

# Piperonylnitrile

<b>Other names:</b>	1,3-Benzodioxole-5-carbonitrile 3,4-Methylene dioxy benzonitrile 3,4-methylenedioxybenzonitrile 5-Cyano-1,3-benzodioxole Benzonitrile, 3,4-methylenedioxy- benzo-1,3-dioxole-5-carbonitrile
<b>Inchi:</b>	InChI=1S/C8H5NO2/c9-4-6-1-2-7-8(3-6)11-5-10-7/h1-3H,5H2
<b>InchiKey:</b>	PKRWWZCDLJSJIF-UHFFFAOYSA-N
<b>Formula:</b>	C8H5NO2
<b>SMILES:</b>	N#Cc1ccc2c(c1)OCO2
<b>Mol. weight [g/mol]:</b>	147.13
<b>CAS:</b>	4421-09-4

## Physical Properties

Property code	Value	Unit	Source
gf	139.03	kJ/mol	Joback Method
hf	-0.84	kJ/mol	Joback Method
hfus	20.79	kJ/mol	Experimental and Computational Thermochemistry of 1,3-Benzodioxole Derivatives
hvap	56.72	kJ/mol	Joback Method
log10ws	-2.04		Crippen Method
logp	1.287		Crippen Method
mcvol	102.080	ml/mol	McGowan Method
pc	4082.92	kPa	Joback Method
tb	586.47	K	Joback Method
tc	833.00	K	Joback Method
tf	371.69	K	Joback Method
vc	0.402	m <sup>3</sup> /kmol	Joback Method

## Temperature Dependent Properties

Property code	Value	Unit	Temperature [K]	Source
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cpg	232.23	J/mol×K	586.47	Joback Method
cpg	241.07	J/mol×K	627.56	Joback Method
cpg	249.18	J/mol×K	668.65	Joback Method
cpg	256.62	J/mol×K	709.74	Joback Method
cpg	263.46	J/mol×K	750.82	Joback Method
cpg	269.77	J/mol×K	791.91	Joback Method
cpg	275.62	J/mol×K	833.00	Joback Method

## Sources

<b>NIST Webbook:</b>	<a href="http://webbook.nist.gov/cgi/cbook.cgi?ID=C4421094&amp;Units=SI">http://webbook.nist.gov/cgi/cbook.cgi?ID=C4421094&amp;Units=SI</a>
<b>Crippen Method:</b>	<a href="http://pubs.acs.org/doi/abs/10.1021/ci990307l">http://pubs.acs.org/doi/abs/10.1021/ci990307l</a>
<b>Crippen Method:</b>	<a href="https://www.cheméo.com/doc/models/crippen_log10ws">https://www.cheméo.com/doc/models/crippen_log10ws</a>
<b>Experimental and Computational Thermochemistry of 1,3-Benzodioxole Derivatives:</b>	<a href="https://www.doi.org/10.1021/je700035m">https://www.doi.org/10.1021/je700035m</a>
<b>Joback Method:</b>	<a href="https://en.wikipedia.org/wiki/Joback_method">https://en.wikipedia.org/wiki/Joback_method</a>
<b>McGowan Method:</b>	<a href="http://link.springer.com/article/10.1007/BF02311772">http://link.springer.com/article/10.1007/BF02311772</a>

## Legend

<b>cpg:</b>	Ideal gas heat capacity
<b>gf:</b>	Standard Gibbs free energy of formation
<b>hf:</b>	Enthalpy of formation at standard conditions
<b>hfus:</b>	Enthalpy of fusion at standard conditions
<b>h vap:</b>	Enthalpy of vaporization at standard conditions
<b>log10ws:</b>	Log10 of Water solubility in mol/l
<b>logp:</b>	Octanol/Water partition coefficient
<b>mcvol:</b>	McGowan's characteristic volume
<b>pc:</b>	Critical Pressure
<b>tb:</b>	Normal Boiling Point Temperature
<b>tc:</b>	Critical Temperature
<b>tf:</b>	Normal melting (fusion) point
<b>vc:</b>	Critical Volume

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