

# 1-Adamantanecarboxamide, N,N-dimethyl-

**Other names:**

Tricyclo[3.3.1.1<sup>3,7</sup>]decane-1-carboxamide, N,N-dimethyl-  
N,N-Dimethyl-1-adamantanecarboxamide  
N,N-Dimethyl-1-adamantylcarboxamide

**Inchi:**

InChI=1S/C13H21NO/c1-14(2)12(15)13-6-9-3-10(7-13)5-11(4-9)8-13/h9-11H,3-8H2,1-2H

**InchiKey:**

RMRZPLJAVNSFCJ-UHFFFAOYSA-N

**Formula:**

C13H21NO

**SMILES:**

CN(C)C(=O)C12CC3CC(CC(C3)C1)C2

**Mol. weight [g/mol]:**

207.31

**CAS:**

1502-00-7

## Physical Properties

Property code	Value	Unit	Source
affp	949.40	kJ/mol	NIST Webbook
basg	917.60	kJ/mol	NIST Webbook
chs	-7733.20 ± 2.10	kJ/mol	NIST Webbook
gf	197.39	kJ/mol	Joback Method
hf	-286.10 ± 2.70	kJ/mol	NIST Webbook
hfs	-383.60 ± 2.70	kJ/mol	NIST Webbook
hfus	21.12	kJ/mol	Joback Method
hsub	97.50 ± 0.30	kJ/mol	NIST Webbook
hsub	97.50 ± 0.30	kJ/mol	NIST Webbook
hvap	51.77	kJ/mol	Joback Method
log10ws	-2.33		Crippen Method
logp	2.291		Crippen Method
mcvol	173.000	ml/mol	McGowan Method
pc	2555.92	kPa	Joback Method
tb	583.21	K	Joback Method
tc	801.50	K	Joback Method
tf	388.63	K	Joback Method
vc	0.647	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
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cpg	575.07	J/mol×K	765.12	Joback Method
cpg	484.64	J/mol×K	583.21	Joback Method
cpg	505.23	J/mol×K	619.59	Joback Method
cpg	524.36	J/mol×K	655.97	Joback Method
cpg	542.24	J/mol×K	692.36	Joback Method
cpg	559.08	J/mol×K	728.74	Joback Method
cpg	590.44	J/mol×K	801.50	Joback Method
hsubt	96.90 ± 0.30	kJ/mol	312.50	NIST Webbook

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C1502007&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C1502007&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>affp:</b>	Proton affinity
<b>basg:</b>	Gas basicity
<b>chs:</b>	Standard solid enthalpy of combustion
<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfs:</b>	Solid phase enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hsub:</b>	Enthalpy of sublimation at standard conditions
<b>hsubt:</b>	Enthalpy of sublimation at a given temperature
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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