

# Glutaric acid, hept-2-yl 3-phenoxybenzyl ester

<b>Inchi:</b>	InChI=1S/C25H32O5/c1-3-4-6-11-20(2)29-25(27)17-10-16-24(26)28-19-21-12-9-15-23(1
<b>InchiKey:</b>	HIJHXBZQBZGPQT-UHFFFAOYSA-N
<b>Formula:</b>	C25H32O5
<b>SMILES:</b>	CCCCC(C)OC(=O)CCCC(=O)OCc1cccc(Oc2ccccc2)c1
<b>Mol. weight [g/mol]:</b>	412.52

## Physical Properties

Property code	Value	Unit	Source
gf	-200.47	kJ/mol	Joback Method
hf	-724.84	kJ/mol	Joback Method
hfus	51.44	kJ/mol	Joback Method
hvap	96.79	kJ/mol	Joback Method
log10ws	-6.85		Crippen Method
logp	6.204		Crippen Method
mvol	336.340	ml/mol	McGowan Method
pc	1198.96	kPa	Joback Method
rinpol	3013.00		NIST Webbook
rinpol	3013.00		NIST Webbook
tb	1004.30	K	Joback Method
tc	1232.41	K	Joback Method
tf	588.42	K	Joback Method
vc	1.280	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	1104.72	J/molxK	1004.30	Joback Method
cpg	1118.21	J/molxK	1042.32	Joback Method
cpg	1130.08	J/molxK	1080.34	Joback Method
cpg	1140.37	J/molxK	1118.36	Joback Method
cpg	1149.13	J/molxK	1156.38	Joback Method
cpg	1156.41	J/molxK	1194.39	Joback Method
cpg	1162.23	J/molxK	1232.41	Joback Method
dvisc	0.0002284	Paxs	588.42	Joback Method

dvisc	0.0001204	Paxs	657.73	Joback Method
dvisc	0.0000717	Paxs	727.05	Joback Method
dvisc	0.0000467	Paxs	796.36	Joback Method
dvisc	0.0000326	Paxs	865.67	Joback Method
dvisc	0.0000240	Paxs	934.99	Joback Method
dvisc	0.0000184	Paxs	1004.30	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U392127&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U392127&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>m<sub>cvol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rin<sub>pol</sub>:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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