

# 2-Propanol, 1-amino-, (S)-

<b>Other names:</b>	(S)-(+)-1-Amino-2-propanol (S)-(+)-1-aminopropan-2-ol
<b>Inchi:</b>	InChI=1S/C3H9NO/c1-3(5)2-4/h3,5H,2,4H2,1H3/t3-/m1/s1
<b>InchiKey:</b>	HXKKHQJGJAFBHI-GSVOUGTGSA-N
<b>Formula:</b>	C3H9NO
<b>SMILES:</b>	CC(O)CN
<b>Mol. weight [g/mol]:</b>	75.11
<b>CAS:</b>	2799-17-9

## Physical Properties

Property code	Value	Unit	Source
gf	-98.43	kJ/mol	Joback Method
hf	-228.97	kJ/mol	Joback Method
hfus	9.29	kJ/mol	Joback Method
hvap	49.20	kJ/mol	Joback Method
log10ws	0.11		Crippen Method
logp	-0.674		Crippen Method
mvol	68.980	ml/mol	McGowan Method
pc	5495.11	kPa	Joback Method
tb	433.20	K	NIST Webbook
tc	613.40	K	Joback Method
tf	252.65	K	Joback Method
vc	0.245	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	142.31	J/molxK	432.31	Joback Method
cpg	149.06	J/molxK	462.49	Joback Method
cpg	155.54	J/molxK	492.67	Joback Method
cpg	161.74	J/molxK	522.85	Joback Method
cpg	167.68	J/molxK	553.03	Joback Method
cpg	173.36	J/molxK	583.21	Joback Method
cpg	178.78	J/molxK	613.40	Joback Method

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.61162e+01
Coeff. B	-4.27299e+03
Coeff. C	-6.15660e+01
Temperature range (K), min.	331.52
Temperature range (K), max.	457.04

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C2799179&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C2799179&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>g<sub>f</sub>:</b>	Standard Gibbs free energy of formation
<b>h<sub>f</sub>:</b>	Enthalpy of formation at standard conditions
<b>h<sub>fus</sub>:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mc<sub>vol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

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