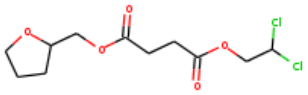
Succinic acid, 2,2-dichloroethyl tetrahydrofurfuryl ester

InChI: InChI=1S/C11H16Cl2O5/c12-9(13)7-18-11(15)4-3-10(14)17-6-8-2-1-5-16-8/h8-9H,1-7H2
InChI Key: OGHLWWUVYCQWQU-UHFFFAOYSA-N
Formula: C11H16Cl2O5
SMILES: O=C(CCC(=O)OCC1CCCO1)OCC(Cl)Cl
Molecular Weight: 299.15



Physical Properties

Property	Value	Unit	Source
$\Delta_f G°$	-501.97	kJ/mol	Joback Method
$\Delta_f H°_{gas}$	-868.25	kJ/mol	Joback Method
$\Delta_{fus} H°$	36.60	kJ/mol	Joback Method
$\Delta_{vap} H°$	71.54	kJ/mol	Joback Method
$logP_{oct/wat}$	1.84		Crippen Method
P_c	2318.07	kPa	Joback Method
T_{boil}	720.31	K	Joback Method
T_c	931.13	K	Joback Method
T_{fus}	440.36	K	Joback Method
V_c	0.75	m^3/kg-mol	Joback Method

Temperature Dependent Properties

Property	Value	Unit	Temperature (K)	Source
$C_{p,gas}$	540.18	J/mol×K	720.31	Joback Method
η	0.00	Pa×s	720.31	Joback Method

Sources

Joback Method: https://en.wikipedia.org/wiki/Joback_method
NIST Webbook: http://webbook.nist.gov/cgi/inchi/InChI=1S/C11H16Cl2O5/c12-9(13)7-18-11(15)4-3-10(14)17-6-8-2-1-5-16-8/h8-9H,1-7H2
Crippen Method: http://pubs.acs.org/doi/abs/10.1021/ci990307l

Legend

$\mathbf{C_{p,gas}}$: Ideal gas heat capacity (J/mol×K).
η: Dynamic viscosity (Pa×s).
$\Delta_f\mathbf{G°}$: Standard Gibbs free energy of formation (kJ/mol).
$\Delta_f\mathbf{H°_{gas}}$: Enthalpy of formation at standard conditions (kJ/mol).
$\Delta_{fus}\mathbf{H°}$: Enthalpy of fusion at standard conditions (kJ/mol).
$\Delta_{vap}\mathbf{H°}$: Enthalpy of vaporization at standard conditions (kJ/mol).
$\boldsymbol{log}\mathbf{P_{oct/wat}}$: Octanol/Water partition coefficient .
$\mathbf{P_c}$: Critical Pressure (kPa).
$\mathbf{T_{boil}}$: Normal Boiling Point Temperature (K).
$\mathbf{T_c}$: Critical Temperature (K).
$\mathbf{T_{fus}}$: Normal melting (fusion) point (K).
$\mathbf{V_c}$: Critical Volume (m^3/kg-mol).