

# Glutaric acid, 3,4-dimethylcyclohexyl 3-methylbut-2-yl ester

Inchi:	InChI=1S/C18H32O4/c1-12(2)15(5)21-17(19)7-6-8-18(20)22-16-10-9-13(3)14(4)11-16/h1
InchiKey:	QOZNEYQQEDLVKI-UHFFFAOYSA-N
Formula:	C18H32O4
SMILES:	CC(C)C(C)OC(=O)CCCC(=O)OC1CCC(C)C(C)C1
Mol. weight [g/mol]:	312.44

## Physical Properties

Property code	Value	Unit	Source
gf	-363.01	kJ/mol	Joback Method
hf	-901.37	kJ/mol	Joback Method
hfus	34.88	kJ/mol	Joback Method
hvap	73.01	kJ/mol	Joback Method
log10ws	-4.48		Crippen Method
logp	4.112		Crippen Method
mvol	268.500	ml/mol	McGowan Method
pc	1372.76	kPa	Joback Method
rinpol	2048.00		NIST Webbook
rinpol	2048.00		NIST Webbook
tb	773.15	K	Joback Method
tc	972.02	K	Joback Method
tf	405.84	K	Joback Method
vc	1.010	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	852.44	J/mol×K	773.15	Joback Method
cpg	872.38	J/mol×K	806.30	Joback Method
cpg	890.99	J/mol×K	839.44	Joback Method
cpg	908.27	J/mol×K	872.59	Joback Method
cpg	924.24	J/mol×K	905.73	Joback Method
cpg	938.88	J/mol×K	938.88	Joback Method
cpg	952.23	J/mol×K	972.02	Joback Method
dvisc	0.0017284	Paxs	405.84	Joback Method

dvisc	0.0007844	Paxs	467.06	Joback Method
dvisc	0.0004275	Paxs	528.28	Joback Method
dvisc	0.0002643	Paxs	589.50	Joback Method
dvisc	0.0001788	Paxs	650.71	Joback Method
dvisc	0.0001294	Paxs	711.93	Joback Method
dvisc	0.0000986	Paxs	773.15	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=U405431&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=U405431&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>g<sub>f</sub>:</b>	Standard Gibbs free energy of formation
<b>h<sub>f</sub>:</b>	Enthalpy of formation at standard conditions
<b>h<sub>fus</sub>:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rinpol:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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