jced.7b00005
Solubility and Thermodynamic Functions of Isatin in Pure Solvents.	https://www.doi.org/10.1021/je500396b

Excess Gibbs energies and volumes of the ternary system chloroform + tetrahydrofuran + cyclohexane at 298.15 K:	https://www.doi.org/10.1016/j.fluid.2010.06.007
Physicochemical properties and activity coefficients at infinite dilution for organic solutes and water in a novel bicyclic guanidinium superbase-derived protic ionic liquid:	https://www.doi.org/10.1016/j.jct.2012.09.033
Interactions of volatile organic compounds with the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.09.028
Study of isobaric vapour-liquid equilibrium of some cyclic ethers with 1-chloropropane: Experimental results and SAFT-VR modelling:	https://www.doi.org/10.1016/j.fluid.2009.01.010
Experimental and theoretically study of interaction between organic compounds and tricyanomethanide based ionic liquids:	https://www.doi.org/10.1016/j.jct.2014.12.027
Density, Viscosity, Refractive Index, and Speed of Sound in the Binary Mixtures of Tri-n-butylamine + Triethylamine, + Tetrahydrofuran, + Tetradecane, + Tetrachloroethylene, + Pyridine, or + Trichloroethylene at (298.15, 303.15, and 308.15) K:	https://www.doi.org/10.1021/je030147g
Phase Equilibria Study of the Binary Systems (N-Butyl-4-methylpyridinium Tosylate Ionic Liquid + Organic Solvent, or Water):	https://www.doi.org/10.1021/je8006309
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Butyl-3-methylpyridinium Trifluoromethanesulfonate:	https://www.doi.org/10.1021/je1000582
Antisolvent Precipitation of Vitamin B6: A Thermodynamic Study:	https://www.doi.org/10.1021/je200853g
Measurement and Correlation of the Solubility of 2-Cyanoacetamide in 14 Pure Solvents and the Mixing Properties of Solutions:	https://www.doi.org/10.1021/acs.jced.8b01205
Solubility Measurement and the Correlation of 1-Naphthaleneacetic Acid in Pure and Methanol + Water Binary Solvents from T = (278.25 to 323.55) K:	https://www.doi.org/10.1021/acs.jced.6b00816
Measurements of activity coefficients at infinite dilution for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tricyanomethanide:	https://www.doi.org/10.1016/j.jct.2013.07.004
Solubility of 1,3,2-dioxaphosphorinane-5,5-dimethyl -2-(2,4,6-tribromophenoxy) 2-oxide in selected solvents:	https://www.doi.org/10.1016/j.fluid.2013.09.023
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 1-(2-methoxyethyl)-1-methylpiperidinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2012.01.019
Thermodynamic and acoustic properties of binary mixtures of oxolane with aniline and substituted anilines at 303.15, 313.15 and 323.15 K:	https://www.doi.org/10.1016/j.tca.2010.04.025
Measurement and correlation of the solubility of rivaroxaban (form I) in binary mixtures of ethyl acetate with tetrahydrofuran, N,N-dimethylformamide, and N,N-dimethylacetamide from T = (278.15 to 318.15) K:	https://www.doi.org/10.1016/j.jct.2015.10.016
Experiments on the continuous separation of gas mixtures via dissolution and hydrate formation in the presence of THF:	https://www.doi.org/10.1016/j.fluid.2013.11.002
Measurement and correlation of vapor liquid equilibria of binary systems containing the ionic liquids [EMIM][(CF3SO2)2N], [BMIM][(CF3SO2)2N], [MMIM][(CH3)2PO4] and oxygenated organic compounds respectively water:	https://www.doi.org/10.1016/j.fluid.2005.01.002
Ultrasonic velocity, viscosity and excess properties of binary mixture of tetrahydrofuran with 1-propanol and 2-propanol:	https://www.doi.org/10.1016/j.fluid.2005.07.013
