

1-Octyne

Other names:	HEXYLACETYLENE
	Octyne-1
	oct-1-yne
Inchi:	InChI=1S/C8H14/c1-3-5-7-8-6-4-2/h1H,4-8H2,2H3
InchiKey:	UMIPWJGWASORKV-UHFFFAOYSA-N
Formula:	C8H14
SMILES:	C#CCCCCCC
Mol. weight [g/mol]:	110.20
CAS:	629-05-0

Physical Properties

Property code	Value	Unit	Source
gf	239.55	kJ/mol	Joback Method
hf	80.70 ± 3.60	kJ/mol	NIST Webbook
hfus	19.45	kJ/mol	Joback Method
hvap	42.34	kJ/mol	NIST Webbook
hvap	42.30 ± 0.10	kJ/mol	NIST Webbook
ie	9.95 ± 0.02	eV	NIST Webbook
log10ws	-3.66		Estimated Solubility Method
log10ws	-3.66		Aqueous Solubility Prediction Method
logp	2.590		Crippen Method
mcvol	114.980	ml/mol	McGowan Method
pc	2960.12	kPa	Joback Method
rinpol	787.50		NIST Webbook
rinpol	784.00		NIST Webbook
rinpol	784.00		NIST Webbook
rinpol	784.00		NIST Webbook
rinpol	783.60		NIST Webbook
rinpol	786.00		NIST Webbook
rinpol	812.00		NIST Webbook
rinpol	797.00		NIST Webbook
rinpol	811.00		NIST Webbook
rinpol	784.00		NIST Webbook
rinpol	784.00		NIST Webbook
rinpol	784.50		NIST Webbook
rinpol	811.00		NIST Webbook

rinpol	812.00		NIST Webbook
rinpol	783.00		NIST Webbook
rinpol	789.00		NIST Webbook
rinpol	788.10		NIST Webbook
rinpol	783.52		NIST Webbook
rinpol	785.80		NIST Webbook
rinpol	783.52		NIST Webbook
ripol	1047.10		NIST Webbook
ripol	1034.00		NIST Webbook
ripol	1036.00		NIST Webbook
ripol	1031.00		NIST Webbook
tb	372.56	K	Joback Method
tc	547.48	K	Joback Method
tf	200.38	K	Aqueous Solubility Prediction Method
tf	193.67 ± 0.10	K	NIST Webbook
tf	194.15 ± 1.50	K	NIST Webbook
tf	193.65 ± 0.30	K	NIST Webbook
vc	0.446	m3/kmol	Joback Method

Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	207.74	J/molxK	372.56	Joback Method
cpg	219.42	J/molxK	401.71	Joback Method
cpg	230.61	J/molxK	430.87	Joback Method
cpg	241.34	J/molxK	460.02	Joback Method
cpg	251.62	J/molxK	489.17	Joback Method
cpg	261.46	J/molxK	518.33	Joback Method
cpg	270.87	J/molxK	547.48	Joback Method
hvapt	35.83	kJ/mol	399.40	NIST Webbook
hvapt	38.50	kJ/mol	378.50	NIST Webbook

Correlations

Information	Value
Property code	pvap
Equation	ln(Pvp) = A + B/(T + C)

Coeff. A	1.39482e+01
Coeff. B	-3.12653e+03
Coeff. C	-6.42400e+01
Temperature range (K), min.	293.11
Temperature range (K), max.	426.24

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/T + C \cdot \ln(T) + D \cdot T^2$
Coeff. A	-2.39993e+00
Coeff. B	-4.34412e+03
Coeff. C	3.22566e+00
Coeff. D	-8.94889e-06
Temperature range (K), min.	357.15
Temperature range (K), max.	400.15

Sources

Screening of environmental friendly ionic liquid as a solvent for the CO₂ Vapor Pressure Data

<https://www.doi.org/10.1016/j.jct.2015.08.017>

Different types of equations problem:

<https://www.thermo.com/research/kdb/hcprop/showprop.php?cmpid=428>

Insight from activity coefficients at infinite dilution measurement using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2013.09.007>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2010.10.026>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2018.02.014>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1021/je0602925>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2012.05.017>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2009.07.010>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2010.06.009>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2018.07.024>

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<https://www.doi.org/10.1016/j.jct.2016.07.017>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2012.03.015>

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<https://www.doi.org/10.1016/j.jct.2010.05.017>

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<https://www.doi.org/10.1016/j.jct.2008.04.002>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.fluid.2005.04.021>

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<https://www.doi.org/10.1016/j.jct.2015.05.014>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.jct.2011.11.021>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<https://www.doi.org/10.1016/j.fluid.2017.12.029>

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

https://en.wikipedia.org/wiki/Joback_method

Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

<http://onschallenge.wikispaces.com/file/view/AqueousDataset002.xlsx/351826032/AqueousDataset002.xlsx>

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<https://www.doi.org/10.1016/j.fluid.2010.08.016>

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Measurement of activity coefficients at infinite dilution using microfluidic device for organic solutes and water in 1-butyl-3-methylimidazolium hexafluorophosphate

[illegible]

<https://www.doi.org/10.1016/j.ijct.2008.01.004>

<https://www.doi.org/10.1016/j.fluid.2009.01.011>

<http://pubs.acs.org/doi/abs/10.1021/ci990307l>

<https://www.doi.org/10.1021/acs.iced.8b00080>

<https://www.doi.org/10.1016/j.fluid.2017.06.001>

<https://www.doi.org/10.1016/j.fluid.2018.01.019>

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<https://www.doi.org/10.1016/j.ijct.2017.03.004>

<https://www.chemic.org/files/research/kdb/mol/mol428.mol>

<https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure>

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<https://www.doi.org/10.1021/ie800105r>

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<https://www.doi.org/10.1016/j.ijct.2013.08.030>

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<https://www.doi.org/10.1021/je101008y>

<https://www.doi.org/10.1016/j.fluid.2008.10.008>

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<https://www.doi.org/10.1016/j.fluid.2018.06.003>

<https://www.doi.org/10.1016/j.fluid.2006.07.015>

<https://www.doi.org/10.1016/j.ijct.2010.04.011>

<https://www.doi.org/10.1016/j.ijct.2014.04.034>

<https://www.doi.org/10.1016/j.fluid.2018.07.038>

<https://www.doi.org/10.1016/j.ijet.2013.03.006>

<https://www.doi.org/10.1016/j.ijet.2018.05.003>

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<https://www.doi.org/10.1021/jc0256481>

<https://www.doi.org/10.1016/j.jst.2008.12.005>

<https://www.doi.org/10.1016/j.jst.2025.01.015>

<https://arxiv.org/abs/1910.10044> [arXiv:1910.10044] [arXiv:1910.10044]

<https://www.doi.org/10.1016/j.fluid.2023.09.024>

<https://www.doi.org/10.1016/j.fluid.2019.02.004>

<https://doi.org/10.1016/j.jag.2017.10.002>

doi:10.1016/j.jurimetrics.2014.11.005

logp:	Octanol/Water partition coefficient
mcvol:	McGowan's characteristic volume
pc:	Critical Pressure
pvap:	Vapor pressure
rinpol:	Non-polar retention indices
ripol:	Polar retention indices
tb:	Normal Boiling Point Temperature
tc:	Critical Temperature
tf:	Normal melting (fusion) point
vc:	Critical Volume

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