

# cis-1-Bromo-1-propene

<b>Inchi:</b>	InChI=1S/C3H5Br/c1-2-3-4/h2-3H,1H3/b3-2-
<b>InchiKey:</b>	NNQDMQVWOWCVEM-IHWYPQMZSA-N
<b>Formula:</b>	C3H5Br
<b>SMILES:</b>	CC=CBr
<b>Mol. weight [g/mol]:</b>	120.98
<b>CAS:</b>	590-13-6

## Physical Properties

Property code	Value	Unit	Source
gf	68.92	kJ/mol	Joback Method
hf	38.30	kJ/mol	Joback Method
hfus	9.01	kJ/mol	Joback Method
hvap	28.67	kJ/mol	Joback Method
log10ws	-1.86		Crippen Method
logp	1.915		Crippen Method
mvol	66.330	ml/mol	McGowan Method
pc	5213.15	kPa	Joback Method
rinpol	582.00		NIST Webbook
tb	338.36	K	Joback Method
tc	534.95	K	Joback Method
tf	178.29	K	Joback Method
vc	0.245	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	81.97	J/molxK	338.36	Joback Method
cpg	87.80	J/molxK	371.12	Joback Method
cpg	93.25	J/molxK	403.89	Joback Method
cpg	98.35	J/molxK	436.65	Joback Method
cpg	103.12	J/molxK	469.42	Joback Method
cpg	107.57	J/molxK	502.18	Joback Method
cpg	111.73	J/molxK	534.95	Joback Method
dvisc	0.0028576	Paxs	178.29	Joback Method

dvisc	0.0015320	Paxs	204.97	Joback Method
dvisc	0.0009481	Paxs	231.65	Joback Method
dvisc	0.0006479	Paxs	258.32	Joback Method
dvisc	0.0004755	Paxs	285.00	Joback Method
dvisc	0.0003679	Paxs	311.68	Joback Method
dvisc	0.0002964	Paxs	338.36	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.45849e+01
Coeff. B	-2.93391e+03
Coeff. C	-3.66250e+01
Temperature range (K), min.	241.83
Temperature range (K), max.	353.00

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R596364&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R596364&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions

<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>rinpola:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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