Thermodynamic properties of betulinic acid in THF + water mixed solvents at different temperatures:	https://www.doi.org/10.1016/j.tca.2014.10.026
Thermodynamics and activity coefficients at infinite dilution for organic solutes in the ionic liquid 1-hexyl-2,3-dimethylimidazolium bis(trifluoromethylsulfonyl)imide:	https://www.doi.org/10.1016/j.jct.2016.07.021
Thermodynamic Properties of Mixtures Containing Precursors of Vitamin B5:	https://www.doi.org/10.1021/je020060f
Solubilities of (a,Z)-2-Amino-a-(methoxyimino)-4-thiazoleethanethioic Acid S-2-Benzothiazolyl Ester in Different Pure Solvents and Binary Mixtures of Tetrahydrofuran + Dichloromethane and + 1,2-Dichloroethane:	https://www.doi.org/10.1021/je100890f
Crystal Structures, Reaction Rates, and Selected Physical Properties of [illegible] (Halo = Fluoride, Chloride, and Bromide):	https://www.doi.org/10.1021/je3005112
The potential of head-space gas chromatography for VLE measurements:	https://www.doi.org/10.1016/j.jct.2012.01.020
Determination of Activity Coefficients at Infinite Dilution of 35 Solutes in the Ionic Liquid, 1-Butyl-3-methylimidazolium Tosylate, Using Gas-Liquid Chromatography:	https://www.doi.org/10.1021/je100410k
Estimated Solubility Method:	http://pubs.acs.org/doi/suppl/10.1021/ci034243x/suppl_file/ci034243xsi20040112_053635.txt
Experimental and predicted viscosities of binary mixtures of cyclic ethers with 1-chloropentane or 1-chlorohexane at 283.15, 298.15, and 313.15K:	https://www.doi.org/10.1016/j.tca.2008.11.009
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium trifluoromethanesulfonate using GLC:	https://www.doi.org/10.1016/j.fluid.2009.01.011
Measurements of activity coefficients at infinite dilution of aliphatic and aromatic hydrocarbons, alcohols, thiophene, tetrahydrofuran, MTBE, and water in ionic liquid [BMIM][SCN] using GLC:	https://www.doi.org/10.1016/j.jct.2008.12.018
Unravelling various types of non-covalent interactions of benzyl amine with ether in n-hexane at 303.15 K by ultrasonic and DFT methods:	https://www.doi.org/10.1016/j.fluid.2018.08.001
High Pressure Vapor-Liquid Equilibria of Some Carbon Dioxide + Organic Binary Systems:	https://www.doi.org/10.1021/je0498560
Excess Molar Enthalpies of Dipropyl Ether + Dibutyl Ether + (1-Hexene or Tetrahydrofuran) at 298.15 K:	https://www.doi.org/10.1021/je100897y
Thermodynamic Properties of Water + Tetrahydrofuran and Water + 1,4-Dioxane Mixtures at (303.15, 313.15, and 323.15) K:	https://www.doi.org/10.1021/je030196t
Vapor-Liquid Equilibrium for Tetrahydrofuran + Methanol + Tetrafluoroborate-based ionic liquids at 101.3 kPa:	https://www.doi.org/10.1016/j.fluid.2013.09.060
A Digital Variable-Angle Rolling Ball Viscometer for Measurement of Viscosity, Density, and Bubble-Point Pressure of CO2 and Organic Liquid Mixtures:	https://www.doi.org/10.1007/s10765-008-0542-6
Effect of solvent on the volumetric behavior of N,N'-salicylidenephenyl diamine (Salophen) Schiff base at different temperatures (288.15 – 318.15) K:	https://www.doi.org/10.1016/j.fluid.2013.05.001
Densities, ultrasonic speeds and refractive indices of phenetole with N-methyl-2-pyrrolidone, N,N-dimethylformamide and tetrahydrofuran binary mixtures at different temperatures:	https://www.doi.org/10.1016/j.jct.2016.07.032
Solubility of Irbesartan (Form A) in Different Solvents between 278 K and 323 K:	https://www.doi.org/10.1021/je700296x
