

# 1,2-Benzenedicarboxylic acid, 2-ethoxy-2-oxoethyl methyl ester

**Other names:**

Phthalic acid, methyl ester, ester with ethyl glycolate  
Ethyl o-[o-(methoxycarbonyl)benzoyl]glycolate  
Methyl carbethoxymethyl phthalate  
Methyl phthalyl ethyl glycolate  
Santicizer M-17  
Ethyl o-(methoxycarbonyl)benzoyloxyacetate  
Glycolic acid, ethyl ester, methyl phthalate  
Phthalic acid, monomethyl ester, ester with ethyl glycolate  
Ethoxykarbonylmethyl-methylester kyseliny ftalove  
Ethoxycarbonylmethyl methyl phthalate  
NSC 4836  
1-(2-Ethoxy-2-oxoethyl) 2-methyl phthalate

**Inchi:** InChI=1S/C13H14O6/c1-3-18-11(14)8-19-13(16)10-7-5-4-6-9(10)12(15)17-2/h4-7H,3,8H

**InchiKey:** OYMDOVKIWFMTAW-UHFFFAOYSA-N

**Formula:** C13H14O6

**SMILES:** CCOC(=O)COC(=O)c1ccccc1C(=O)OC

**Mol. weight [g/mol]:** 266.25

**CAS:** 85-71-2

## Physical Properties

Property code	Value	Unit	Source
gf	-540.40	kJ/mol	Joback Method
hf	-820.99	kJ/mol	Joback Method
hfus	31.44	kJ/mol	Joback Method
hvap	74.94	kJ/mol	Joback Method
log10ws	-2.07		Crippen Method
logp	1.193		Crippen Method
mcvol	192.590	ml/mol	McGowan Method
pc	2470.27	kPa	Joback Method
tb	757.37	K	Joback Method
tc	970.44	K	Joback Method
tf	491.69	K	Joback Method
vc	0.728	m <sup>3</sup> /kmol	Joback Method

# Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	527.41	J/molxK	757.37	Joback Method
cpg	539.51	J/molxK	792.88	Joback Method
cpg	550.67	J/molxK	828.39	Joback Method
cpg	560.88	J/molxK	863.91	Joback Method
cpg	570.13	J/molxK	899.42	Joback Method
cpg	578.41	J/molxK	934.93	Joback Method
cpg	585.70	J/molxK	970.44	Joback Method
dvisc	0.0006662	Paxs	491.69	Joback Method
dvisc	0.0004289	Paxs	535.97	Joback Method
dvisc	0.0002953	Paxs	580.25	Joback Method
dvisc	0.0002143	Paxs	624.53	Joback Method
dvisc	0.0001623	Paxs	668.81	Joback Method
dvisc	0.0001273	Paxs	713.09	Joback Method
dvisc	0.0001026	Paxs	757.37	Joback Method

## Sources

<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>
<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C85712&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C85712&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient

<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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