

# Butane, 2-methyl-3-(methylthio)-

<b>Other names:</b>	3,4-Dimethyl-2-thiapentane Methyl (1,2-dimethylpropyl) sulfide
<b>Inchi:</b>	InChI=1S/C6H14S/c1-5(2)6(3)7-4/h5-6H,1-4H3
<b>InchiKey:</b>	AUNQXXJGFDKEMS-UHFFFAOYSA-N
<b>Formula:</b>	C6H14S
<b>SMILES:</b>	CSC(C)C(C)C
<b>Mol. weight [g/mol]:</b>	118.24
<b>CAS:</b>	53897-51-1

## Physical Properties

Property code	Value	Unit	Source
gf	27.88	kJ/mol	Joback Method
hf	-135.86	kJ/mol	Joback Method
hfus	8.38	kJ/mol	Joback Method
hvap	34.99	kJ/mol	Joback Method
log10ws	-2.09		Crippen Method
logp	2.394		Crippen Method
mcvol	111.750	ml/mol	McGowan Method
pc	3224.64	kPa	Joback Method
rinpol	862.00		NIST Webbook
rinpol	862.00		NIST Webbook
rinpol	862.00		NIST Webbook
rinpol	862.00		NIST Webbook
tb	404.58	K	Joback Method
tc	602.65	K	Joback Method
tf	161.78	K	Joback Method
vc	0.413	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	206.45	J/mol×K	404.58	Joback Method
cpg	218.68	J/mol×K	437.59	Joback Method
cpg	230.42	J/mol×K	470.60	Joback Method

cpg	241.66	J/mol×K	503.61	Joback Method
cpg	252.43	J/mol×K	536.62	Joback Method
cpg	262.72	J/mol×K	569.63	Joback Method
cpg	272.54	J/mol×K	602.65	Joback Method

## Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.44961e+01
Coeff. B	-3.47042e+03
Coeff. C	-5.26700e+01
Temperature range (K), min.	296.92
Temperature range (K), max.	430.52

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C53897511&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C53897511&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume

<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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