Separation of binary mixtures based on gamma infinity data using [OMMIM][NTf2] ionic liquid and modelling of thermodynamic functions:	https://www.doi.org/10.1016/j.jct.2017.12.012
Measurement and Correlation of the Solubility of Mifepristone in Eight Pure and Water + Methanol Mixed Solvents at Temperatures from 273.15 to 318.15 K:	https://www.doi.org/10.1021/acs.jced.8b01062
Vapor-Liquid Equilibria of the Binary Carbon Dioxide-Tetrahydrofuran Mixture System:	https://www.doi.org/10.1021/je0202228
Solubility Study of (2E)-1-(3-Pyridyl)-3-(dimethylamino)-2-propen-1-one in Different Pure Solvents and Binary Solvent Mixtures from 278.15 to 328.15 K:	https://www.doi.org/10.1021/acs.jced.9b00047

Excess Enthalpies of the Ternary Mixtures: Tetrahydrofuran + (Hexane or Cyclohexane) + Decane at 298.15K:	https://www.doi.org/10.1021/je020159m
Activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-ethyl-3-methylimidazolium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.02.012
Solubilities of Phosphorus-Containing Compounds in Selected Solvents:	https://www.doi.org/10.1021/je100341q
4-Amino-3,6-dichloropyridazine Solubility Measurement and Correlation in Seven Pure Organic Solvents from (278.15 to 333.15) K:	https://www.doi.org/10.1021/je500286x
Excess molar enthalpies of the ternary mixtures: {tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether + n-dodecane} at the temperature 298.15 K:	https://www.doi.org/10.1016/j.jct.2004.12.005
Activity Coefficients at Infinite Dilution of Organic Solutes and Water in Tributylethylphosphonium Diethylphosphate Using Gas Liquid Chromatography: Thermodynamic Properties of Mixtures Containing Ionic Liquids	https://www.doi.org/10.1021/acs.jced.5b00980
KDB Vapor Pressure Data:	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1035
Dissolution behavior and thermodynamic properties of lapatinib ditosylate in pure and mixed organic solvents from T = (283.15-323.15) K:	https://www.doi.org/10.1016/j.fluid.2019.01.005
Thermodynamic study of molecular interaction selectivity in separation processes based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2018.02.014
Solubility of Decabromodiphenyl Ether in Different Solvents at (283.0 to 323.0) K:	https://www.doi.org/10.1021/je100316k
(Liquid + liquid), (solid + liquid), and (solid + liquid + liquid) equilibria of systems containing cyclic ether (tetrahydrofuran or 1,3-dioxolane), water, and a biological buffer MOPS:	https://www.doi.org/10.1016/j.jct.2014.11.002
Solubility of Phenacetinum in Methanol, Ethanol, 1-Propanol, 1-Butanol, 2-Propanol, Tetrahydrofuran, Ethyl Acetate, and Benzene between 282.65 K and 333.70 K:	https://www.doi.org/10.1021/je700209v
Low cost apparatus for rapid boiling point determination of small air sensitive samples under inert atmosphere:	https://www.doi.org/10.1016/j.tca.2017.05.005
Densities, Viscosities, Refractive Indexes, and Surface Tensions for Binary and Ternary Mixtures of Tetrahydrofuran, 2-Propanol, and 2,2,4-Trimethylpentane:	https://www.doi.org/10.1021/je700626v
Volumetric properties of binary mixtures of N-ethylformamide with tetrahydrofuran, 2-butanone and ethylacetate from (293.15 to 313.15) K:	https://www.doi.org/10.1016/j.jct.2012.02.033
KDB	https://www.cheric.org/files/research/kdb/mol/mol1035.mol
Experimental Isobaric Vapor-Liquid Equilibrium Data for Binary Mixtures of Cyclic Ethers with (1-Methylethyl)benzene:	https://www.doi.org/10.1021/je800087p
