

# 2-Furanmethanamine, tetrahydro-

<b>Other names:</b>	Furfurylamine, tetrahydro- Tetrahydro-2-furanmethanamine Tetrahydrofurfurylamine 2-Aminomethyltetrahydrofuran (.+/-)-Tetrahydrofurfurylamine USAF Q-2 UN 2943 Tetrahydro-2-furanylmethylamine NSC 76037
<b>Inchi:</b>	InChI=1S/C5H11NO/c6-4-5-2-1-3-7-5/h5H,1-4,6H2
<b>InchiKey:</b>	YNOGYQAEJGADFJ-UHFFFAOYSA-N
<b>Formula:</b>	C5H11NO
<b>SMILES:</b>	NCC1CCCO1
<b>Mol. weight [g/mol]:</b>	101.15
<b>CAS:</b>	4795-29-3

## Physical Properties

Property code	Value	Unit	Source
gf	8.10	kJ/mol	Joback Method
hf	-184.26	kJ/mol	Joback Method
hfus	15.82	kJ/mol	Joback Method
hvap	42.13	kJ/mol	Joback Method
log10ws	-0.44		Crippen Method
logp	0.124		Crippen Method
mcvol	86.300	ml/mol	McGowan Method
pc	4627.70	kPa	Joback Method
tb	428.56	K	Joback Method
tc	642.71	K	Joback Method
tf	266.84	K	Joback Method
vc	0.306	m3/kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
---------------	-------	------	-----------------	--------

cpg	181.24	J/mol×K	428.56	Joback Method
cpg	194.15	J/mol×K	464.25	Joback Method
cpg	206.34	J/mol×K	499.94	Joback Method
cpg	217.85	J/mol×K	535.63	Joback Method
cpg	228.69	J/mol×K	571.32	Joback Method
cpg	238.89	J/mol×K	607.02	Joback Method
cpg	248.48	J/mol×K	642.71	Joback Method

## Pressure Dependent Properties

Property code	Value	Unit	Pressure [kPa]	Source
tbrp	426.70	K	99.20	NIST Webbook

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C4795293&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C4795293&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mccvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tbrp:</b>	Boiling point at reduced pressure
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

Latest version available from:

<https://www.chemeo.com/cid/89-138-1/2-Furanmethanamine-tetrahydro.pdf>

Generated by Cheméo on 2024-05-02 22:50:23.894579849 +0000 UTC m=+16979472.815157164.

Cheméo (<https://www.chemeo.com>) is the biggest free database of chemical and physical data for the process industry.