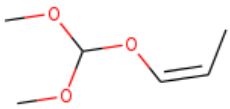
1-Propene, 1-(dimethoxymethoxy)-, (E)-

InChI: InChI=1S/C6H12O3/c1-4-5-9-6(7-2)8-3/h4-6H,1-3H3/b5-4-
InChI Key: DDINHWIZOMBDDD-PLNGDYQASA-N
Formula: C6H12O3
SMILES: CC=COC(OC)OC
Molecular Weight: 132.16
CAS: 66178-20-9



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-237.58	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-451.89	kJ/mol	Joback Method
$\Delta_{fus} H°$	11.54	kJ/mol	Joback Method
$\Delta_{vap} H°$	35.75	kJ/mol	Joback Method
$logP_{oct/wat}$	1.11		Crippen Method
P_c	3138.51	kPa	Joback Method
T_{boil}	407.66	K	Joback Method
T_c	585.73	K	Joback Method
T_{fus}	203.99	K	Joback Method
V_c	0.40	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	211.99	J/mol×K	407.66	Joback Method
η	0.00	Pa×s	407.66	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method
NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C6H12O3/c1-4-5-9-6(7-2)8-3/h4-6H,1-3H3/b5-4-
Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
η: Dynamic viscosity (Pa×s).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).