Investigation on thermophysical and excess properties of binary mixtures of imidazolium based ionic liquids at temperatures (293.15 to 323.15) K: III $[C_nmim][PF_6]$ (n = 4, 6, 8) + THF:	https://www.doi.org/10.1016/j.jct.2015.02.011
Determination and Correlation of Solubilities of Lithium Bis(oxalato)borate in Six Different Solvents from (293.15 to 363.15) K:	https://www.doi.org/10.1021/je500068b
Ion-Pair and Triple-Ion Formation by Some Tetraalkylammonium Iodides in Binary Mixtures of 1,4-Dioxane + Tetrahydrofuran:	https://www.doi.org/10.1021/je7004787
Salt effect on liquid-liquid equilibria of tetrahydrofuran/water/5-hydroxymethylfurfural systems:	https://www.doi.org/10.1016/j.fluid.2019.04.018
Thermodynamics and activity coefficients at infinite dilution measurements for organic solutes and water in the ionic liquid 1-butyl-1-methylpyrrolidinium tetracyanoborate:	https://www.doi.org/10.1016/j.jct.2011.06.007
Activity coefficients at infinite dilution and physicochemical properties for organic solutes and water in the ionic liquid 4-(3-hydroxypropyl)-4-methylmorpholinium bis(trifluoromethylsulfonyl)-amide:	https://www.doi.org/10.1016/j.jct.2015.02.024
Activity Coefficients at Infinite Dilution Measurements for Organic Solutes and Water in the Ionic Liquid 1-Hexyl-3-methylimidazolium Thiocyanate:	https://www.doi.org/10.1021/je900890u
Solubilities of Phosphonic Acid, P,P'-(1,4-Piperazinediyl)bis-, P,P,P',P'-tetraphenyl Ester in Selected Solvents:	https://www.doi.org/10.1021/je400412z
Experimental Determination and Thermodynamic Models for Solid-Liquid Equilibrium of 4-(4-Aminophenoxy)-N-methylpyridine-2-carboxamide in Pure and Binary Solvent Mixtures for (278.15-328.15) K:	https://www.doi.org/10.1021/acs.jced.8b00165
Solubilities of some organic solutes in 1-ethyl-3-methylimidazolium acetate. Chromatographic measurements and predictions from COSMO-RS:	https://www.doi.org/10.1016/j.jct.2011.11.005
Solubilities of Tris(o-phenylenedioxy)cyclotriphosphazene in Selected Solvents:	https://www.doi.org/10.1021/je200252c
Thermodynamics and limiting activity coefficients measurements for organic solutes and water in the ionic liquid 1-dodecyl-3-methylimidazolium bis(trifluoromethylsulfonyl) imide:	https://www.doi.org/10.1016/j.jct.2016.08.008
Thermodynamic Models for Determination of the Solubility of Boc-(R)-3-Amino-4-(2,4,5-trifluorophenyl)butanoic acid in Different Pure Solvents and (Tetrahydrofuran + n-Butanol) Binary Mixtures with Temperatures from 280.15 to 330.15 K:	https://www.doi.org/10.1021/acs.jced.5b00758
Ammonium ionic liquids in extraction of bio-butan-1-ol from water phase using activity coefficients at infinite dilution:	https://www.doi.org/10.1016/j.fluid.2018.09.024
Vapor-Liquid Equilibrium for Ternary and Binary Mixtures of Tetrahydrofuran, Cyclohexane, and 1-Propanol at 101.3 kPa:	https://www.doi.org/10.1021/je4004993
Measurement, correlation and dissolution thermodynamics of biological active chalcone in organic solvents at different temperatures:	https://www.doi.org/10.1016/j.jct.2014.03.026
Excess molar enthalpies of the ternary mixtures: (tetrahydrofuran or 2-methyltetrahydrofuran + methyl tert-butyl ether + n-octane) at the temperature 298.15 K:	https://www.doi.org/10.1016/j.jct.2005.07.017
Solubility Measurement and Phase Equilibrium Modeling of 2-Aminobenzamide in 15 Pure Solvents:	https://www.doi.org/10.1021/acs.jced.9b00350
Solubility of citric, malonic, and malic acids in different solvents from 303.2 to 333.2 K:	https://www.doi.org/10.1016/j.fluid.2011.09.033
