

# 2,5-Cyclohexadiene-1,4-dione, 2-chloro-

<b>Other names:</b>	2-Chloro-1,4-benzoquinone 2-Chloro-2,5-cyclohexadien-1,4-dione 2-Chloro-p-benzoquinone 2-Chloroquinone Chloro-1,4-benzoquinone Chloro-p-benzoquinone Chloroquinone Monochloro-p-benzoquinone Monochlorobenzoquinone Monochloroquinone NSC 526777 p-Benzoquinone, 2-chloro- p-Benzoquinone, chloro-
<b>Inchi:</b>	InChI=1S/C6H3ClO2/c7-5-3-4(8)1-2-6(5)9/h1-3H
<b>InchiKey:</b>	WOGWYSWDBYCVDY-UHFFFAOYSA-N
<b>Formula:</b>	C6H3ClO2
<b>SMILES:</b>	O=C1C=CC(=O)C(Cl)=C1
<b>Mol. weight [g/mol]:</b>	142.54
<b>CAS:</b>	695-99-8

## Physical Properties

Property code	Value	Unit	Source
chs	-2580.00	kJ/mol	NIST Webbook
chs	-2590.00	kJ/mol	NIST Webbook
chs	-2592.80 ± 8.40	kJ/mol	NIST Webbook
gf	-175.02	kJ/mol	Joback Method
hf	-279.56	kJ/mol	Joback Method
hfs	-231.00	kJ/mol	NIST Webbook
hfus	7.33	kJ/mol	Joback Method
hsub	69.04	kJ/mol	NIST Webbook
hvap	43.81	kJ/mol	Joback Method
log10ws	-1.14		Crippen Method
logp	0.817		Crippen Method
mvol	91.320	ml/mol	McGowan Method
pc	4652.99	kPa	Joback Method
tb	537.27	K	Joback Method
tc	795.71	K	Joback Method

tf	349.40	K	Joback Method
vc	0.341	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	218.73	J/mol×K	752.64	Joback Method
cpg	174.23	J/mol×K	537.27	Joback Method
cpg	184.11	J/mol×K	580.34	Joback Method
cpg	193.57	J/mol×K	623.42	Joback Method
cpg	202.53	J/mol×K	666.49	Joback Method
cpg	210.94	J/mol×K	709.57	Joback Method
cpg	225.83	J/mol×K	795.71	Joback Method
hsubt	69.00 ± 8.30	kJ/mol	276.50	NIST Webbook

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	2.17617e+01
Coeff. B	-7.98104e+03
Temperature range (K), min.	371.66
Temperature range (K), max.	485.16

## Sources

**NIST Webbook:**

<http://webbook.nist.gov/cgi/cbook.cgi?ID=C695998&Units=SI>

**The Yaws Handbook of Vapor Pressure:  
Crippen Method:**

<https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure>

<http://pubs.acs.org/doi/abs/10.1021/ci9903071>

**Crippen Method:**

[https://www.chemeo.com/doc/models/crippen\\_log10ws](https://www.chemeo.com/doc/models/crippen_log10ws)

**Joback Method:**

[https://en.wikipedia.org/wiki/Joback\\_method](https://en.wikipedia.org/wiki/Joback_method)

**McGowan Method:**

<http://link.springer.com/article/10.1007/BF02311772>

# Legend

<b>chs:</b>	Standard solid enthalpy of combustion
<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfs:</b>	Solid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hsub:</b>	Enthalpy of sublimation at standard conditions
<b>hsubt:</b>	Enthalpy of sublimation at a given temperature
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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