

# Glutaric acid, 2,2-dichloroethyl 2,6-dimethoxyphenyl ester

<b>Inchi:</b>	InChI=1S/C15H18Cl2O6/c1-20-10-5-3-6-11(21-2)15(10)23-14(19)8-4-7-13(18)22-9-12(16)
<b>InchiKey:</b>	QFFCWRBUIPKUGD-UHFFFAOYSA-N
<b>Formula:</b>	C15H18Cl2O6
<b>SMILES:</b>	COc1cccc(OC)c1OC(=O)CCCC(=O)OCC(Cl)Cl
<b>Mol. weight [g/mol]:</b>	365.21

## Physical Properties

Property code	Value	Unit	Source
gf	-535.57	kJ/mol	Joback Method
hf	-930.14	kJ/mol	Joback Method
hfus	40.69	kJ/mol	Joback Method
hvap	84.10	kJ/mol	Joback Method
log10ws	-3.89		Crippen Method
logp	3.126		Crippen Method
mvol	249.550	ml/mol	McGowan Method
pc	1787.88	kPa	Joback Method
rinpol	2512.00		NIST Webbook
rinpol	2512.00		NIST Webbook
tb	851.08	K	Joback Method
tc	1064.18	K	Joback Method
tf	543.89	K	Joback Method
vc	0.944	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	696.37	J/molxK	851.08	Joback Method
cpg	708.09	J/molxK	886.60	Joback Method
cpg	718.66	J/molxK	922.11	Joback Method
cpg	728.05	J/molxK	957.63	Joback Method
cpg	736.26	J/molxK	993.15	Joback Method
cpg	743.25	J/molxK	1028.67	Joback Method
cpg	749.01	J/molxK	1064.18	Joback Method
dvisc	0.0003200	Paxs	543.89	Joback Method

dvisc	0.0001986	Paxs	595.09	Joback Method
dvisc	0.0001330	Paxs	646.29	Joback Method
dvisc	0.0000944	Paxs	697.48	Joback Method
dvisc	0.0000702	Paxs	748.68	Joback Method
dvisc	0.0000543	Paxs	799.88	Joback Method
dvisc	0.0000433	Paxs	851.08	Joback Method

## Sources

<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U392002&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U392002&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307l">http://pubs.acs.org/doi/abs/10.1021/ci990307l</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>g<sub>f</sub>:</b>	Standard Gibbs free energy of formation
<b>h<sub>f</sub>:</b>	Enthalpy of formation at standard conditions
<b>h<sub>fus</sub>:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>w<sub>s</sub>:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mc<sub>vol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rin<sub>pol</sub>:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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