Crippen Method:	http://pubs.acs.org/doi/abs/10.1021/ci990307l
A volumetric and viscosity study for the binary mixtures of 1-hexyl-3-methylimidazolium tetrafluoroborate with some molecular solvents:	https://www.doi.org/10.1016/j.jct.2010.12.027
KDB Pure (Korean Thermophysical Properties Databank):	https://www.cheric.org/research/kdb/hcprop/showprop.php?cmpid=1035
Measurement of heat capacities of ionic liquids by differential scanning calorimetry:	https://www.doi.org/10.1016/j.fluid.2006.03.015
Biomass-Derived Platform Chemicals: Thermodynamic Studies on the Extraction of 5-Hydroxymethylfurfural from Ionic Liquids:	https://www.doi.org/10.1021/je300529j
Separation of binary mixtures hexane/hex-1-ene, cyclohexane/cyclohexene and ethylbenzene/styrene based on limiting activity coefficients:	https://www.doi.org/10.1016/j.jct.2017.03.004
Solid Liquid Phase Equilibria of N,N'-[1,3-Phenylenebis(methylene)]bis(phosphoramidic acid) P,P,P',P'-Tetraphenyl Ester in 10 Pure Solvents:	https://www.doi.org/10.1021/je4010917

Solubility and Thermodynamic Behavior of Veramoss in Different Pure Solvents and (Ethanol + Water) Binary Solvents	https://www.doi.org/10.1021/acs.jced.6b00646
On the solubility of nicotinic acid and isonicotinic acid in water and organic solvents	https://www.doi.org/10.1016/j.jct.2013.01.024
Measurement and correlation solubility and mixing properties of dimethyl succinyl succinate in pure and mixture organic solvents from (278.15 to 323.15) K	https://www.doi.org/10.1016/j.fluid.2013.06.037
Liquid-Liquid Equilibria for the Ternary Systems of (Water + Tetrahydrofuran + 1-Alkyl-3-methylimidazolium bis(trifluoromethylsulfonyl)imide at 101.3 kPa	https://www.doi.org/10.1021/je049747l
Solubility of benzoic acid in (ethanol + water) mixtures from (278.15 to 333.15) K	https://www.doi.org/10.1016/j.tca.2004.08.001
Measurement and correlation of compressibilities, refractive indexes, and viscosities of tetrahydrofuran with alkanols at 298.15 K	https://www.doi.org/10.1016/j.jct.2009.11.005
Partial molar volumes of organic solutes in water. XVIII. Cyclic ethers at 308.15 K	https://www.doi.org/10.1016/j.jct.2018.05.017
Measurement and correlation of isobaric vapor-liquid equilibria of methanol + tetrahydrofuran + [HMIMI][BF4]	https://www.doi.org/10.1016/j.fluid.2016.04.012

Legend

affp:	Proton affinity
basg:	Gas basicity
chg:	Standard gas enthalpy of combustion
chl:	Standard liquid enthalpy of combustion
cpg:	Ideal gas heat capacity
cpl:	Liquid phase heat capacity
dvisc:	Dynamic viscosity
gf:	Standard Gibbs free energy of formation
hf:	Enthalpy of formation at standard conditions
hfus:	Enthalpy of fusion at standard conditions
hfust:	Enthalpy of fusion at a given temperature
hvap:	Enthalpy of vaporization at standard conditions
hvapt:	Enthalpy of vaporization at a given temperature
ie:	Ionization energy
kvisc:	Kinematic viscosity
log10ws:	Log10 of Water solubility in mol/l
logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pvap:	Vapor pressure
rfi:	Refractive Index
rhol:	Liquid Density
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
sfust:	Entropy of fusion at a given temperature
sg:	Molar entropy at standard conditions
sl:	Liquid phase molar entropy at standard conditions
speedsl:	Speed of sound in fluid
srf:	Surface Tension
tb:	Normal Boiling Point Temperature

tbp:	Boiling point at given pressure
tc:	Critical Temperature
tf:	Normal melting (fusion) point
tt:	Triple Point Temperature
vc:	Critical Volume
volm:	Molar Volume