

# 2-Heptanethiol

Inchi:	InChI=1S/C7H16S/c1-3-4-5-6-7(2)8/h7-8H,3-6H2,1-2H3
InchiKey:	DAZNOIJJKASGS-UHFFFAOYSA-N
Formula:	C7H16S
SMILES:	CCCCC(C)S
Mol. weight [g/mol]:	132.27
CAS:	1628-00-2

## Physical Properties

Property code	Value	Unit	Source
gf	35.01	kJ/mol	Joback Method
hf	-154.61	kJ/mol	Joback Method
hfus	14.40	kJ/mol	Joback Method
hvap	37.52	kJ/mol	Joback Method
log10ws	-2.94		Crippen Method
logp	2.885		Crippen Method
mcvol	125.840	ml/mol	McGowan Method
pc	3025.61	kPa	Joback Method
rinpol	953.00		NIST Webbook
rinpol	953.00		NIST Webbook
ripol	1170.00		NIST Webbook
ripol	1152.00		NIST Webbook
ripol	1170.00		NIST Webbook
tb	421.98	K	Joback Method
tc	613.54	K	Joback Method
tf	190.11	K	Joback Method
vc	0.475	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	243.65	J/molxK	421.98	Joback Method
cpg	256.79	J/molxK	453.91	Joback Method
cpg	269.35	J/molxK	485.83	Joback Method
cpg	281.37	J/molxK	517.76	Joback Method

cpg	292.85	J/mol×K	549.69	Joback Method
cpg	303.81	J/mol×K	581.61	Joback Method
cpg	314.27	J/mol×K	613.54	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.44593e+01
Coeff. B	-3.78714e+03
Coeff. C	-6.25660e+01
Temperature range (K), min.	329.80
Temperature range (K), max.	476.56

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R603744&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R603744&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure

<b>pvap:</b>	Vapor pressure
<b>rinpol:</b>	Non-polar retention indices
<b>ripol:</b>	Polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

Latest version available from:

<https://www.cheméo.com/cid/76-457-1/2-Heptanethiol.pdf>

Generated by Cheméo on 2024-04-18 06:17:26.148966529 +0000 UTC m=+15710295.069543842.

Cheméo (<https://www.cheméo.com>) is the biggest free database of chemical and physical data for the process industry.