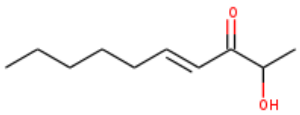
2-hydroxy-(E)-4-decen-3-one

InChI: InChI=1S/C10H18O2/c1-3-4-5-6-7-8-10(12)9(2)11/h7-9,11H,3-6H2,1-2H3/b8-7+
InChI Key: QGMOZIDHLFOEDQ-BQYQJAHWSA-N
Formula: C10H18O2
SMILES: CCCCCC=CC(=O)C(C)O
Molecular Weight: 170.25



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-154.64	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-402.60	kJ/mol	Joback Method
$\Delta_{fus} H°$	24.02	kJ/mol	Joback Method
$\Delta_{vap} H°$	60.85	kJ/mol	Joback Method
$logP_{oct/wat}$	2.073		Crippen Method
P_c	2595.13	kPa	Joback Method
T_{boil}	577.97	K	Joback Method
T_c	754.09	K	Joback Method
T_{fus}	293.13	K	Joback Method
V_c	0.595	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	383.58	J/mol×K	577.97	Joback Method
η	0.0000818	Pa×s	577.97	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method
NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C10H18O2/c1-3-4-5-6-7-8-10(12)9(2)11/h7-9,11H,3-6H2,1-2H3/b8-7+
Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
η: Dynamic viscosity (Pa×s).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).