

# 3-chloro-2-ethoxy-tetrahydro-pyran

<b>Inchi:</b>	InChI=1S/C7H13ClO2/c1-2-9-7-6(8)4-3-5-10-7/h6-7H,2-5H2,1H3
<b>InchiKey:</b>	OZJLVSAMSVIOFG-UHFFFAOYSA-N
<b>Formula:</b>	C7H13ClO2
<b>SMILES:</b>	CCOC1OCCCC1Cl
<b>Mol. weight [g/mol]:</b>	164.63

## Physical Properties

Property code	Value	Unit	Source
gf	-178.25	kJ/mol	Joback Method
hf	-433.79	kJ/mol	Joback Method
hfus	20.16	kJ/mol	Joback Method
hvap	42.60	kJ/mol	Joback Method
log10ws	-1.69		Crippen Method
logp	1.767		Crippen Method
mvol	122.610	ml/mol	McGowan Method
pc	3152.62	kPa	Joback Method
rinpol	1085.00		NIST Webbook
rinpol	1085.00		NIST Webbook
tb	461.24	K	Joback Method
tc	669.78	K	Joback Method
tf	250.51	K	Joback Method
vc	0.448	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
cpg	258.49	J/molxK	461.24	Joback Method
cpg	327.87	J/molxK	635.02	Joback Method
cpg	315.35	J/molxK	600.27	Joback Method
cpg	302.15	J/molxK	565.51	Joback Method
cpg	288.27	J/molxK	530.75	Joback Method
cpg	273.72	J/molxK	496.00	Joback Method
cpg	339.72	J/molxK	669.78	Joback Method
dvisc	0.0003039	Paxs	461.24	Joback Method

dvisc	0.0003880	Paxs	426.12	Joback Method
dvisc	0.0005175	Paxs	391.00	Joback Method
dvisc	0.0007306	Paxs	355.88	Joback Method
dvisc	0.0011124	Paxs	320.75	Joback Method
dvisc	0.0018782	Paxs	285.63	Joback Method
dvisc	0.0036730	Paxs	250.51	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=R132935&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=R132935&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci9903071">http://pubs.acs.org/doi/abs/10.1021/ci9903071</a>
<b>Crippen Method:</b>	<a href="https://www.chemeo.com/doc/models/crippen_log10ws">https://www.chemeo.com/doc/models/crippen_log10ws</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cp<sub>g</sub>:</b>	Ideal gas heat capacity
<b>dvisc:</b>	Dynamic viscosity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>h<sub>vap</sub>:</b>	Enthalpy of vaporization at standard conditions
<b>log<sub>10</sub>ws:</b>	Log <sub>10</sub> of Water solubility in mol/l
<b>log<sub>p</sub>:</b>	Octanol/Water partition coefficient
<b>m<sub>cvol</sub>:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>rin<sub>pol</sub>:</b>	Non-polar retention indices
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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