

# 1-Iodo-2-trifluoromethylpropane

<b>Other names:</b>	1,1,1-Trifluoro-3-iodo-2-methylpropane
<b>Inchi:</b>	InChI=1S/C4H6F3I/c1-3(2-8)4(5,6)7/h3H,2H2,1H3
<b>InchiKey:</b>	UEEFSGKHKKFYRL-UHFFFAOYSA-N
<b>Formula:</b>	C4H6F3I
<b>SMILES:</b>	CC(Cl)C(F)(F)F
<b>Mol. weight [g/mol]:</b>	237.99
<b>CAS:</b>	26653-47-4

## Physical Properties

Property code	Value	Unit	Source
gf	-543.11	kJ/mol	Joback Method
hf	-651.38	kJ/mol	Joback Method
hfus	8.82	kJ/mol	Joback Method
hvap	29.74	kJ/mol	Joback Method
log10ws	-2.87		Crippen Method
logp	2.620		Crippen Method
mcvol	98.350	ml/mol	McGowan Method
pc	3333.53	kPa	Joback Method
tb	378.20	K	Joback Method
tc	566.53	K	Joback Method
tf	182.09	K	Joback Method
vc	0.385	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	163.64	J/mol×K	378.20	Joback Method
cpg	172.21	J/mol×K	409.59	Joback Method
cpg	180.19	J/mol×K	440.98	Joback Method
cpg	187.63	J/mol×K	472.36	Joback Method
cpg	194.53	J/mol×K	503.75	Joback Method
cpg	200.94	J/mol×K	535.14	Joback Method
cpg	206.89	J/mol×K	566.53	Joback Method
hvapt	30.40	kJ/mol	333.00	NIST Webbook

# Correlations

Information	Value
Property code	pvap
Equation	$\ln(P_{vp}) = A + B/(T + C)$
Coeff. A	1.45084e+01
Coeff. B	-3.65190e+03
Temperature range (K), min.	256.80
Temperature range (K), max.	397.08

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C26653474&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C26653474&amp;Units=SI</a>
<b>The Yaws Handbook of Vapor Pressure:</b>	<a href="https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure">https://www.sciencedirect.com/book/9780128029992/the-yaws-handbook-of-vapor-pressure</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>hvap:</b>	Enthalpy of vaporization at standard conditions
<b>hvapt:</b>	Enthalpy of vaporization at a given temperature
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>pvap:</b>	Vapor pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature

**tf:** Normal melting (fusion) point

**vc:** Critical Volume

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